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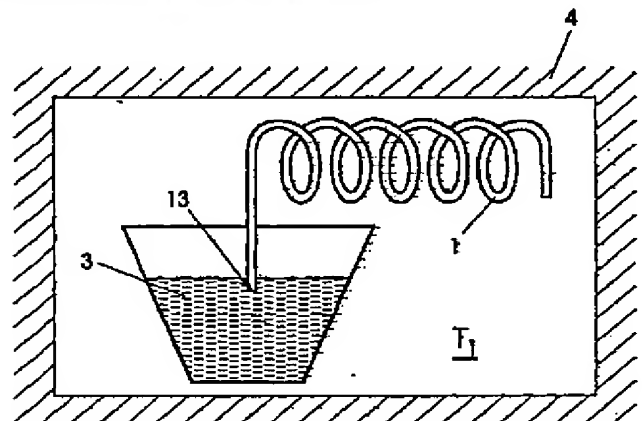
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⑤⑤ **Für die Beurteilung der Patentfähigkeit in Betracht
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GB 20 09 137 A

⑤④ **Lichtleiter zum Leiten von Licht im Infrarotbereich und Verfahren zu seiner Herstellung**

⑤⑦ Lichtleiter zum Leiten von Licht im Infrarotbereich von $\geq 2,5 \mu\text{m}$, bestehend aus einem Kern (12) und einem Mantel (11), wobei der Mantel (11) als Kapillare oder als dünnes Rohr (1) ausgebildet ist und aus einem Material besteht, das eine Viskosität zwischen 10^7 bis 10^{13} dPas bei einer Temperatur T_1 aufweist, bei der ein den Kern bildendes Material bereits in Schmelze vorliegt, wobei das Kernmaterial entweder aus Silizium oder aus Germanium besteht.



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Beschreibung

[0001] Die Erfindung betrifft einen Lichtleiter zum Leiten von Licht im Infrarotbereich bei Wellenlängen $\geq 2,5 \mu\text{m}$, welcher bevorzugt bei Laseranwendungen, wo das Infrarotlicht zur Materialbearbeitung oder -veränderung dient, zum Einsatz gelangt. Insbesondere sollen erfindungsgemäße Lichtleiter in der Laserchirurgie Anwendung finden.

[0002] Als Materialien für Lichtleiter für Wellenlängen oberhalb $2,5 \mu\text{m}$ sind beispielsweise Fluoridgläser (ZrF_4 , HfF_4), Chalkogenid-Gläser (AsS , AsSe , AsSeTe), Erdalkali-halogenid-Gläser (BeF_2 , ZnCl_2), Silberhalogenidfasern (AgCl/Br), Thallium/Cäsiumhalogenid-Fasern als geeignet bekannt. Daraus gebildete Lichtleitfasern haben für den vorgesehenen bevorzugten Anwendungsfall mindestens einen, meist sogar mehrere der folgenden Nachteile:

die Materialien bzw. daraus gebildete Lichtleiter sind

- toxisch,
- chemisch unbeständig (Wasserlöslichkeit/Hygro-skopizität),
- lichtempfindlich, wobei sie zur Degradation unter der zu leitenden IR-Strahlung neigen,
- mechanisch empfindlich, so daß sie nur relativ geringe Krümmungen erlauben,

[0003] Aus diesen Gründen konnten sich derartige Lichtleitfasern bislang in der Lasermedizin nicht durchsetzen, wo neben einer erforderlichen Toxizitätsfreiheit auch Krümmungsradien im Zentimeterbereich gegeben sein müssen. Andere konstruktive Ausführungen zur Lichtleitung, wie mit einer Flüssigkeit (CCl_4) gefüllte Schläuche erweisen sich hierbei wegen der Toxizität als ebenso unbrauchbar, wie innenverspiegelte Metallrohre, die nur geringe Krümmungen ermöglichen.

[0004] In GB 2 009 137 A ist ein Verfahren zur Herstellung IR-leitfähiger Fasern, nämlich das Taylor-Verfahren, beschrieben, bei dem das Kernmaterial in Form von Mischungen aus verschiedenen Materialien in ein Glasrohr mit einem sehr viel größeren Durchmesser als der späteren Lichtleitfaser eingefüllt und erschmolzen wird. Erst danach erfolgt ein Ausziehen des Körpers zu einer Faser unter weiterer Wärmeeinwirkung, wobei es aufgrund der langen Zeiten zu starken chemischen Reaktionen zwischen Kern- und Mantelmaterial kommt. Darüber hinaus reagieren Silizium und Quarzglas bei Temperaturen oberhalb ca. 1200°C zu gasförmigen SiO , weshalb bis heute keine Fasern unter Einsatz von Kernmaterialien aus reinem Germanium oder Silizium bekannt geworden sind.

[0005] Der Erfindung liegt die Aufgabe zugrunde, einen Lichtleiter zum Leiten von Licht im Infrarotbereich von $\geq 2,5 \mu\text{m}$ anzugeben, der toxikologisch unbedenklich, wasserunlöslich und chemisch beständig ist und der eine solche mechanische Festigkeit besitzt, daß er bei Durchmessern für Lichtleiter, wie sie bspw. in der Laserchirurgie in Betracht kommen, Krümmungsradien im Zentimeterbereich standhält. Auch sollen durch die Erfindung Lichtleiter bereitgestellt werden, die, bei Verzicht auf die gewünschte Verbiegbarkeit, ansonsten die gleichen oben genannten Eigenschaften aufweisen. Weiterhin liegt der Erfindung die Aufgabe zugrunde, ein Verfahren zur Herstellung eines solchen Lichtleiters anzugeben.

[0006] Die Aufgabe der Erfindung wird durch die Merkmale der Patentansprüche 1 und 8 gelöst. Vorteilhafte weitere Ausgestaltungen sind durch die nachgeordneten Ansprüche erfaßt.

[0007] Das Wesen der Erfindung besteht darin, dünne Rohre, insbesondere Kapillaren, aus einem Material, wel-

ches Licht im sichtbaren Spektralbereich leitet, mit schmelzflüssigem Germanium oder Silizium zu füllen und anschließend abzukühlen.

[0008] Die Erfindung soll nachstehend anhand von Ausführungsbeispielen näher erläutert werden. Es zeigen:

[0009] Fig. 1 einen Schnitt durch einen Teil des vorgeschlagenen neuen Lichtleiters und

[0010] Fig. 2 schematisch den Aufbau einer Anordnung zur Herstellung des genannten Lichtleiters.

[0011] In einem ersten Ausführungsbeispiel findet eine Kapillare 1, bestehend aus einem Glas, wie z. B. hochschmelzende Borosilicatgläser, Erdalkali-Alumo-Silicatgläser oder Übergangsgläser, Anwendung. Die Materialauswahl für das für die Kapillare 1 in Betracht kommende Glas erfolgt unter der Maßgabe, daß dieses bei einer Temperatur T_1 , die geringfügig oberhalb der Schmelztemperatur des den späteren Kern 12 bildenden Kernmaterials liegt, eine Viskosität zwischen 10^7 bis 10^{13} dPas annimmt, und daß es bei den weiter unten beschriebenen Erstarrungszeiten des Kernmaterials zu keinen störenden chemischen Reaktionen zwischen Kern- und Kapillarenmaterial kommt. Im Rahmen der Erfindung ist der Kapillareninnendurchmesser d_2 in einem Bereich von $10 \mu\text{m}$ bis einige Millimeter, bevorzugt von $10 \mu\text{m}$ bis $200 \mu\text{m}$ festlegbar, wobei der Kapillarenwandung 11, insbesondere bei kleinen Kapillareninnendurchmessern, bspw. von 10 – $100 \mu\text{m}$, jeweils eine Wandstärke d_1 gegeben ist, die in der Größenordnung des Kapillareninnendurchmessers d_2 liegt.

[0012] Bei diesem ersten Ausführungsbeispiel wird in einem geeigneten Ofen 4 in einem Tiegel eine Germaniumschmelze bei einer Temperatur T_1 wenig oberhalb ihres Schmelzpunktes vorrätig gehalten. Die nach obigen Maßgaben ausgewählte Glaskapillare 1 wird in den Ofen 4 eingebracht und dadurch ebenfalls auf die Temperatur T_1 erwärmt. Genannte Temperatur T_1 liegt in diesem Ausführungsbeispiel oberhalb 940°C . Bei Erreichen der vorgebbaren Temperatur T_1 wird ein Ende 13 der Kapillare 1 (siehe Fig. 2) in die Germaniumschmelze 3 getaucht. Je nach gewähltem Kapillareninnendurchmesser d_2 wird die Schmelze 3 durch Kapillarwirkung oder mit Hilfe einer nicht dargestellten Pumpe in die Kapillare 1 gesaugt. Ist die Kapillare 1 mit der Schmelze 3 in einer gewünschten Länge gefüllt, wird das Ende 13 aus dem Schmelzbehälter entfernt und die gefüllte Kapillare 1 einer Abkühlung unterworfen, wobei das Kernmaterial 12 erstarrt und der gewünschte Lichtleiter gebildet ist.

[0013] In einem weiteren Ausführungsbeispiel ist die Kapillare 1 aus Quarzglas gebildet und das Kernmaterial 12 soll aus Silizium bestehen. Die Art der Herstellung des Lichtleiters erfolgt analog zum ersten Ausführungsbeispiel, wobei in diesem Fall T_1 oberhalb 1420°C zu wählen ist.

[0014] Wie weit in jedem Fall die Temperatur T_1 oberhalb der Schmelztemperatur des Kernmaterials gewählt werden kann, ist vornehmlich abhängig davon, welche Viskosität das Kapillarenmaterial in diesem Temperaturbereich annimmt.

[0015] Die Materialauswahl für die im Ausführungsbeispiel verwendeten Kapillaren 1 ist für den bevorzugten Verwendungszweck besonders vorteilhaft, beschränkt die Erfindung jedoch nicht einzig darauf. Durch die vorgeschlagenen Materialien für die zum Einsatz gelangenden Kapillaren und die Dimensionierung des Kapillarenmantels 11 in bezug auf den Kapillareninnendurchmesser d_2 ist eine Leitung von sichtbarem Licht im Mantel 11 der Lichtleitfaser gewährleistet, was für das bevorzugte Anwendungsgebiet zu Justierzwecken von erheblichem Vorteil ist, wohingegen im Kern 12 des Lichtleiters infrarotes Licht geleitet wird, dessen Wellenlänge im Falle der Verwendung von Silizium zwi-

schen 2 μm –5 μm und im Falle von Germanium zwischen 3 μm –13 μm liegt.

[0016] Die mechanische Festigkeit der erhaltenen Lichtleiter ist so groß, daß mit einer Faser mit einem Kerndurchmesser von bspw. 50 μm ein Krümmungsradius von 1 cm durchaus möglich ist. Silizium und Germanium sind chemisch beständig, nicht wasserlöslich und lichtunempfindlich.

[0017] Durch die Erfindung werden Lichtleiter erhalten, die die Nachteile der bekannten IR-Fasern nicht aufweisen.

[0018] Auch liegt es im Rahmen der Erfindung, für den Fall, daß der Kapillarenmantel 11 bei speziellen anderweitigen Verwendungen der Lichtleitfaser nicht erforderlich ist, diesen bspw. durch einen Ätzschritt zu entfernen. Ebenso liegt es im Rahmen der Erfindung, wenn für bestimmte Verwendungen des Lichtleiters darauf verzichtet werden kann, daß dieser genannten Krümmungen schadlos standhält, diesen stab- bzw. taperförmig mit einem entsprechend größeren Durchmesser zu fertigen als auch der Kapillare eine vorgebbare Krümmung im noch plastischen Zustand aufzuprägen.

Patentansprüche

1. Lichtleiter zum Leiten von Licht im Infrarotbereich von $\geq 2,5 \mu\text{m}$, bestehend aus einem Kern (12) und einem Mantel (11), wobei der Mantel (11) als Kapillare oder als dünnes Rohr (1) ausgebildet ist und aus einem Material besteht, das eine Viskosität zwischen 10^7 bis 10^{13} dPas bei einer Temperatur T_1 aufweist, bei der ein den Kern bildendes Material bereits in Schmelze vorliegt, wobei das Kernmaterial entweder aus Silizium oder aus Germanium besteht.
2. Lichtleiter nach Anspruch 1, dadurch gekennzeichnet, daß dem Mantel (11) der Kapillare eine Wandstärke (d_1) gegeben ist, die bei kleinen Kapillareninnendurchmessern, in der Größenordnung von 10 bis 100 μm , in der Größenordnung des Durchmessers (d_2) des Kerns (12) liegt.
3. Lichtleiter nach Anspruch 1 oder 2, dadurch gekennzeichnet, daß der Mantel (11) aus einem Material gefertigt ist, welches in einem anderen Spektralbereich als der Kern (12) Licht leitet.
4. Lichtleiter nach einem der vorstehenden Ansprüche, dadurch gekennzeichnet, daß der Mantel (11) aus einem Glas besteht, wenn das Kernmaterial durch Germanium gebildet ist.
5. Lichtleiter nach einem der Ansprüche 1 bis 3, dadurch gekennzeichnet, daß der Mantel (11) aus Quarzglas besteht, wenn das Kernmaterial durch Silizium gebildet ist.
6. Lichtleiter nach einem der vorstehenden Ansprüche, dadurch gekennzeichnet, daß die Wandstärke (d_1) des Mantels (11) und der Durchmesser (d_2) des Kerns (12) im Bereich von 10 μm bis 1 mm festgelegt sind.
7. Lichtleiter nach einem der vorstehenden Ansprüche, dadurch gekennzeichnet, daß die Wandstärke (d_1) des Mantels (11) und der Durchmesser (d_2) des Kerns (12) im Bereich von 10 μm bis 200 μm festgelegt sind.
8. Verfahren zur Herstellung eines Lichtleiters zum Leiten von Licht im Infrarotbereich von $\geq 2,5 \mu\text{m}$, bei dem ein den IR-lichtleitenden Kern (12) bildendes Material (3) in eine Schmelze überführt wird, eine Kapillare oder ein dünnes Rohr (1), bestehend aus einem Glas oder Quarz ausgewählt wird, welches bei einer Temperatur T_1 , die geringfügig oberhalb der Schmelztemperatur des Kernmaterials liegt, eine Viskosität zwischen 10^7 bis 10^{13} dPas annimmt,

diese Kapillare oder dieses dünne Rohr (1) auf die genannte Temperatur T_1 erwärmt wird und bei Erreichen der Temperatur T_1 einseitig in das geschmolzene Material (3) eingetaucht und dieses in die Kapillare oder das Rohr (1) eingesaugt wird und die Kapillare oder das Rohr (1) bis zur gewünschten Verfüllung auf der Temperatur T_1 gehalten und anschließend im verfüllten Zustand einer unmittelbaren Abkühlung unterworfen wird.

9. Verfahren nach Anspruch 8, dadurch gekennzeichnet, daß für das den Kern (12) bildende Material Silizium und für die Kapillare oder das dünne Rohr (1) Quarzglas eingesetzt wird.

10. Verfahren nach Anspruch 8, dadurch gekennzeichnet, daß für das den Kern (12) bildende Material Germanium und für die Kapillare oder das dünne Rohr (1) ein Glas eingesetzt wird.

Hierzu 1 Seite(n) Zeichnungen

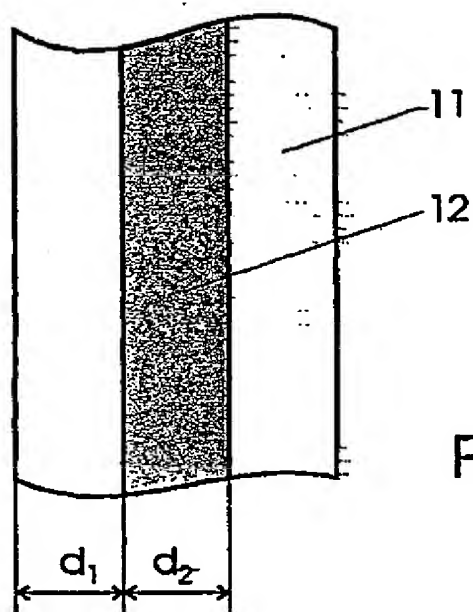


Fig. 1

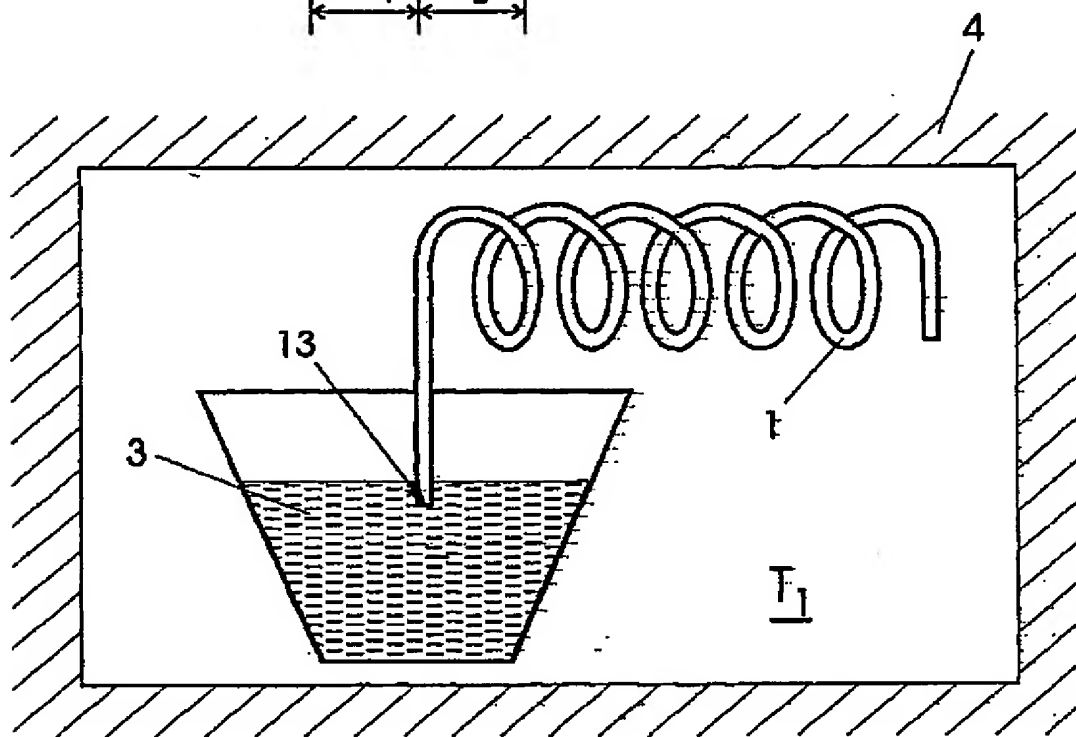
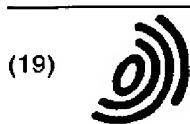


Fig. 2



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(54) Apparatus and Method of Modifying the Birefringence in Optical Fibres

(57) The invention relates to optical fibre devices and methods of tuning the birefringence of waveguides, especially to all-fibre optical devices which have undergone microstructuring. Embodiments of the invention include an optical fibre device (10) such as a tunable birefringent optical fibre having a core region (12), a cladding layer (14) therearound, and a controllable active material (18) selectively disposed in, for example, capillaries or pockets (16) formed in the cladding layer. The active materials include, for example, electro-optic material, magneto-optic material, photorefractive material,

thermo-optic material and/or materials such as laser dyes that provide tunable gain or loss. The application of, for example, temperature, light or an electric or magnetic field modifies the optical properties of the active material, which, in turn, modifies the propagation properties of optical signals in the device. The optical device can include a tapered region or long period grating that causes the core mode to spread or couple into the cladding region and, simultaneously, allows the active material to be relatively close to the propagated modes, thus allowing interaction between the active material and the propagating modes.

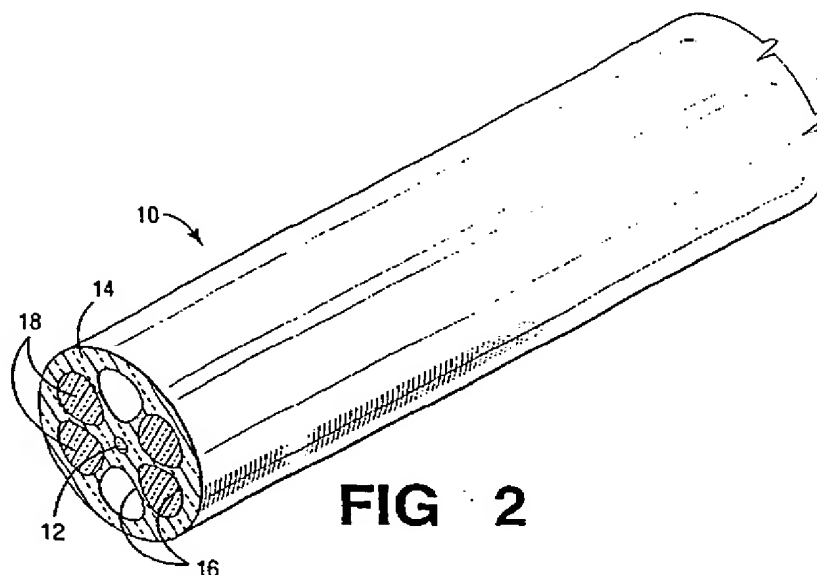


FIG. 2

Description**Field of the Invention**

[0001] The invention relates to optical fiber devices and methods for tuning the birefringence of waveguides (polarization of optical signals) using such devices and methods. More particularly, the invention relates to all-fiber optical devices, methods and systems for modifying the birefringence in microstructured optical fibers.

2. Background of the Invention

[0002] Optical devices that modify the properties of optical signals include devices such as modulators, attenuators, and polarization controllers. Such devices use various means to vary refractive properties of one or more regions of the device to change the amplitude/phase of a signal propagating through the device. Conventional devices of this kind use lithium niobate, electroabsorption and/or other configurations to affect the propagation properties (for example, attenuation, birefringence) of an optical fiber or other waveguide arrangement. Typically, conventional modulators and other devices such as polarization controllers are relatively costly and bulky devices that introduce an undesirable amount of loss. However, all-fiber devices inherently exhibit lower loss and are easier to package than conventional devices.

[0003] For example, with birefringent optical fibers (or polarization maintaining fibers), the birefringence is categorized as geometrical birefringence or stress-induced birefringence. Geometrical birefringence is produced by the axially asymmetrical core or structures surrounding the core. Stress induced birefringence is generated by a non-symmetrical stress in the core, where stress applying rods with larger thermal expansion coefficients than silica are introduced in the vicinity of the core. For example, see Fig. 1. However, the configuration of such devices makes them relatively difficult to attain tunability in birefringence.

[0004] Accordingly, it is desirable to have available an all-fiber, optical device such as an optical birefringence device that has enhanced tunable birefringence, has easier manufacturability and generally is smaller in size than conventional polarization controllers and modulators, and is relatively easy to splice to conventional fibers.

Summary of the Invention

[0005] The invention is as defined by the claims. Embodiments of the invention include an optical fiber device such as a tunable polarization dependent loss element, a polarization controller or a system for use therein, and a method for making the optical device. According to embodiments of the invention, polarization of light is manipulated in microstructured optical fibers. The ability to control light propagation is achieved by establishing spatial asymmetry within a device such as a microstructured optical fiber, for example, by filling selected pockets or air-holes formed in the microstructured optical fiber with active/tunable materials. The fiber includes of a core and a surrounding cladding layer. Pockets are introduced in the cladding region and extend, for example, in the axial direction of the fiber. Active materials are infused in the pockets to change the optical properties of the optical signal. The active materials include, for example, electro-optic material, magneto-optic material, photo-refractive material and thermo-optic material. Those materials change their intrinsic optical properties such as their refractive index according to external field applied (for example, temperature). That is, the application of, for example, temperature, light (optical field) or an electric or magnetic field varies optical properties such as refractive index, loss, scattering, or birefringence of the active material, which, in turn, varies or affects the propagation properties of optical signals in the device.

[0006] According to an embodiment of the invention, the optical device includes a tapered region that reduces the diameter of the fiber device but maintains the relative dimensional proportions of the cross-sectional index as in the non-tapered regions. In the tapered region, the mode field is not supported by the doped core and spreads into the cladding region, where it interacts with the active materials. Simultaneously, the tapered region allows the active material to be physically closer to the propagated modes compared to conventional arrangements, thus allowing interaction between the active material and the propagating modes. The tapered region also is designed such that the fiber can be spliced to conventional optical fibers with relatively low insertion loss.

[0007] According to an embodiment of the invention, the optical device includes a grating such as a Fiber Bragg grating (FBG) or a long period grating (LPG) written in the photosensitive core of the optical fiber. The FBG or LPG includes periodic perturbations in the index of the core and, like the tapered region, enhances the interaction between the optical signal and the active material. The FBG or LPG couples light from the core mode of the fiber into a mode whose field distribution is spread in the cladding, which is sensitive to the change in the refractive index at the silica-air-holes interface. In this manner, the coupled light interacts with the active material.

[0008] To induce birefringence in the optical signal, materials are infused in the air-holes in such a way to break the axial symmetry of the fiber. Axial asymmetry is achieved by sealing specific air holes with epoxy, so that only the open

air holes are filled with polymer. In this manner, there will be a difference in the propagation constants of the orthogonally polarized modes, which is exhibited as birefringence in the fiber. Control (generating and tuning) of birefringence is achieved by filling selective air-holes with active material and by applying an external field to change the refractive index of the material.

[0009] For example, temperature-dependent polymers are infused in one or many holes to provide different birefringence. Such arrangements exploit enhanced tunability because the refractive index of the polymers have relatively large temperature dependences. Birefringence tuning is achieved by changing the index of the material (for example, thermally in the case of polymers). Also, birefringence tuning is achieved electrically by using liquid crystals, whose refractive index varies electrically, in one or many of the holes.

[0010] According to other embodiments of the invention, birefringence tuning is achieved by infusing different materials in different holes (for example, $dn/dT > 0$ for some holes, $dn/dT < 0$ for other holes). Also, birefringence is turned on/off by electrically driving micro-fluids, in one or different holes again, into the waist of the taper or in the region where a LPG is written in the core of the fiber.

Brief Description of the Drawings

[0011] In the drawings:

Fig. 1 is a simplified, cross-sectional view of a conventional birefringent fiber;

Fig. 2 is a side, partial cross-sectional view of an optical fiber device in which embodiments of the invention are applicable;

Fig. 3a is a side, partial cross-sectional view of a tapered microstructured optical fiber having a symmetric mode field (that is, all six holes are filled with active material);

Fig. 3b is an end, cross-sectional view of the fiber of Fig. 3a;

Fig. 3c is a corresponding Poincare sphere for the fiber of Fig. 3a;

Fig. 4a is a side, partial cross-sectional view of a tapered microstructured optical fiber having an asymmetric mode field (for example, only two holes are filled with active material) according to an embodiment of the invention;

Fig. 4b is an end, cross-sectional view of the fiber of Fig. 4a;

Fig. 4c is a corresponding Poincare sphere for the fiber of Fig. 4a;

Fig. 5a is a graphical diagram of the calculated birefringence (and mode fields) as a function of temperature or index of polymer of a microstructured optical fiber having an asymmetric mode field;

Fig. 5b is a graphical diagram of the measured birefringence as a function of temperature or index of polymer of a microstructured optical fiber having an asymmetric mode field;

Fig. 6a is a side, partial cross-sectional view of a microstructured optical fiber having a long period grating written therein and having a symmetric mode field (that is, all six holes are filled with active material);

Fig. 6b is a graphical diagram of the corresponding LPG transmission spectra at different temperatures for the orthogonal polarization states (solid and dotted lines) of the microstructured optical fiber of Fig. 6a;

Fig. 7a is a side, partial cross-sectional view of a tapered microstructured optical fiber having a long period grating written therein and having an asymmetric mode field (for example, only two holes are filled with active material);

Fig. 7b is a graphical diagram of the corresponding LPG transmission spectra at different temperatures for the orthogonal polarization states (solid and dotted lines) of the microstructured optical fiber of Fig. 7a;

Fig. 8 is a graphical diagram of the wavelength splitting of the resonance ($\delta\lambda$) as a function of temperature for microstructure optical fiber arrangements in which all pockets are filled with index tunable active material (- -), two adjacent pockets are filled with index tunable active material (-Δ-), three alternate pockets are filled with index tunable active material (-◊-), and two opposite pockets are filled with index tunable active material (-O-);

Fig. 9a is a simplified schematic diagram of an optical system in which embodiments of the invention are useful; Fig. 9b is another simplified schematic diagram of an optical system in which embodiments of the invention are useful; and

Fig. 10 is a simplified block diagram of a method for making optical fiber devices according to embodiments of the invention.

Detailed Description

[0012] In the following description similar components are referred to by the same reference numeral to enhance the understanding of the invention through the description of the drawings. Also, unless otherwise explicitly specified herein, the drawings are not drawn to scale.

[0013] Although specific features, configurations and arrangements are discussed hereinbelow, it should be understood that such is done for illustrative purposes only. A person skilled in the relevant art will recognize that other steps,

configurations and arrangements are useful without departing from the spirit and scope of the invention.

[0014] Referring now to Fig. 2, shown is a side, partial cross-sectional view of an optical fiber device 10 in which embodiments of the invention are applicable. The optical fiber device 10 includes an optical fiber comprised of a core or core region 12 surrounded by a cladding region or layer 14. The core region 12 is made of, for example, silica (that is, silicon dioxide, SiO_2) doped with germanium or other suitable material. The cladding layer 14, which has a refractive index less than that of the core region 12, typically is made of silica or silica doped with fluorine or other suitable material. The optical fiber is, for example, an air-silica, microstructured fiber (MF) or microstructured optical fiber (MOF).

[0015] Multimode fiber supports many different modes (paths) of a particular wavelength, whereas singlemode fiber is designed to support only the fundamental mode (LP_{01}) of a particular wavelength. For multimode fiber, the diameter of the core region typically is approximately 50 or approximately 62.5 μm ; for singlemode fiber, the diameter of the core region typically is less than approximately 10 microns (μm), for example, 5-8 μm . For both singlemode and multimode fibers, the total diameter of the cladding layer 14 surrounding the core region 12 typically is approximately 125 μm . Typically, the cladding layer 14 is covered, for example, for protection and strength, with one or more coating or buffer layers (not shown), resulting in a total outer diameter of approximately 250-1000 μm .

[0016] The fiber device 10 includes active material formed or otherwise disposed in the cladding layer 14. For example, the cladding layer 14 includes one or more capillaries, voids air-holes or pockets 16 having active or tunable material (shown as 18) infused or otherwise disposed therein. The active material 18 varies the optical properties of optical signals propagating through the optical device 10. For example, the active material 18 is a controllable material whose refractive index can be varied to affect, for example, the signal strength of optical information propagating through the fiber device 10.

[0017] The active material 18 includes, for example, electro-optic material such as polymer dispersed liquid crystals and organic non-linear materials, and/or magneto-optic material such as Europium-based magnetic materials, and/or photorefractive material such as azo compounds or stilbene derivatives, and/or thermo-optic material such as liquid crystals or polymers with sufficiently high dn/dT (where n is refractive index and T is temperature), and/or material that provides tunable gain or loss (for example, laser dyes such as coumarins, substances containing rare-earth elements, and/or composite materials such as polymers dispersed with erbium particles), and/or other suitable material. According to embodiments of the invention, the application of an externally controlled force to the active material(s) results in changes in the optical properties, for example, refractive index, of the active material(s), which, in turn, results in changes in the propagation properties of the optical signal. By establishing spatial asymmetry in the optical device by filling selected pockets with active material(s), embodiments of the invention tunably control the birefringence of the device.

[0018] The pockets or capillaries 16 typically are formed at the preform stage. For example, according to the "stack and draw" method, silica tubes, rods and core rods are bundled into a close-packed arrangement, and the assembly is overclad to obtain the desired dimensions. The preform then is drawn into a fiber in such a way that the capillaries remain open by air pressure inside them, while unwanted holes are kept purged open so surface tension causes them to collapse. Other methods to form the capillaries in the preform are possible, for example, sol-gel methods. Other methods to draw the preforms with capillaries also are available, for example, the "cane" method, in which the preform is drawn into fiber in two high-tension steps; if the draw tension is much higher than the surface tension, hole collapse is prevented.

[0019] The active material 18 is introduced into the capillaries 16 of the device 10 preferably as a liquid. One way of introducing the active material 18 into the capillaries 16 is by immersing one end of the fiber device 10 into the liquid, and applying vacuum on the other end of the fiber device 10. Alternately, the is forced into the capillaries 16 by using positive pressure. Once the active material 18 is inside the capillary, the active material 18 typically is left as a liquid (for example, in the case of liquid crystals) and permanently sealed inside the fiber device 10 by plugging the capillaries 16 with a sealant or collapsing the capillaries 16 with heat.

[0020] Alternatively, the active material 18 is introduced as a liquid comprising polymerizable monomeric units. These monomeric units polymerize inside the capillaries 16 and are activated, for example, by UV or visible light, by heat, or at room temperature as a result of mixing two or more reactive components. Once polymerized, the active material 18 is permanently trapped inside the capillaries 16. Yet another alternative is to introduce a solid active material (such as a polymer) as a liquid solution in a suitable solvent, in such a way that the solid material deposits inside the walls of the capillaries 16. The solvent then is removed, for example, by evaporation, leaving behind a solid layer of the active material 18 deposited in the capillaries 16.

[0021] Alternatively, the active material 18 is introduced into the fiber device 10 before the fiber device 10 is drawn (that is, at the preform stage).

[0022] According to embodiments of the invention, polarization of light is manipulated in optical devices such as the microstructured optical fiber 10 by establishing spatial asymmetry in microstructured optical fiber 10. Spatial asymmetry is established, for example, by selectively filling the pockets 16 with the active material 18 and keeping the remaining pockets 16 filled with air or other suitable non-active material. Such selectively is done, for example, by sealing specific

pockets 16 with, for example, epoxy or other suitable sealant, thus allowing only the open or non-sealed pockets to be filled with the active material 18. According to embodiments of the invention, establishing spatial asymmetry in the optical device 10 introduces birefringence in the fiber structure.

[0023] Often, the pockets 16 in the microstructured fiber device are sufficiently far from the core that they do not affect the core mode. Therefore, to study the birefringence of the polymer-air-silica microstructure, various configurations are used to couple light from the core mode to the cladding region where it can interact with the microstructure. For example, the fiber is tapered to achieve such interaction. Alternatively, a grating such as a fiber Bragg grating (FBG) or a long period grating (LPG) is written in the core to provide the desired interaction between the coupled light and the microstructure. The FBG or LPG couples light from the LP_{01} mode in the core to the co-propagating LP_{02} mode.

[0024] For example, see the tapered fibers shown in Figs. 3a-c and Figs. 4a-c. The fibers are tapered, for example, by heating and stretching the fiber down to approximately 30 microns (μm) with a length of the waist of approximately 2 centimeters (cm). However, despite the reduced size of the tapered region, the refractive index profile of the fiber remains the same as in the non-tapered regions. For example, in the embodiment shown in Figs. 3a and 4a, in the non-tapered regions, the outer diameter is approximately 125 μm , the diameter of the core region is approximately 5-10 μm , and the average diameter of the pockets is approximately 40 μm . In the tapered region, such relative proportions remain intact: the outer diameter is approximately 30 μm , the diameter of the core region is approximately 2 μm , and the average diameter of the pockets 16 is approximately 9 μm .

[0025] Fig. 3a shows a side, partial cross-sectional view of a tapered microstructured fiber having a symmetric mode field (that is, all six holes are filled with active material). Fig. 3b shows an end, cross-sectional view of the tapered fiber, and Fig. 3c shows is a corresponding Poincare sphere for the tapered fiber. When all six air-holes are filled with polymer, the fiber structure retains its axial symmetry.

[0026] Fig. 4a shows a side, partial cross-sectional view of a tapered microstructured fiber having an asymmetric mode field according to embodiments of the invention. In this arrangement, only two holes are filled with active material. However, other asymmetric arrangements are possible according to embodiments of the invention.

[0027] In the Poincare sphere, the polarization state does not change by heating the fiber, as can be seen when comparing Fig. 4c to Fig. 3c. The Poincare sphere is a measure of the polarization state of the transmitted signal and illustrates changes in phase (ϕ between the orthogonal modes, which is also an indication in the change in birefringence from the following relation:

$$\phi = (\beta_x - \beta_y)L = 2\pi(n_x - n_y)L/\lambda = 2\pi BL/\lambda \quad (1)$$

where B = the birefringence, β_x = the propagation constant of the mode in the x direction, β_y = the propagation constant of the mode in the y direction, n_x = the effective index in the x direction, n_y = the effective index in the y direction, L = the length of the fiber waist, and λ = the operating wavelength. The relatively small change in the phase shift detected on the sphere when heated is an indication of low birefringence in the fiber. Such is expected since the effective indices of the orthogonal modes are the same.

[0028] However, when the structure of the fiber is asymmetric, for example, by filling two opposite air-holes, as shown in Fig. 4b, the birefringence is relatively large. The relatively large birefringence is achieved, for example, by heating the fiber and observing the change in the phase on the sphere. Because the effective index of the polymer has relatively large thermal coefficients (dn/dT is approximately $-4 \times 10^{-4}/^\circ\text{C}$), heating the fiber results in relatively dramatic changes in birefringence. The trace of the phase change on the Poincare sphere is a demonstration of the enhanced tunability attained in the birefringence of the fiber. Each rotation around the sphere corresponds to a phase shift of 2π . In Fig. 4c, several rotations are shown on the sphere, which indicates a relatively high tunability attained in the fiber. The hysteresis effect of the device is relatively small as observed, thus the trace of the phase change on the sphere goes back to its original state after the fiber is cooled.

[0029] In Figs. 5a-b, the birefringence is plotted as a function of temperature and refractive index of the active material. Also, the calculated mode profiles in the fiber are shown corresponding to various refractive index values or ambient temperatures. The mode profile is distorted as shown at lower temperatures (and higher refractive index values), thus indicating birefringent effects in the fiber. The effective indices of the orthogonal modes are different along the fast and slow axes. The symmetry of the mode is restored at higher temperatures, as the refractive index value of the active material decreases to the point where the effective refractive index values of the orthogonal modes are approximately the same.

[0030] Referring now to Figs. 6a-b and 7a-b, birefringence structures are characterized, for example, by measuring polarization splitting in the transmission spectra of a long period grating (LPG) written in the core of the structure. The results shown demonstrate the ability to tune and manipulate birefringence in the cladding region and also indicate that a microstructure optical fiber with either six-fold symmetry (Fig. 6a) or two-fold symmetry (Fig. 7a) exhibits reduced

birefringence. The coupling is measured as a wavelength dependent loss peak in transmission determined by the phase matching condition

$$\lambda = \Lambda_{\text{LPG}}(n_{\text{co}} - n_{\text{cl}}) \quad (2)$$

where Λ_{LPG} is the LPG period, n_{co} and n_{cl} are the effective indices of the LP_{01} and LP_{02} modes, respectively. If there is birefringence in either the core or the cladding regions, there will be two resonant wavelengths for the LPG, corresponding to the two principal axes of polarization (denoted x and y below), which have different effective mode indices. The wavelength splitting in the LPG spectra is then related to the birefringence B by

$$\delta\lambda = |\lambda_x - \lambda_y| = \Lambda_{\text{LPG}}(n_{\text{co}} - n_{\text{cl}_x}) - \Lambda_{\text{LPG}}(n_{\text{co}} - n_{\text{cl}_y}) = \Lambda_{\text{LPG}}|n_{\text{cl}_y} - n_{\text{cl}_x}| = \Lambda_{\text{LPG}}B \quad (3)$$

where B is the birefringence of the structure.

[0031] Figs. 6a-b and 7a-b show measured transmission spectra for two different devices, one with all six holes filled with active material such as polymer (Fig. 6a) and the other with only two opposing holes filled with active material (Fig. 7a). The solid and dashed lines in the graphs of Figs. 6b and 7b correspond to separate measurements taken for the two principle axes. At room temperature, the wavelength splitting is relatively small in the device in Fig. 6a (see Fig. 6b) and relatively large in the device in Fig. 7a (see Fig. 7b). From equation (3), it is seen that the six-fold symmetric structure (Fig. 6a) has relatively low birefringence, while the two-fold symmetric structure (Fig. 7a) has high birefringence. It should be noted that the small splitting in the six-fold symmetry case is caused by birefringence in the core induced during the process of writing the grating. Since the pockets are filled with a relatively high dn/dT polymer or other suitable active material, n_{cl} is highly temperature dependent. Thus, a relatively significant shift in the resonance peak of the grating occurs with relatively small increases in temperature.

[0032] In the device shown in Fig. 6a, the shift is the same for both polarization axes. See, Fig. 6b. In the device shown in Fig. 7a, the shift is not the same for both polarization axes, and the wavelength splitting decreases as the device is heated. This is because n_{cl} of the slow axis (that is, when the polarization is aligned with the polymer-filled holes) is highly dependent on n_{pol} , while n_{cl} of the fast axis (that is, when polarization is aligned with the air-filled holes) is not. The splitting decreases because dn_{pol}/dT is negative; heating brings n_{pol} closer to n_{air} , thus restoring symmetry in the cladding. However, it should be noted that heating the device by only 25° Celsius eliminates most of the birefringence, even though n_{pol} at this temperature (approximately 1.424) is still much greater than n_{air} .

[0033] Referring now to Fig. 8, shown is a graphical diagram of the wavelength splitting of the resonance ($\delta\lambda$) as a function of temperature for microstructure optical fiber arrangements in which all pockets are filled with index tunable active material (—), two adjacent pockets are filled with index tunable active material (—Δ—), three alternate pockets are filled with index tunable active material (—○—), and opposite pockets are filled with index tunable active material (—O—). The right axis shows the corresponding birefringence derived from equation (3). For structure arrangements with less than three-fold symmetry, the birefringence is much greater than, for example, the three- and six-fold symmetric structure arrangements.

[0034] Referring now to Figs. 9a-b, shown are simplified schematic diagrams of various system arrangements in which microstructure optical fiber devices according to embodiments of the invention are useful. The system 100 includes one or more sources 110 for transmitting optical information, an optical transmission medium, and one or more receivers 120 for receiving the transmitted information. According to an embodiment of the invention as shown in Fig. 9a, the optical transmission medium includes at least one conventional fiber 130 coupled to the source 110 and at least one microstructured optical fiber device 10 according to embodiments of the invention, for example, as described hereinabove. Also, for example, at least one conventional fiber 130 is coupled between the fiber device 10 and the receiver 120. According to an embodiment of the invention as shown in Fig. 9b, the optical transmission medium includes at least one microstructured optical fiber or fiber device 10 coupled between the transmitter 110 and the receiver 120 and having a tapered region, for example, as discussed hereinabove.

[0035] Referring now to Fig. 10, shown is a simplified block diagram of a method 120 for making optical fiber devices according to embodiments of the invention. The method 120 includes a step 122 of forming an optical fiber device having a core region, a cladding layer, and at least one air pocket in the cladding layer for disposing the active material. Another step 124 of the method 120 is to dispose the active material in the pocket(s) formed in the cladding layer in such a way that establishes spatial asymmetry therein. Alternatively, another step 126 of the method 120 is to taper at least a portion of the optical device.

Claims

1. An optical fiber device (10), comprising:

a length of optical fiber having a core region (12), a cladding layer (14) formed around the core region, and a plurality of pockets (16) formed in the cladding layer radially about the core region,

wherein at least one but not all of the plurality of pockets includes an active material (18) disposed therein to cause the optical fiber device to be spatially asymmetric, and

wherein the active material is capable of changing optical properties in such a way that varies the polarization of light propagating in the optical fiber.

2. The optical fiber device as recited in claim 1, wherein the length of optical fiber has a first diameter, and wherein at least a portion of the length of optical fiber includes a tapered region having a second diameter less than the first diameter.

3. The optical fiber device as recited in claim 2, wherein the core region and the cladding layer are configured in such a way that changing optical properties of the active material affects the propagation properties of optical signals in the tapered region and does not affect the propagation properties of optical signals not in the tapered region.

4. The optical fiber device as recited in claim 1, where the core region includes at least one grating written therein.

5. The optical fiber device as recited in claim 1, wherein the active material further comprises a magneto-optic material, wherein the magneto-optic material has an index of refraction that changes in response to changes in a magnetic field, and wherein at least one of the plurality of pockets is formed in the cladding layer in such a way that changes in the index of refraction of the magneto-optic material vary the polarization of light propagating in the optical fiber.

6. The optical fiber device as recited in claim 1, wherein the active material further comprises a thermo-optic material, wherein the thermo-optic material has an index of refraction that changes in response to changes in temperature, and wherein at least one of the plurality of pockets is formed in the cladding layer in such a way that changes in the index of refraction of the thermo-optic material vary the polarization of light propagating in the optical fiber.

7. The optical fiber device as recited in claim 1, wherein the active material further comprises a photorefractive material, wherein the photorefractive material has an index of refraction that changes in response to exposure to light at a given wavelength, and wherein at least one of the plurality of pockets is formed in the cladding in such a way that changes in the index of refraction of the photorefractive material vary the polarization of light propagating in the optical fiber.

8. The optical fiber device as recited in claim 1, wherein the active material further comprises an electro-optic material, wherein the electro-optic material has an index of refraction that changes in response to changes in electric field, and wherein at least one of the plurality of pockets is formed in the cladding layer in such a way that changes in the index of refraction of the electro-optic material vary the polarization of light propagating in the optical fiber.

9. The optical fiber device as recited in claim 1, wherein the index of refraction of the active material changes from a first value greater than the index of refraction of the core region to a second value less than the index of refraction of the core region.

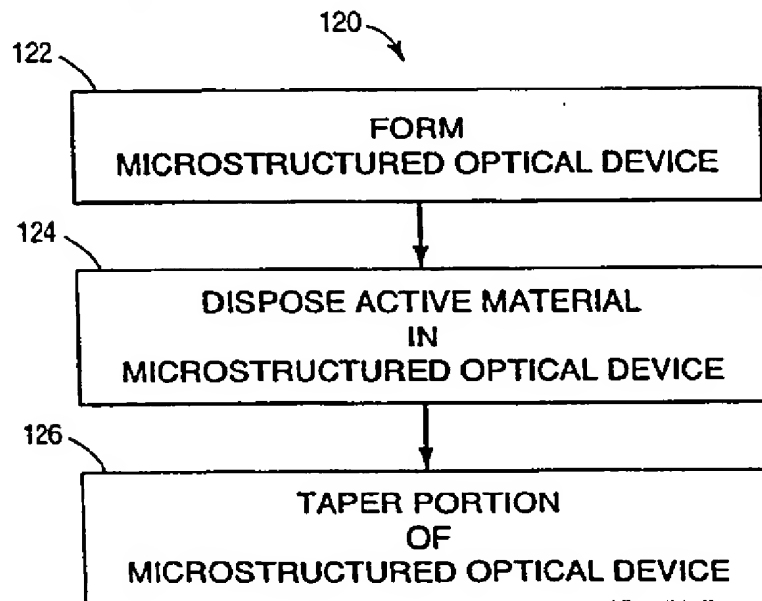
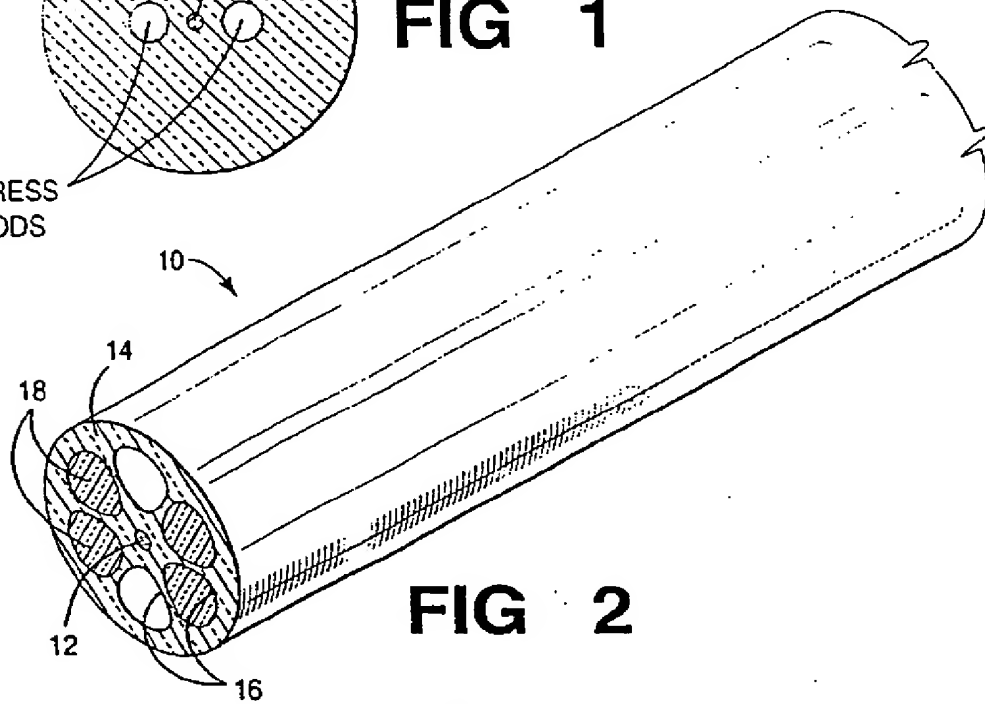
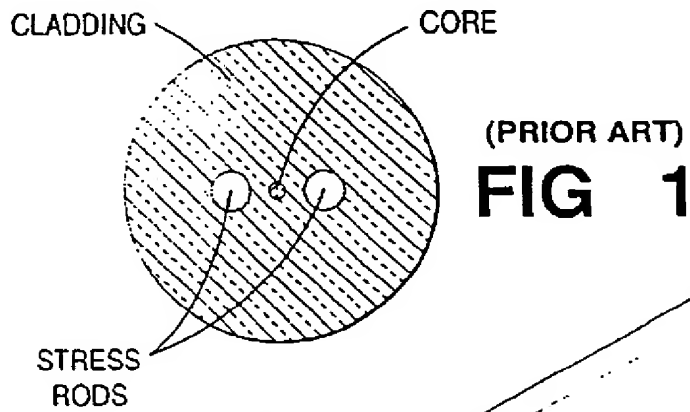


FIG 10

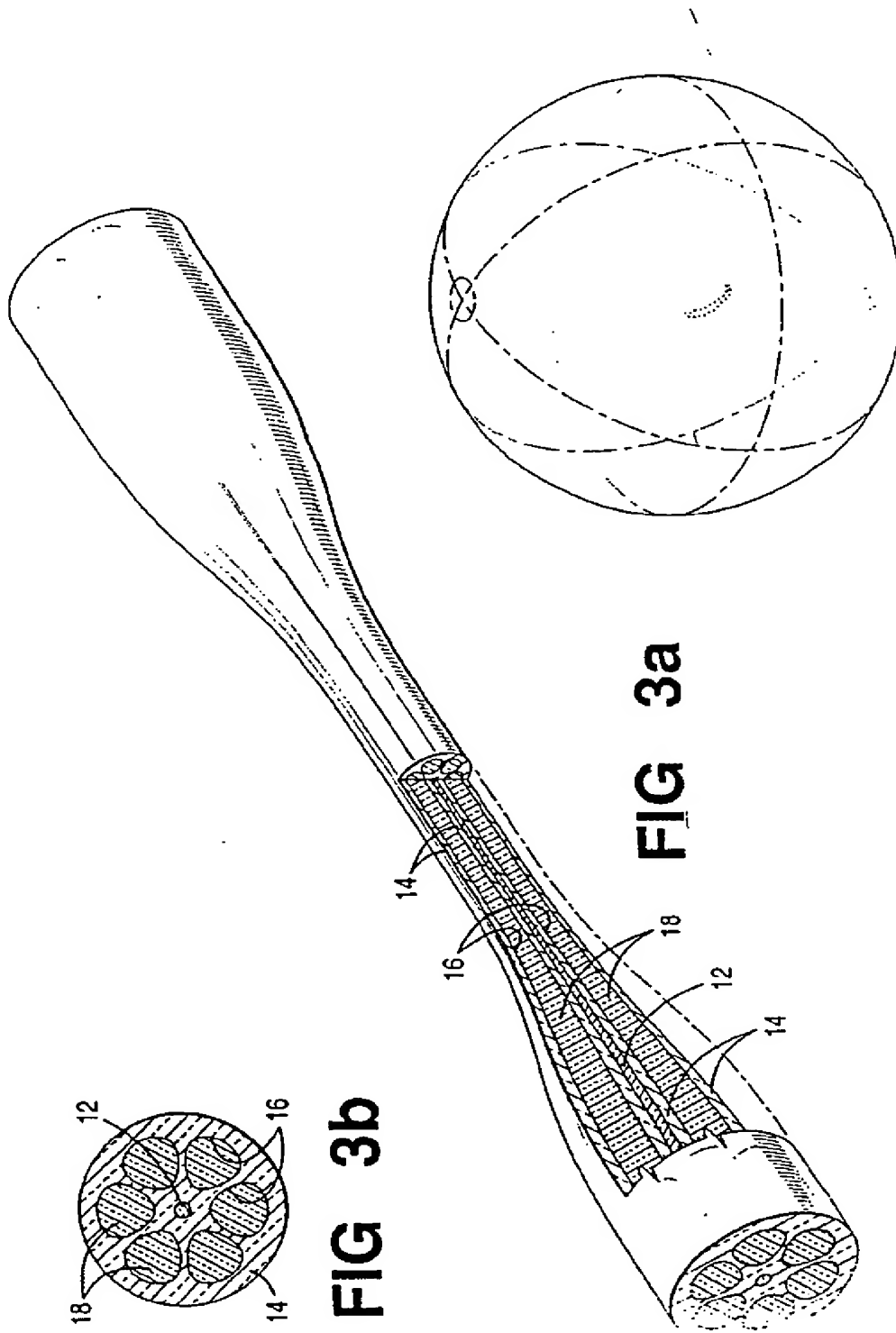


FIG 3b

FIG 3a

FIG 3c

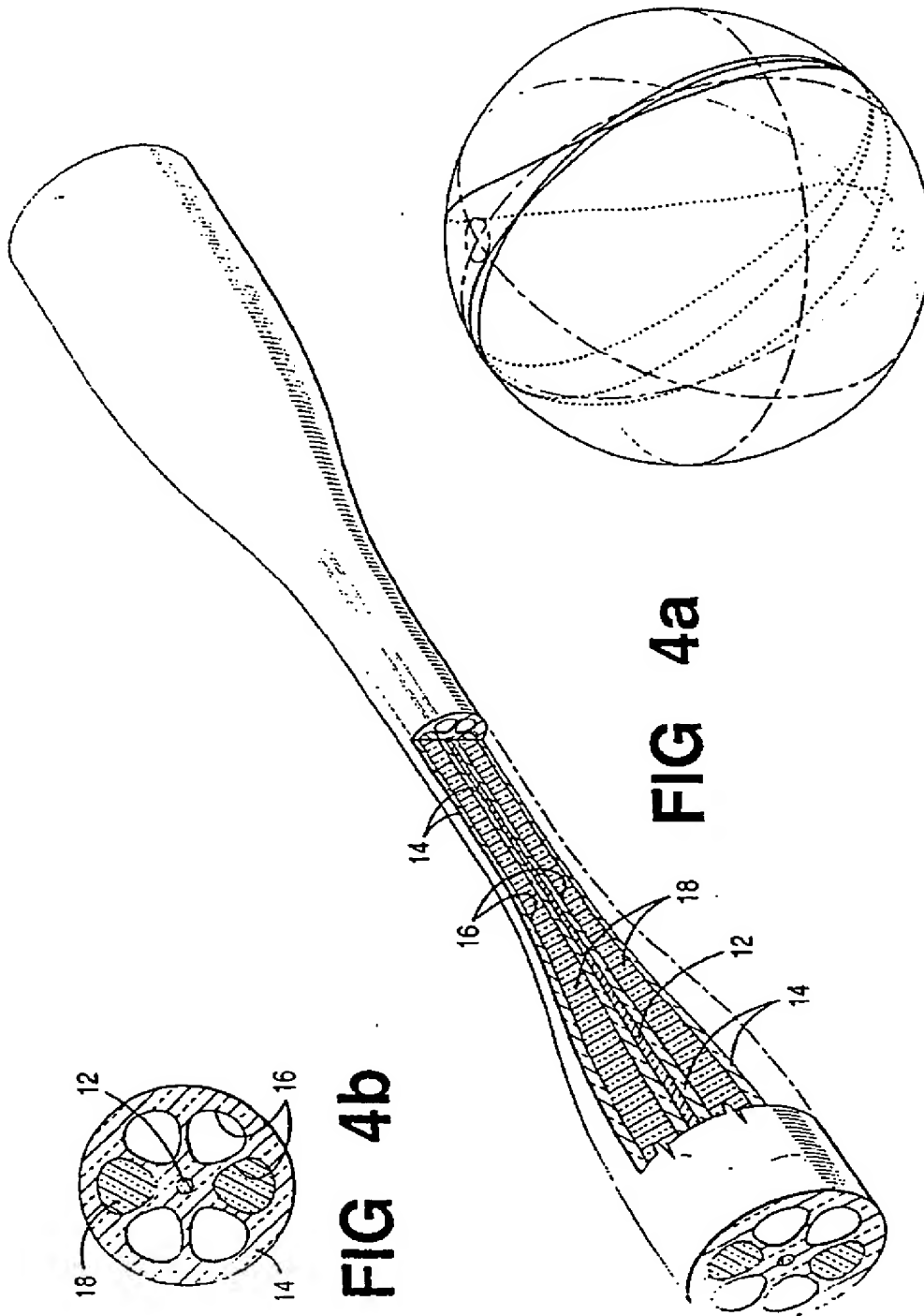


FIG 4b

FIG 4a

FIG 4c

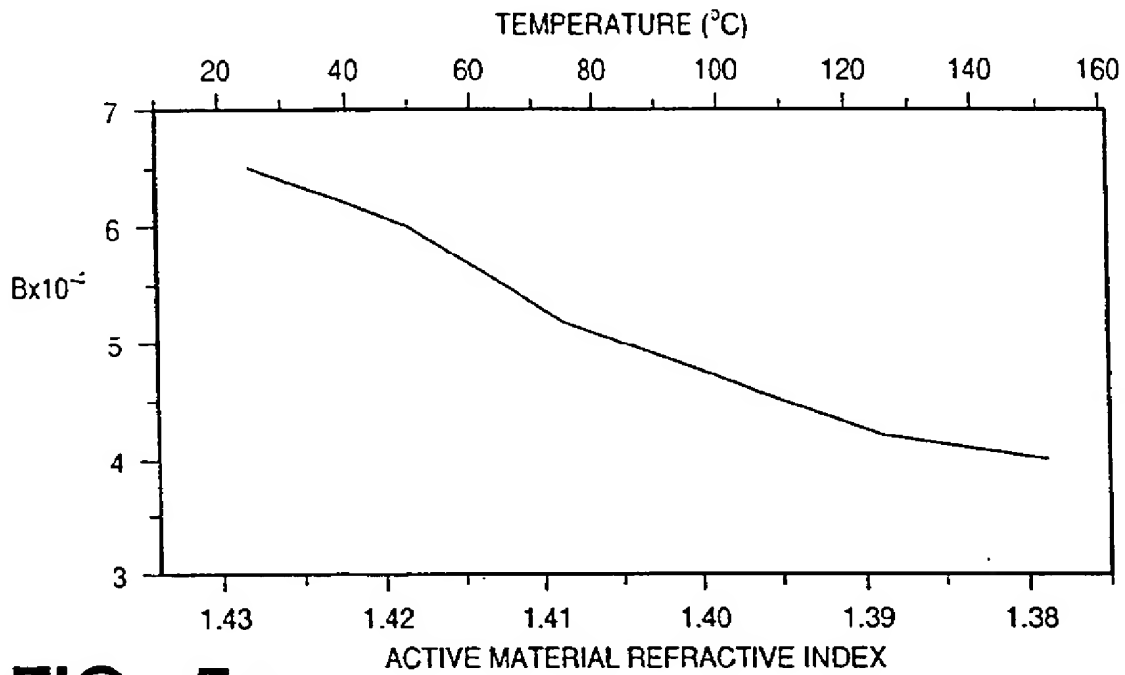


FIG 5a

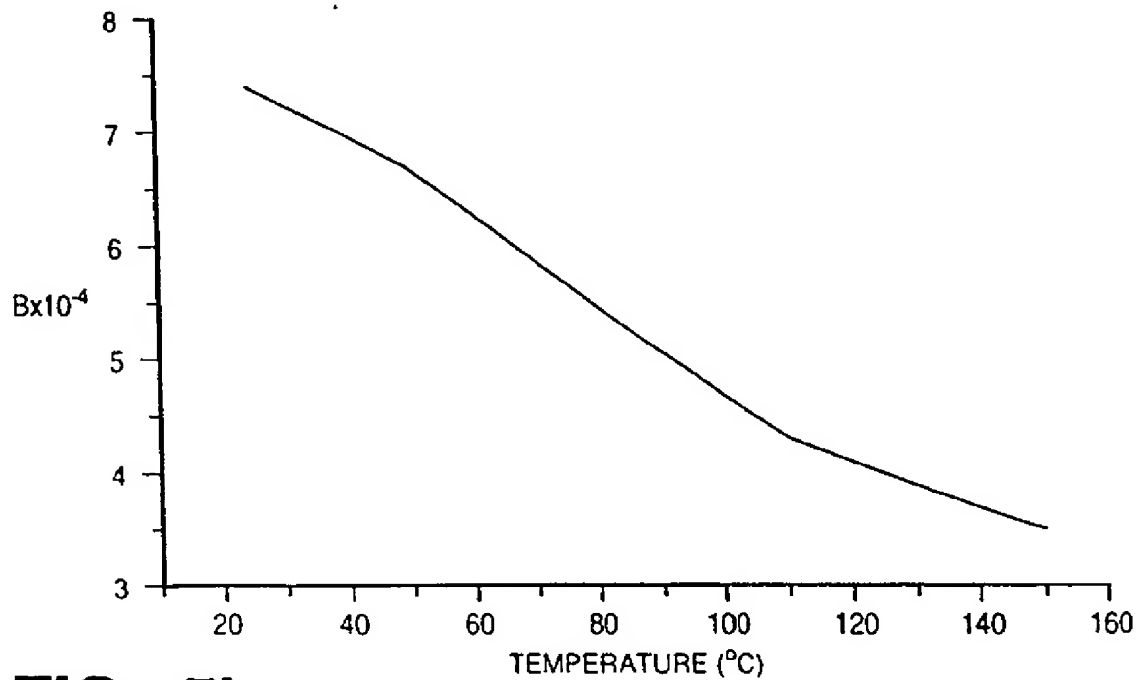
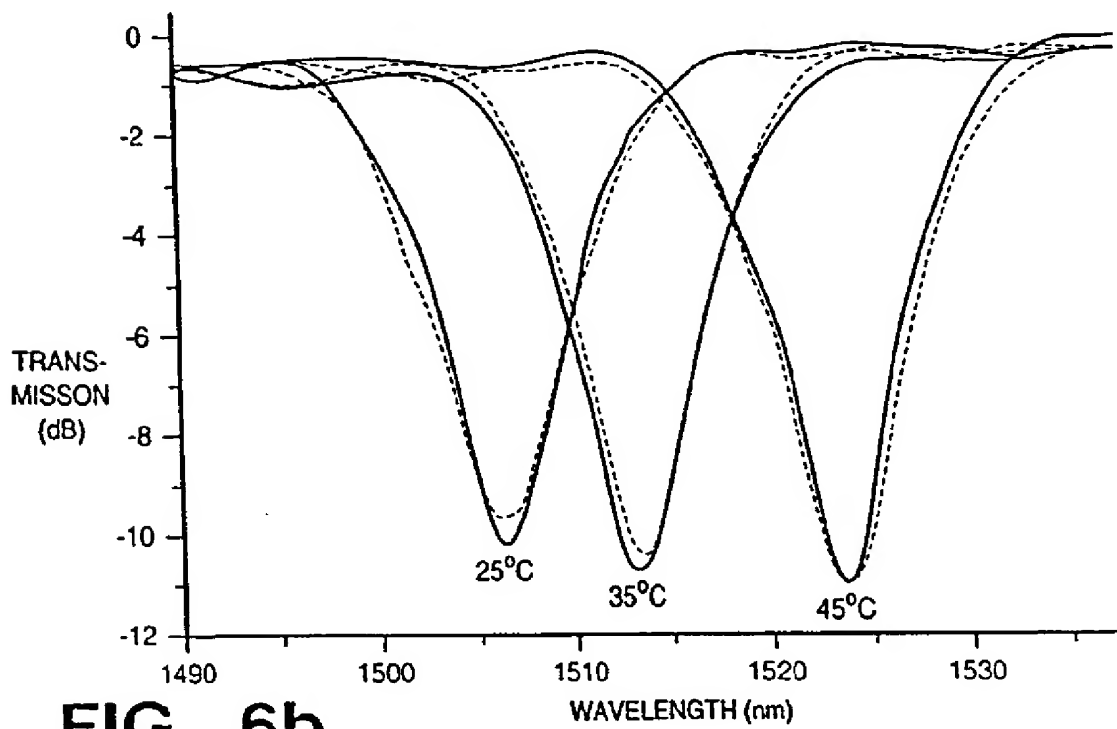
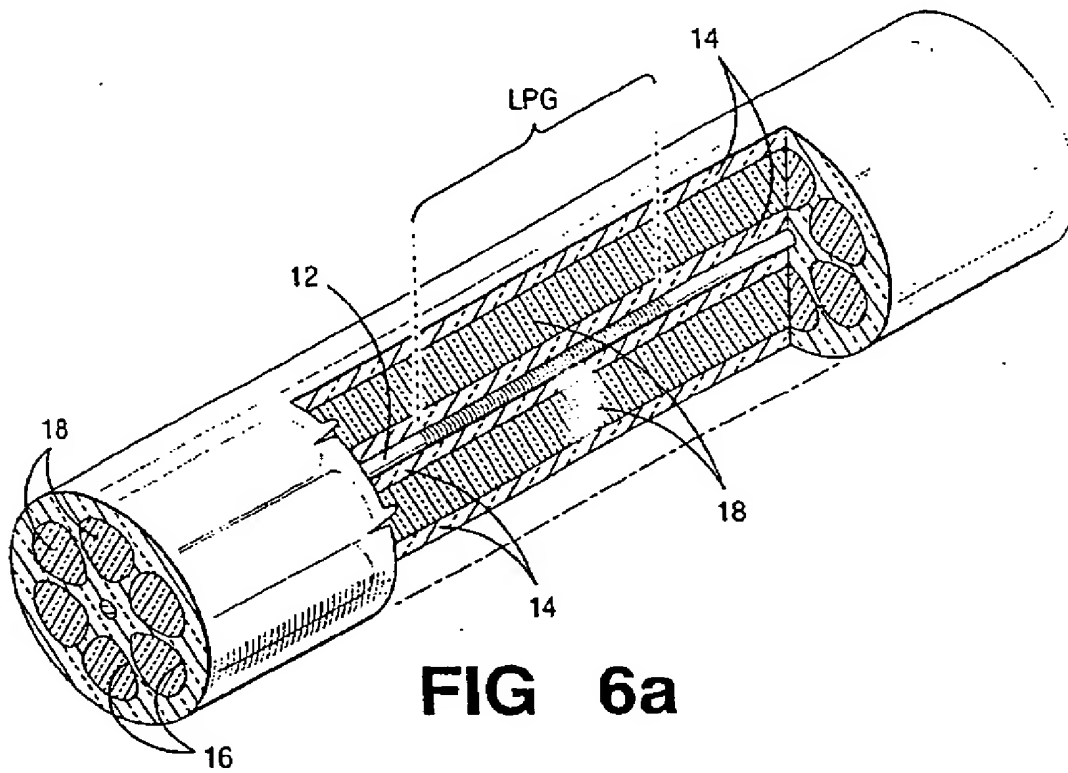
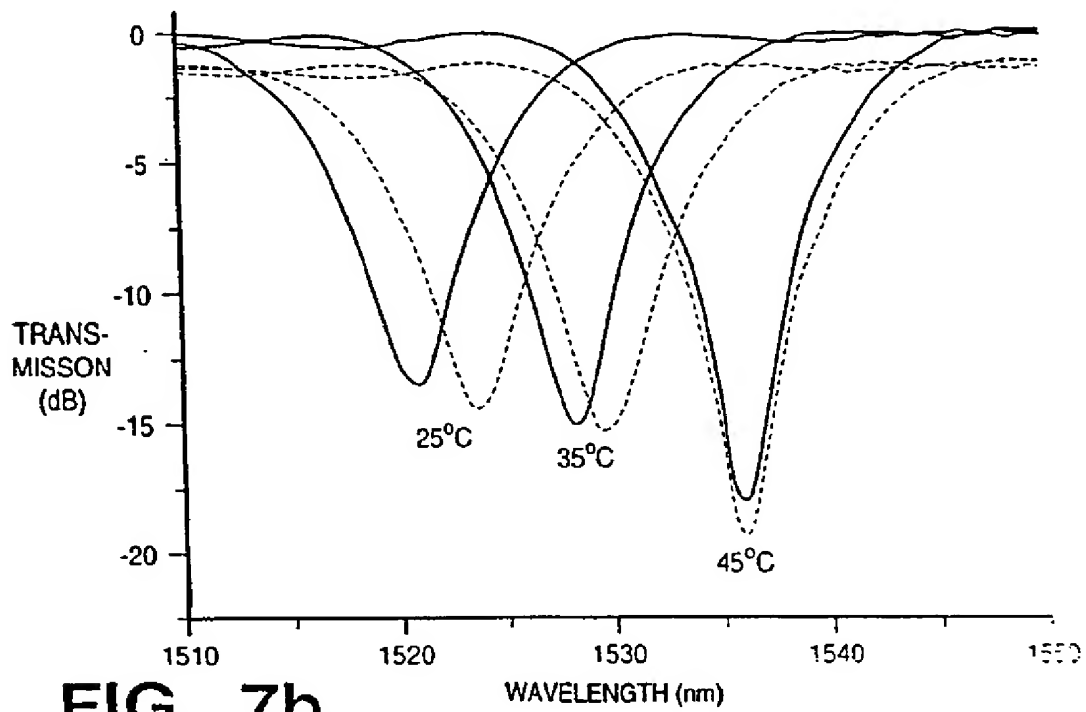
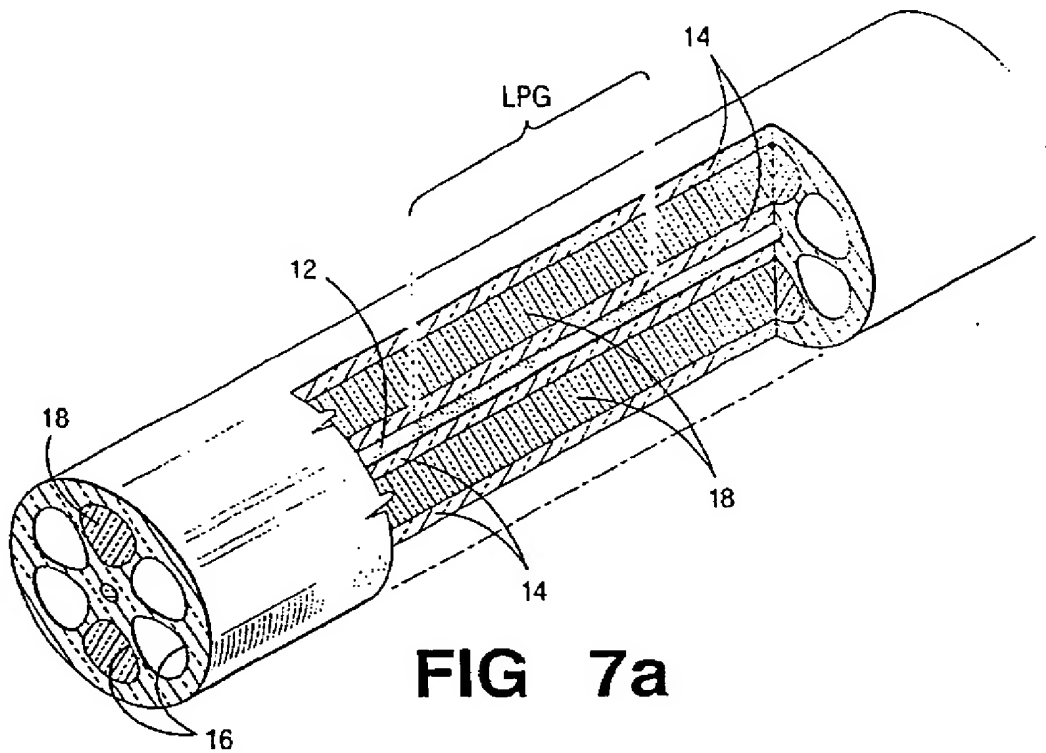
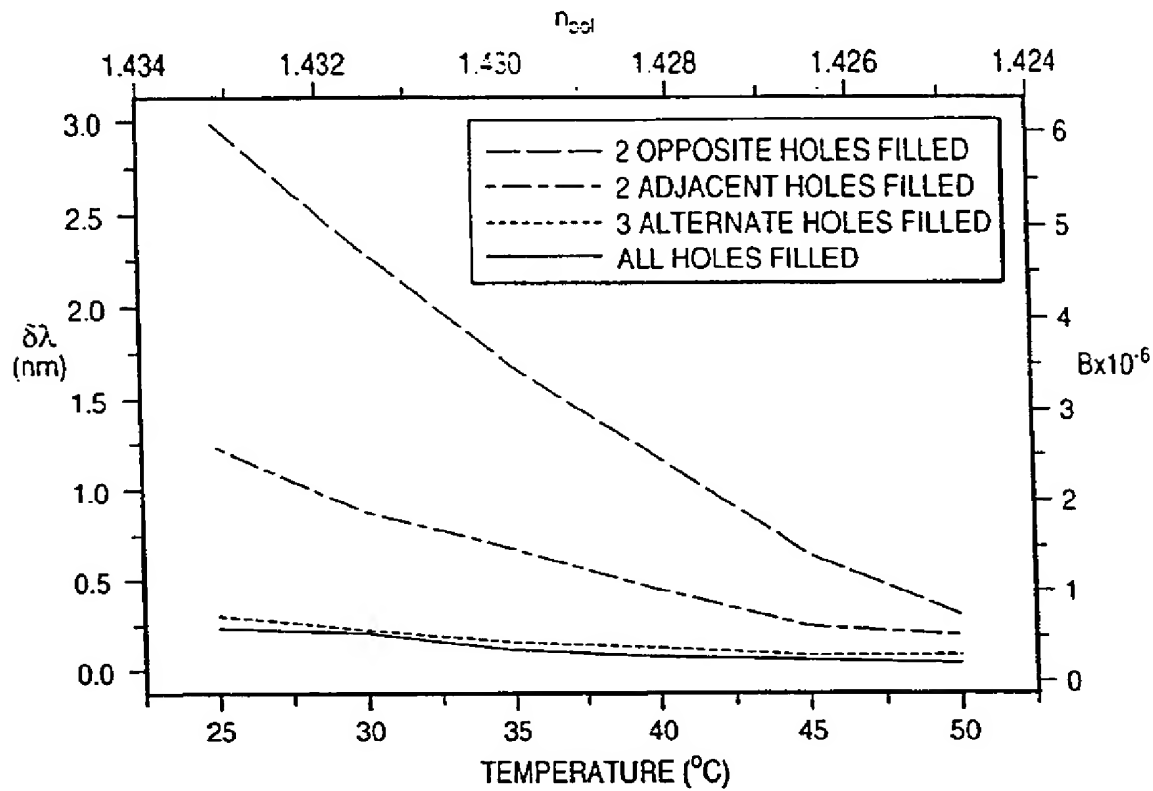
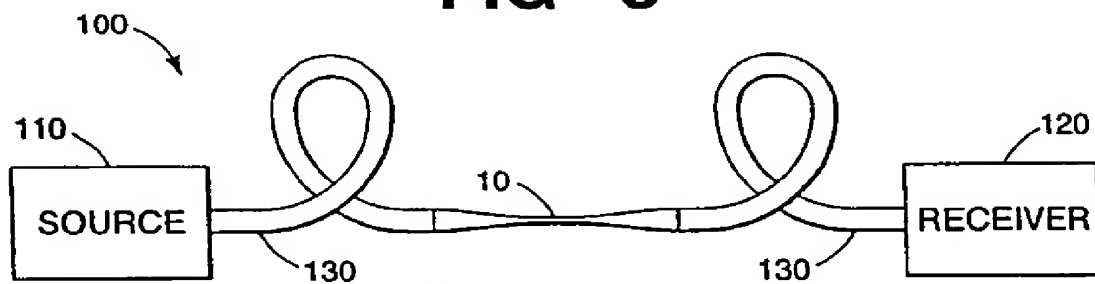
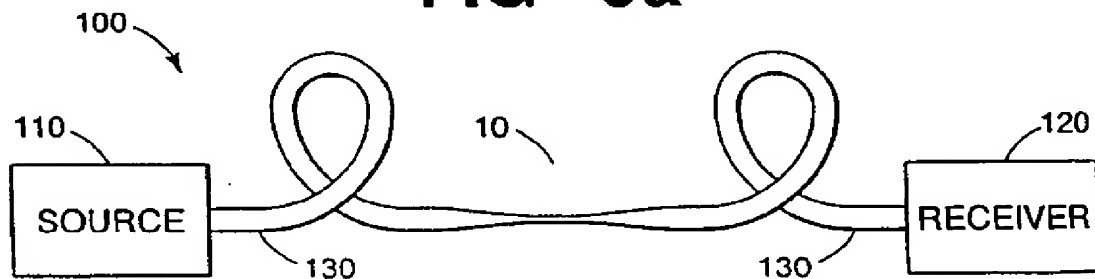
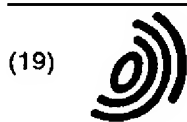


FIG 5b





**FIG 8****FIG 9a****FIG 9b**



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(54) Apparatus and Method of Modifying the Birefringence in Optical Fibres

(57) The invention relates to optical fibre devices and methods of tuning the birefringence of waveguides, especially to all-fibre optical devices which have undergone microstructuring. Embodiments of the invention include an optical fibre device (10) such as a tunable birefringent optical fibre having a core region (12), a cladding layer (14) therearound, and a controllable active material (18) selectively disposed in, for example, capillaries or pockets (16) formed in the cladding layer. The active material/s include, for example, electro-optic material, magneto-optic material, photorefractive material,

thermo-optic material and/or materials such as laser dyes that provide tunable gain or loss. The application of, for example, temperature, light or an electric or magnetic field modifies the optical properties of the active material, which, in turn, modifies the propagation properties of optical signals in the device. The optical device can include a tapered region or long period grating that causes the core mode to spread or couple into the cladding region and, simultaneously, allows the active material to be relatively close to the propagated modes, thus allowing interaction between the active material and the propagating modes.

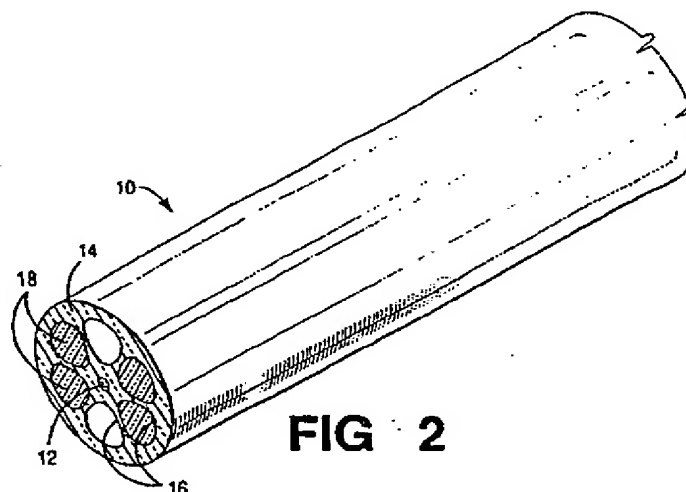


FIG. 2



European Patent
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EUROPEAN SEARCH REPORT

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Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
X	KERBAGE C ET AL: "Numerical analysis and experimental design of tunable birefringence in microstructured optical fiber" OPTICS EXPRESS, 11 MARCH 2002, OPT. SOC. AMERICA, USA, vol. 10, no. 5, XP001154568 ISSN: 1094-4087	1-9	G02F1/01 G02B6/16
X	* figures 1,5 *	1	
X	* page 247, paragraph 4 - page 248, paragraph 2; figure 2 *	9	
X	* figure 2 *	2	
X	* page 247, paragraph 2 *	1,6	
X	* page 247, paragraph 4 - line 1,2 *	1,6	
X	* page 248, paragraph 1 - line 1-6 *	3	
Y	whole document	4,6	
Y	--- WESTBROOK P S ET AL: "CLADDING-MODE RESONANCES IN HYBRID POLYMER-SILICA MICROSTRUCTURED OPTICAL FIBER GRATINGS" IEEE PHOTONICS TECHNOLOGY LETTERS, IEEE INC. NEW YORK, US, vol. 12, no. 5, May 2000 (2000-05), pages 495-497, XP000950309 ISSN: 1041-1135	1	
Y	* page 495, column 1, paragraph 1 - line 5-16; figures 1A,1B *	1,6	
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The present search report has been drawn up for all claims			
Place of search MUNICH		Date of completion of the search 16 September 2003	Examiner Thomas, K
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For more details about this annex : see Official Journal of the European Patent Office, No. 12/82

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For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.

(54) Title: COMPOSITION AND PROCESS FOR TREATING METALS

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(57) Abstract: A chromium free conversion coating at least equivalent in corrosion protective quality to conventional chromate conversion coatings can be formed on metals, particularly cold rolled steel, by a dry-in-place aqueous acidic liquid that preferably has a pH value between 0.5 and 5.0 and comprises: (A) "fluorometallate" anions consisting of (i) at least four fluorine atoms, (ii) at least one atom of an element selected from the group consisting of titanium, zirconium, hafnium, silicon, aluminum, and boron, and, optionally, one or more of (iii) ionizable hydrogen atoms and (iv) oxygen atoms; (B) a component of divalent or tetravalent cations of elements selected from the group consisting of cobalt, magnesium, manganese, zinc, nickel, tin, copper, zirconium, iron, and strontium in such an amount that the ratio of the total number of cations of this component to the number of anions in component (A) is at least about 1:5 but not greater than about 3:1; (C) a component selected from the group consisting of phosphorus-containing inorganic oxyanions and phosphonate anions; and (D) a component of polymers of hydroxy styrene, modified by substitution on the aromatic rings of the polymers of substituted aminomethyl moieties, in which the substituents (other than the carbon atom that is directly bonded to an aromatic ring in the polymer) on the amino nitrogen atom jointly contain at least two carbon atoms and at least one hydroxy moiety but neither of these substituents on the amino nitrogen atom individually contains more than half as many hydroxyl moieties as it has carbon atoms, unless it contains only one carbon atom.

COMPOSITION AND PROCESS FOR TREATING METALS**CROSS-REFERENCE TO RELATED APPLICATIONS**

Not Applicable

STATEMENT REGARDING FEDERALLY SPONSORED RESEARCH OR DEVELOPMENT

Not Applicable

5 BACKGROUND OF THE INVENTION

This invention relates to compositions and processes for treating metal surfaces with acidic aqueous compositions for forming conversion coatings on the metals; the conversion coatings provide excellent bases for subsequent painting. The invention is well suited to treating iron and steel, galvanized iron and steel, zinc and those of its al-
10 loys that contain at least 50 atomic percent zinc, and aluminum and its alloys that contain at least 50 atomic percent aluminum. Preferably the surface treated is predominantly ferrous; most preferably the surface treated is cold rolled steel.

This invention is very closely related to that disclosed in U. S. Patent 5,449,415, from which it differs primarily in using a particularly advantageous type of water soluble
15 and/or dispersible polymer. The object of this invention is to achieve better corrosion resistance under at least one set of corrosion promoting conditions than does the invention illustrated by examples in U. S. Patent 5,449,415, without using any more hexavalent chromium in the process than is used in preferred examples in U. S. Patent 5,449,415.

Except in the claims and the operating examples, or where otherwise expressly
20 indicated, all numerical quantities in this description indicating amounts of material or conditions of reaction and/or use are to be understood as modified by the word "about" in describing the broadest scope of the invention. Practice within the numerical limits stated is generally preferred, however. Also, throughout this description unless expressly stated to the contrary: percent, "parts of", and ratio values are by weight; the term
25 "polymer" includes oligomer; the description of a group or class of materials as suitable or preferred for a given purpose in connection with the invention implies that mixtures of any two or more of the members of the group or class are equally suitable or preferred; description of constituents in chemical terms refers to the constituents at the time of ad-
30 dition to any combination specified in the description, and does not necessarily preclude chemical interactions among the constituents of a mixture once mixed; specification of materials in ionic form implies the presence of sufficient counterions to produce electrical neutrality for the composition as a whole; any counterions thus implicitly specified should preferably be selected from among other constituents explicitly specified in ionic form,

to the extent possible; otherwise such counterions may be freely selected, except for avoiding counterions that act adversely to an object of the invention; and the term "mole" and its variations means "gram-mole" and its variations and may be applied to elemental, ionic, hypothetical, unstable, and any other chemical species defined by number and type of atoms present, as well as to compounds with well defined molecules.

BRIEF SUMMARY OF THE INVENTION

It has been found that improved resistance to corrosion, particularly after subsequent conventional coating with an organic binder containing protective coating such as a paint or lacquer, can be imparted to active metal surfaces, particularly to iron and steel and other ferrous surfaces, by contacting the metal surfaces for a sufficient time at a sufficient temperature with an acidic aqueous composition as described in detail below. Such a composition differs from some preferred embodiments illustrated by example in U. S. 5,449,415 most markedly in the specific chemical nature of the substituents on the polymers of hydroxy styrene, modified by substitution on the aromatic rings of the polymers of substituted aminomethyl moieties, which are characteristic of both this invention and the preferred examples of U. S. 5,449,415. In the latter, at least one of the two substituents, exclusive of the single carbon atom that is bonded both to the amino nitrogen and the aromatic ring, on the amino nitrogen atom of each substituent moiety is a polyhydroxy moiety, but in the present invention both of these substituents jointly contain at least two carbon atoms and at least one hydroxy moiety but neither of these substituents on the amino nitrogen atoms individually contains more than half as many hydroxyl moieties as it has carbon atoms, unless it contains only one carbon atom. Preferably, the composition is coated over the metal surface to be treated and then dried in place on the surface of the metal, without intermediate rinsing.

Various embodiments of the invention include working compositions for direct use in treating metals, concentrates from which such working compositions can be prepared by dilution with water, processes for treating metals with a composition according to the invention, and extended processes including additional steps that are conventional *per se*, such as precleaning, rinsing, and, particularly advantageously, painting or some similar overcoating process that puts into place an organic binder containing protective coating over the conversion coating formed according to a narrower embodiment of the invention. Articles of manufacture including surfaces treated according to a process of the invention are also within the scope of the invention.

DETAILED DESCRIPTION OF THE INVENTION AND PREFERRED EMBODIMENTS

An acidic aqueous composition according to the invention comprises, preferably consists essentially of, or more preferably consists of, water and:

- (A) a component of "fluorometallate" anions, each of said anions consisting of (i) at least four fluorine atoms, (ii) at least one atom of an element selected from the group consisting of titanium, zirconium, hafnium, silicon, aluminum, and boron, and, optionally, one or both of (iii) ionizable hydrogen atoms, and (iv) one or more oxygen atoms; preferably the anions are fluorotitanate (i.e., TiF_6^{-2}) or fluorozirconate (i.e., ZrF_6^{-2}), most preferably fluorotitanate;
- (B) a component of divalent or tetravalent cations of elements selected from the group consisting of cobalt, magnesium, manganese, zinc, nickel, tin, copper, zirconium, iron, and strontium; independently preferably at least, with increasing preference in the order given, 60, 70, 80, 85, 90, 95, or 99 % by weight of the total of component (B) consisting of divalent manganese, cobalt, nickel, or magnesium, more preferably of divalent manganese, cobalt, or nickel; most preferably of divalent manganese;
- (C) a component of phosphorus-containing inorganic oxyanions and/or phosphonate anions; and
- (D) a component of water-soluble and/or water-dispersible polymers of vinyl phenol having at least mono-substituted aminomethyl moieties as substituents on the aromatic rings of said polymers; and, optionally, one or more of the following components:
- (E) a dissolved oxidizing agent, preferably a peroxy compound, more preferably hydrogen peroxide;
- (F) a component selected from the group consisting of tungstate, molybdate, silicotungstate, and silicomolybdate anions; and
- (G) a component selected from dissolved or dispersed complexes stabilized against settling, said complexes resulting from reaction between:
- "fluorometallate" anions, each of said anions consisting of (i) at least four fluorine atoms, (ii) at least one atom of an element selected from the group consisting of titanium, zirconium, hafnium, silicon, aluminum, and boron, and, optionally, one or both of (iii) ionizable hydrogen atoms, and (iv) one or more oxygen atoms and
 - one or more materials selected from the group consisting of metallic and metalloid elements and the oxides, hydroxides, and carbonates of these metallic or metalloid elements
- to produce a reaction product that is not part of any of components (A) through (F) as recited above; preferably this component results from reaction with silica or vanadium(V) oxide.

It should be understood that the components listed need not necessarily all be provided by separate chemicals. For example, "fluorometallate" salts of protonated polymer molecules of component (D) can be used to provide at least part of both of components (A) and (D). Also, if the acidity of the composition is sufficiently high and the substrate that is contacted with it is predominantly ferrous, component (B) can be provided by iron dissolved from the substrate and need not be present in the liquid composition when the liquid composition is first contacted with the substrate.

The pH value of an acidic aqueous liquid composition according to the invention preferably is at least, with increasing preference in the order given, 0.5, 1.0, 1.4, 1.7, 2.0, or 2.3 and independently preferably is not more than, with increasing preference in the order given, 5.0, 4.0, 3.5, 3.2, 2.9, 2.6, or 2.4.

Component (C) as defined above is to be understood as including all of the following inorganic acids and their salts that may be present in the composition: hypophosphorous acid (H_3PO_2), orthophosphorous acid (H_3PO_3), pyrophosphoric acid ($\text{H}_4\text{P}_2\text{O}_7$), orthophosphoric acid (H_3PO_4), tripolyphosphoric acid ($\text{H}_5\text{P}_3\text{O}_{10}$), and further condensed phosphoric acids having the formula $\text{H}_{x+2}\text{P}_x\text{O}_{3x+1}$, where x is a positive integer greater than 3. Component (C) also includes all phosphonic acids and their salts. In a concentrated composition, the concentration in the total composition of phosphorus atoms contained in component (C) is preferably at least, with increasing preference in the order given, 0.05, 0.10, 0.15, 0.20, 0.25, 0.30, 0.34, 0.38, 0.42, or 0.45 moles of phosphorus atoms per kilogram of total composition (this unit of measure being hereinafter freely applied to any other constituent as well as to phosphorus and being hereinafter usually abbreviated as "M/kg") and independently preferably is not more than, with increasing preference in the order given, 2.0, 1.5, 1.0, 0.8, 0.60, 0.55, 0.50, or 0.47 M/kg.

Generally, inorganic phosphates, particularly orthophosphates, phosphites, hypophosphites, and/or pyrophosphates, especially orthophosphates, are preferred for component (C) because they are more economical. Phosphonates are also suitable and may be advantageous for use with very hard water, because the phosphonates are more effective chelating agents for calcium ions. Acids and their salts in which phosphorous has a valence less than five may be less stable than the others to oxidizing agents and are therefore less preferred in compositions according to the invention that are to contain oxidizing agents; such acids and their salts are less preferred in all instances for economy.

The polymers required for component (D) may be prepared by processes as described in one or more of the following U. S. Patents, the entire disclosures of all of which, except for any extent to which they may be inconsistent with any explicit statement

herein, are hereby incorporated herein by reference: 5,891,952, 5,116,912, 4,517,028, 4,433,015, and 4,376,000. Preferably, the polymers required for component (D) are made in a manner that reduces or eliminates the presence of organic solvent in the source of the polymer that is added to a mixture to make an acidic aqueous liquid composition according to the invention. Most preferably, these polymers are made by a method that does not use any organic solvent, as set forth further in the paragraphs following immediately below.

An aqueous solution of substituted aminomethylated polyphenol polymers produced by this preferred process is a product of reaction of:

- (A') at least one precursor phenolic polymer or copolymer, which normally does not bear any substituted aminomethyl substituents on its aromatic rings;
- (B') at least one aldehyde, ketone, or mixture thereof; and
- (C') at least one amine.

The preferred process comprises, preferably consists essentially of, or more preferably consists of, the following operations:

- (I') reacting the precursor phenolic polymer component (A') in water with an organic or inorganic alkalinizing agent to form an aqueous solution of the corresponding phenoxide salt;
- (II') mixing the aqueous solution from operation (I') with the amine component (C'), which preferably consists of one or more secondary amines, and the component (B') of aldehyde, ketone, or mixture thereof to form a single aqueous solution in which chemical reaction among components (A'), (B'), and (C') occurs at a temperature in the range from 20 to 100 °C, preferably from 50 to 80 °C, to attach substituted aminomethyl moieties to at least some of the aromatic rings in the precursor polymer and produce an aqueous solution of substituted aminomethylated phenolic polymer molecules;
- (III') adding at least one acid to the aqueous solution formed at the end of operation (II'), the quantity of acid added being sufficient to neutralize the alkalinizing agent added in operation (I') and to protonate a sufficient fraction of the amino nitrogen atoms in the substituted aminomethylated phenolic polymer to stabilize against settling the solution of the substituted aminomethylated phenolic polymer formed in operation (II'); and
- (IV') contacting the resulting aqueous solution from the end of operation (III') with a cation exchange resin in its protonated form to remove at least, with increasing preference in the order given, 50, 75, 90, 95, 99.0, 99.50, 99.70, 99.90, 99.95, or 99.98 % of any inorganic and/or quaternary ammonium cations dissolved in

said aqueous solution from the end of operation (III').

Ordinarily, as is known in the art, contacting is most conveniently, economically, and effectively, and therefore preferably, carried out by passing the aqueous solution through a bed of ion-exchange resin beads arranged in a vertical column of sufficient
5 length that the desired level of removal of alkali metal and quaternary ammonium cations is achieved by the time the solution has passed through the entire column, and the ion-exchange resin in the column can later be returned to its protonated form by treatment with strong acid; however, many other methods of establishing contact between the solution and the ion-exchange resin for a sufficient time to remove the unwanted cations
10 from the aqueous solution are known to those skilled in ion-exchange, and any of these methods may be used.

If it is desired, as is usually preferred, to remove unreacted amine as well as inorganic cations added as part of the alkalizing agent, a strong acid cation exchange resin is used in operation (IV'). If it is desired to remove only inorganic cations, a weak
15 acid cation exchange resin may be used instead. Suitable strong acid cation exchange resins are those of the sulfonic acid or phosphonic acid types, and suitable weak acid cation exchange resins are those of the carboxylic acid type.

The quantities of components (A'), (B') and (C') used to prepare the substituted aminomethylated polyphenol polymer product in aqueous solution preferably are such
20 as to have the following ratios to one another, independently for each ratio specified:

- the number of moles of carbonyl groups in component (B') has a ratio to the number of moles of primary and secondary amino nitrogen atoms in component (C') that is at least, with increasing preference in the order given, 0.5:1.00, 0.7:1.00, 0.80:1.00, 0.85:1.00, 0.90:1.00, 0.95:1.00, or 0.99:1.00 and independently preferably is not more than, with increasing preference in the order given,
25 1.5:1.00, 1.3:1.00, 1.20:1.00, 1.15:1.00, 1.10:1.00, 1.05:1.00, or 1.01:1.00;
- the number of moles of carbonyl groups in component (B') has a ratio to the number of moles of aromatic rings in component (A') that is at least, with increasing preference in the order given, 0.10:1.00, 0.20:1.00, 0.30:1.00, 0.40:1.00,
30 0.50:1.00, 0.60:1.00, 0.70:1.00, 0.80:1.00, 0.85:1.00, 0.90:1.00, or 0.94:1.00 and independently preferably is not more than, with increasing preference in the order given, 2.00:1.00, 1.90:1.00, 1.80:1.00, 1.70:1.00, 1.60:1.00, 1.50:1.00, 1.40:1.00, 1.30:1.00, 1.20:1.00, 1.15:1.00, 1.10:1.00, 1.05:1.00, 1.00:1.00, or 0.96:1.00; and
- the number of moles of primary and secondary amino nitrogen atoms in component (C') has a ratio to the number of moles of aromatic rings in component (A')
35 that is at least, with increasing preference in the order given, 0.10:1.00,

0.20:1.00, 0.30:1.00, 0.40:1.00, 0.50:1.00, 0.60:1.00, 0.70:1.00, 0.80:1.00, 0.85:1.00, 0.90:1.00, or 0.94:1.00 and independently preferably is not more than, with increasing preference in the order given, 2.00:1.00, 1.90:1.00, 1.80:1.00, 1.70:1.00, 1.60:1.00, 1.50:1.00, 1.40:1.00, 1.30:1.00, 1.20:1.00, 1.15:1.00, 1.10:1.00, 1.05:1.00, 1.00:1.00, or 0.96:1.00.

Additionally and independently, the quantities of components (A'), (B'), and (C') preferably are such as to provide an aqueous solution at the end of operation (IV') that contains at least, with increasing preference in the order given, 1.0, 3.0, 5.0, 6.0, 7.0, 7.5, 8.0, 8.5, 9.0, 9.5, or 10.0 percent of the total mass of the aqueous solution as polymer molecules that satisfy the definition for component (A) as given above. To avoid impractically high viscosity of the solution, such polymer molecules independently preferably constitute not more than, with increasing preference in the order given, 50, 35, or 30 % of the total mass of the aqueous solution.

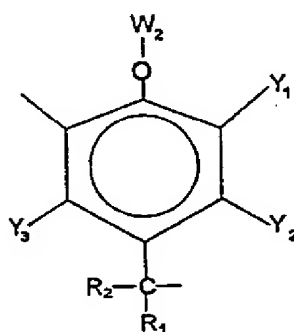
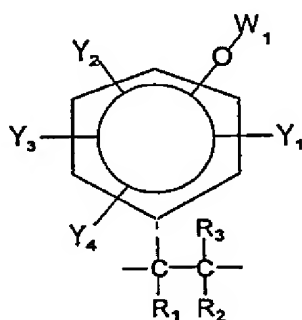
In operation (I') the alkalinizing agent is preferably, for economy, an alkali metal hydroxide, e.g. sodium or potassium hydroxide, although tetraalkylammonium hydroxides, e.g. tetrabutylammonium hydroxide, or tetraarylammonium hydroxides can also be used with technically satisfactory results. A sufficient number of moles of alkalinizing agent should be present to correspond to at least, with increasing preference in the order given, 10, 15, 20, or 25 % of the number of moles of phenolic hydrogen atoms in component (A').

Operation (I') is preferably carried out at a temperature that is at least, with increasing preference in the order given, 30, 40, 50, 55, 60, or 65 °C, in order to minimize the amount of neutralizing agent that is needed in operation (I') by increasing the solubility, or at least the speed of dissolution, in water of at least partially neutralized polymer molecules formed in this step. Primarily for operating convenience when using process equipment that is not easily pressurized, the temperature during operation (I') independently preferably is not more than, with increasing preference in the order given, 95, 90, 80, or 70 °C.

In operation (III') the acid used to neutralize the alkalinizing agent can be organic or inorganic. Suitable acids for this purpose include carbonic acid, acetic acid, citric acid, oxalic acid, ascorbic acid, phenylphosphonic acid, chloromethylphosphonic acid; mono-, di- and tri-chloroacetic acids, trifluoroacetic acid, nitric acid, phosphoric acid, hydrofluoric acid, tetrafluoroboric acid, hexafluorotitanic acid, hexafluorosilicic acid, hexafluorozirconic acid, sulfuric acid, boric acid, hydrochloric acid, and the like. The most preferred acid is a mixture of at least one of the "fluorometallic" acids with phosphoric acid. This mixture is preferred at least for economy, because when these

two acids are used as neutralizers in the preparation of an aqueous solution and/or dispersion of necessary component (D), this solution and/or dispersion provides at least part of both components (A) and (C), which are also needed for a composition according to this invention.

- 5 Component (A'), i.e., the precursor phenolic polymer or copolymer, is preferably selected from polymer molecules in which at least, with increasing preference in the order given, 10, 20, 30, 40, 50, 60, 70, 80, or 90 % of the mass of the polymer molecules is constituted of moieties that conform to one of the two following general formulas, or would so conform if one of the open bonds shown in the following general formulas were replaced by a bond to a hydrogen atom:
- 10



wherein, in either or both of these general formulas when present:

- each of R_1 through R_3 is independently selected from the group consisting of a hydrogen atom, alkyl moieties having from 1 to 5 carbon atoms, and aryl moieties having from 6 to 18 carbon atoms;
 - each of Y_1 through Y_4 is independently selected from the group consisting of a
- 15

hydrogen atom and alkyl and aryl moieties having from 1 to 18 carbon atoms;
and

- W_1 or W_2 is selected from the group consisting of a hydrogen atom and hydrolyzable moieties, preferably an acyl group, e.g. acetyl, benzoyl, and the like.

5 Most preferably, independently for each such element of the formula, each of R_1 through R_3 , Y_1 through Y_4 , W_1 , and W_2 , is a hydrogen atom.

The weight average molecular weight of component (A') preferably is at least, with increasing preference in the order given, 500, 1000, 1500, 2000, 2400, 2800, 3200, 3600, 4000, 4300, 4600, 4800, or 4900 Daltons and independently preferably is not more
10 than, with increasing preference in the order given, 30,000, 25,000, 20,000, 15,000, 10,000, 8000, 6000, or 5100 Daltons.

Component (B') is preferably an aldehyde, and most preferably is formaldehyde, especially in the form of paraformaldehyde. Liquid formaldehyde is generally commercially available only in a form that contains a significant quantity of methanol, e.g. 15 %
15 methanol, as a polymerization inhibitor for the formaldehyde. Since the preferred process is carried out in the absence of organic solvents, formaldehyde free from methanol, such as uninhibited aqueous formaldehyde or paraformaldehyde, is preferably used.

Component (C') is selected from amine molecules, preferably exclusively from
20 secondary amines, still more preferably from secondary amines in which the total number of carbon atoms is not more than, with increasing preference in the order given, 10, 8, 6, 4, or 3. Examples of suitable secondary amines are methyl methanol amine, methyl ethanol amine, methyl butanol amine, ethyl methanol amine, ethyl ethanol amine, pentyl ethanol amine, pentyl pentanol amine, hexyl ethanol amine, dimethanol amine, diethanol
25 amine, dipropanol amine, dibutanol amine, dipentanol amine, nonyl methanol amine, octyl ethanol amine, and the like. Primary amines, such as C_1 - C_{12} alkyl and alkanol amines and the like, can also be used. Most preferably, component (C') is selected from secondary amines in which one of the organic moieties bonded directly to the amino
30 nitrogen atom in the amine is an unsubstituted alkyl moiety and the other is a hydroxyl substituted but otherwise unsubstituted alkyl moiety. The single most preferred substance for component (C') is methyl ethanol amine with the formula $H_3C-NH-(CH_2)_2OH$.

Even if component (D) of a composition according to this invention is in fact made by some other method than the preferred method described in the immediately preceding paragraphs, its chemical characteristics preferably are, with the same degree of prefer-
35 ence, those that would result from being made according to this preferred method with preferred choices as indicated above.

The term "stabilize(d) against settling" in the description above of component (G) and of operation (III') means that the composition containing the material does not suffer any visually detectable settling or separation into distinct liquid phases when stored, without mechanical agitation, for a period of 100, or more preferably 1000, hours at 25 °C.

5 Materials for component (G) may be prepared by mixing the two types of reagents noted in the definition for component (G). A spontaneous chemical reaction normally ensues, converting the added element, oxide, hydroxide, or carbonate into a soluble species. The reaction to form this soluble species can be accelerated by use of heat and by stirring or other agitation of the composition. The formation of the soluble
10 species is also aided by the presence in the composition of suitable complexing ligands, such as peroxide and fluoride.

For a variety of reasons, it is preferred that compositions according to the invention as defined above should be substantially free from many ingredients used in compositions for similar purposes in the prior art. Specifically, it is increasingly preferred in
15 the order given, independently for each preferably minimized component listed below, that these compositions, when directly contacted with metal in a process according to this invention, contain no more than 1.0, 0.35, 0.10, 0.08, 0.04, 0.02, 0.01, 0.001, or 0.0002, percent of each of the following constituents: organic materials that are liquid at 25 °C under normal atmospheric pressure and have a vapor pressure of at least 0.05 bar
20 at 25 °C; hexavalent chromium; ferricyanide; ferrocyanide; sulfates and sulfuric acid; alkali metal and ammonium cations; pyrazole compounds; sugars; gluconic acid and its salts; glycerine; α -glucoheptanoic acid and its salts; and myoinositol phosphate esters and salts thereof.

Furthermore, in a process according to the invention that includes other operations than the drying in place on the surface of the metal of a layer of a composition as
25 described above, it is preferred that none of these other operations include contacting the surfaces with any composition that contains more than, with increasing preference in the order given, 1.0, 0.35, 0.10, 0.08, 0.04, 0.02, 0.01, 0.003, 0.001, or 0.0002 % of hexavalent chromium, except that a final protective coating system including an organic
30 binder, more particularly those including a primer coat, may include hexavalent chromium as a constituent. Any such hexavalent chromium in the protective coating is generally adequately confined by the organic binder, so as to avoid adverse environmental impact.

In one embodiment of the invention, it is preferred that the acidic aqueous composition as noted above be applied to the metal surface and dried thereon within a short
35 time interval. With increasing preference in the order given, the time interval during which the liquid coating is applied to the metal being treated and dried in place thereon,

when heat is used to accelerate the process, is not more than 25, 15, 9, 7, 4, 3, 1.8, 1.0, or 0.7 second (hereinafter often abbreviated "sec"). In order to facilitate this rapid completion of a process according to this invention, it is preferred to apply the acid aqueous composition used in the invention to a heated metal surface, such as one rinsed with hot water and subsequently dried after initial cleaning and very shortly before treating with the aqueous composition according to this invention, and/or to use infrared or microwave radiant heating and/or convection heating in order to effect very fast drying of the applied coating. Preheating of the metal substrate before application of a composition according to the invention is preferred over postheating of the applied liquid composition when practical, because the latter is more likely to result in unwanted deformation of the coating film or inhomogeneous properties of the film as a result of more rapid drying in some areas than in others. Whether preheating, postheating, or both are used, the peak metal temperature preferably is in a range from, with increasing preference in the order given, 10 - 100, 15 - 95, 20 - 90, 20 - 80, or 20 - 70, °C.

In an alternative embodiment, which is equally effective technically and is satisfactory when ample time is available at acceptable economic cost, a composition according to this invention may be applied to the metal substrate and allowed to dry at a temperature not exceeding 40 °C. In such a case, there is no particular advantage to fast drying. This alternative embodiment is particularly advantageously used for "touching up" thinned or damaged coatings that are already in place over most of the surface of some article that is too large to fit into any conveniently available oven or other heating device. For such uses a composition according to the invention is advantageously applied to the substrate to be treated with the aid of an applicator as taught in U. S. Patent 5,702,759 of Dec. 30, 1997 to White et al., the entire disclosure of which, except to any extent that may be inconsistent with any explicit statement herein, is hereby incorporated herein by reference.

The effectiveness of a treatment according to the invention appears to depend predominantly on the total amounts of the active ingredients that are dried in place on each unit area of the treated surface, and on the nature and ratios of the active ingredients to one another, rather than on the concentration of the acidic aqueous composition used. Thus, if the surface to be coated is a continuous flat sheet or coil and precisely controllable coating techniques such as roll coaters are used, a relatively small volume per unit area of a concentrated composition may effectively be used for direct application. On the other hand, with some coating equipment, it is equally effective to use a more dilute acidic aqueous composition to apply a heavier liquid coating that contains about the same amount of active ingredients. As a general guide, it is normally

preferable, independently for each preference stated, for a working composition according to the invention that is intended to be dried without substantial external heating and/or to be applied without precise control of the total amount applied to have:

- 5 - a total concentration of "fluorometallate" anions of component (A) that is at least, with increasing preference in the order given, 0.005, 0.010, 0.020, 0.030, 0.035, 0.040, 0.050, or 0.055 M/kg;
- a total concentration of metal cations of component (B) that is at least, with increasing preference in the order given, 0.007, 0.011, 0.020, 0.030, 0.035, 0.040, 0.045, 0.050, 0.054, 0.058, or 0.062 M/kg;
- 10 - a ratio of total concentration in M/kg of metal cations of component (B) in M/kg to the total concentration in M/kg of "fluorometallate" anions of component (A) in the same composition that is at least, with increasing preference in the order given, 0.2:1.0, 0.4:1.0, 0.60:1.00, 0.70:1.00, 0.80:1.00, 0.90:1.00, 0.95:1.00, 1.00:1.00, 1.05:1.00, 1.10:1.00, or 1.12:1.00 and independently preferably is not
15 more than, with increasing preference in the order given, 3.0:1.00, 2.6:1.00, 2.2:1.00, 1.8:1.00, 1.6:1.00, 1.40:1.00, 1.30:1.00, 1.20:1.00, or 1.13:1.00;
- a concentration of phosphorus from component (C) that is at least, with increasing preference in the order given, 0.007, 0.012, 0.017, 0.022, 0.027, 0.032, 0.037, 0.042, 0.047, 0.052, 0.057, 0.062, or 0.067 M/kg;
- 20 - a ratio of moles of phosphorus from component (C) to moles of "fluorometallate" ions from component (A) in the same composition that is at least, with increasing preference in the order given, 0.2:1.0, 0.4:1.0, 0.60:1.00, 0.70:1.00, 0.80:1.00, 0.90:1.00, 0.95:1.00, 1.00:1.00, 1.05:1.00, 1.10:1.00, 1.15:1.00, or 1.19:1.00 and
25 independently preferably is not more than, with increasing preference in the order given, 3.0:1.00, 2.6:1.00, 2.2:1.00, 1.8:1.00, 1.6:1.00, 1.40:1.00, 1.30:1.00, 1.25:1.00, or 1.20:1.00;
- a ratio of moles of phosphorus from component (C) to moles of total metal cations from component (B) in the same composition that is at least, with increasing preference in the order given, 0.2:1.0, 0.4:1.0, 0.60:1.00, 0.70:1.00, 0.80:1.00, 0.90:1.00, 0.95:1.00, 1.00:1.00, or 1.05:1.00 and independently preferably
30 is not more than, with increasing preference in the order given, 2.6:1.00, 2.2:1.00, 1.8:1.00, 1.6:1.00, 1.40:1.00, 1.30:1.00, 1.25:1.00, 1.20:1.00, 1.15:1.00, 1.11:1.00, or 1.07:1.00;
- a concentration of moles of substituted phenol moieties from component (D) that
35 is at least, with increasing preference in the order given, 0.004, 0.008, 0.012, 0.018, 0.024, 0.028, or 0.031 M/kg;

- a ratio of moles of substituted phenol moieties from component (D) to moles of total "fluorometallate" anions from component (A) in the same composition that is at least, with increasing preference in the order given, 0.060:1.00, 0.12:1.00, 0.16:1.00, 0.20:1.00, 0.25:1.00, 0.30:1.00, 0.35:1.00, 0.40:1.00, 0.43:1.00, 0.46:1.00, 0.49:1.00, or 0.54:1.00 and independently preferably is not more than, with increasing preference in the order given, 2.0:1.00, 1.5:1.00, 1.0:1.00, 0.80:1.00, 0.70:1.00, 0.65:1.00, 0.61:1.00, or 0.57:1.00;
- a ratio of moles of substituted phenol moieties from component (D) to moles of total metal cations from component (B) in the same composition that is at least, with increasing preference in the order given, 0.060:1.00, 0.12:1.00, 0.16:1.00, 0.20:1.00, 0.25:1.00, 0.30:1.00, 0.35:1.00, 0.40:1.00, 0.43:1.00, 0.46:1.00, or 0.49:1.00 and independently preferably is not more than, with increasing preference in the order given, 2.0:1.00, 1.5:1.00, 1.0:1.00, 0.80:1.00, 0.70:1.00, 0.65:1.00, 0.61:1.00, 0.57:1.00, 0.54:1.00 or 0.51:1.00;
- a ratio of moles of substituted phenol moieties from component (D) to moles of phosphorus from component (C) in the same composition that is at least, with increasing preference in the order given, 0.060:1.00, 0.12:1.00, 0.16:1.00, 0.20:1.00, 0.25:1.00, 0.30:1.00, 0.35:1.00, 0.40:1.00, 0.43:1.00, or 0.46:1.00 and independently preferably is not more than, with increasing preference in the order given, 2.0:1.00, 1.5:1.00, 1.0:1.00, 0.80:1.00, 0.70:1.00, 0.65:1.00, 0.61:1.00, 0.57:1.00, 0.54:1.00, 0.51:1.00, or 0.48:1.00.

Working compositions containing up to from five to ten times these amounts of active ingredients are also generally fully practical to use when coating control is precise enough to meter relatively thin uniform films of working composition onto the metal surface to be treated according to the invention. Concentrations of at least six times the values given above are also preferred for concentrated compositions from which working compositions within the more preferred ranges given above are to be made by dilution with water.

Ordinarily, at least for economy, all of the optional components indicated above are preferably omitted, because satisfactory quality can be obtained without them. They may be useful in special situations, however.

Preferably the amount of composition applied in a process according to this invention is chosen so as to result in a total add-on mass (after drying) in the range from 5 to 500 milligrams per square meter of the substrate surface treated (this unit of add-on mass being hereinafter usually abbreviated as "mg/m²"), more preferably from 10 to 400 mg/m², or still more preferably from 50 to 300 mg/m². The add-on mass of the protective

film formed by a process according to the invention may be conveniently monitored and controlled by measuring the add-on weight or mass of the metal atoms in the anions of component (A) as defined above. The amount of these metal atoms may be measured by any of several conventional analytical techniques known to those skilled in the art.

5 The most reliable measurements generally involve dissolving the coating from a known area of coated substrate and determining the content of the metal of interest in the resulting solution. The total add-on mass can then be calculated from the known relationship between the amount of the metal in component (A) and the total mass of the part of the total composition that remains after drying. For the purpose of this calculation it is assumed that all water in the working composition, including any water of hydration in
10 any solid constituent added to the composition during its preparation, is expelled by drying but that all other constituents of the liquid film of working composition coated onto the surface measured remain in the dried coating. In many instances, fully practically satisfactory quality can be achieved by experience in judging the visual appearance of the coating, without directly measuring the amount of coating added on at all.
15

A working composition according to the invention may be applied to a metal workpiece and dried thereon by any convenient method, several of which will be readily apparent to those skilled in the art. For example, coating the metal with a liquid film may be accomplished by immersing the surface in a container of the liquid composition, spraying the composition on the surface, coating the surface by passing it between upper
20 and lower rollers with the lower roller immersed in a container of the liquid composition, and the like, or by a mixture of methods. Excessive amounts of the liquid composition that might otherwise remain on some part of the surface prior to drying may be removed before drying by any convenient method, such as drainage under the influence of gravity, squeegees, passing between rolls, wiping with a towel or other absorbent material, and
25 the like. Alternatively, the excess may simply be dried into place on the surface, and any area of the dried substrate that exhibits a powdery appearance as a result of dried excess amounts of coating material can usually be made uniform in appearance by wiping the powdery areas with a soft cloth or similar material to remove the powdery excess of dried coating. (Insofar as it has been measured, the protective quality of the
30 coating is also uniform after such wiping to remove any powdery excess.) Drying also may be accomplished by any convenient method, such as a hot air oven, exposure to infra-red radiation, microwave heating, and the like.

For flat and particularly for continuous flat workpieces such as sheet and coil
35 stock, application by a roller set in any of several conventional arrangements, followed by drying in a separate stage, is generally preferred. The temperature during application

of the liquid composition may be any temperature within the liquid range of the composition, although for convenience and economy in application by roller coating, normal room temperature, i.e., from 20 - 30 °C, is usually preferred. In most cases for continuous processing of coils, rapid operation is favored, and in such cases drying by infrared radiative heating, to produce a peak metal temperature in the range already given above, is generally preferred.

Alternatively, particularly if the shape of the substrate is not suitable for roll coating, a composition may be sprayed onto the surface of the substrate and allowed to dry in place. Such cycles can be repeated as often as needed until the desired thickness of coating, generally measured in mg/m², is achieved. For this type of operation, it is preferred that the temperature of the metal substrate surface during application of the working composition be in the range from 20 to 300, more preferably from 30 to 100, or still more preferably from 30 to 90 °C.

Preferably, the metal surface to be treated according to the invention is first cleaned of any contaminants, particularly organic contaminants and foreign metal fines and/or inclusions. Such cleaning may be accomplished by methods known to those skilled in the art and adapted to the particular type of metal substrate to be treated. For example, if predominantly chemical cleaning is desired, for galvanized steel substrates, the surface is most preferably cleaned with a conventional hot alkaline cleaner, then rinsed with hot water, squeegeed, and dried. For aluminum substrates, the surface to be treated most preferably is first contacted with a conventional hot alkaline cleaner, then rinsed in hot water, then, optionally, contacted with a neutralizing acid rinse, before being contacted with an acid aqueous composition as described above. Abrasive cleaning, particularly with high-loft type coated abrasive products, may also be used effectively before treatment according to this invention.

The invention is particularly well adapted to treating surfaces that are to be subsequently further protected by applying conventional organic protective coatings such as paint, lacquer, and the like over the surface produced by treatment according to the invention.

The practice of this invention may be further appreciated by consideration of the following, non-limiting, working and comparison examples.

An aqueous solution of a polymer for component (D) of a composition according to the invention was prepared as follows: To a 500 milliliter (hereinafter usually abbreviated as "ml") size three-neck round bottom flask equipped with an overhead stirrer, reflux condenser, nitrogen gas inlet, heating mantle and thermocouple, about 350 grams (hereinafter usually abbreviated as "g") of deionized water, 4.2 grams of sodium hydroxide,

and 29.1 g of N-methyl ethanolamine were added and mixed till dissolved. 48.4 g of solid poly(4-hydroxystyrene) with a weight average molecular weight of about 5000 Daltons was then added with mixing. The mixture was then heated to 65 °C with mechanical agitation. After 1.5 hours of mixing at 65 °C the polymer was completely dissolved. The resulting solution was then allowed to cool to 30 °C and 12.5 g of paraformaldehyde containing 92 % stoichiometric equivalent as formaldehyde (i.e., HCHO) was added with continued mixing. After 30 minutes, the paraformaldehyde was completely dissolved. The solution was then heated with mixing to 65 °C and held at that temperature for 1.5 hours. Heating was then discontinued, and after the solution had cooled to about 30 °C, the solution was transferred to a 1500 ml beaker equipped with a stirrer, and therein diluted with about 500 grams of deionized water. The resulting solution was well mixed; 33.4 g of 75 % phosphoric acid (i.e., H_3PO_4) in water was added quickly. After this addition, the solution changed from transparent to very turbid; the solution then became transparent again within one hour. The pH at this stage was 6.7.

The solution was then passed through an ion-exchange column containing at least 100 ml of Rohm and Haas IR-120+ cation exchange resin. The resulting pH was 5.4; the sodium concentration was less than 1 part of sodium per million parts of the solution, a concentration unit that may be freely used hereinafter for specifying the concentration of any noted material and is hereinafter usually abbreviated as "ppm", and the residual formaldehyde concentration was about 90 ppm. 3.0 g of 60 % fluorotitanic acid (i.e., H_2TiF_6) and sufficient deionized water to give a total mass of 1000 g of the solution were then added with mixing. The final solution contained 10.2 % non-volatile solids on drying.

A concentrate according to the invention was prepared as follows: 24.6 parts of deionized water, 3.51 parts of 75 % H_3PO_4 in water solution, and 8.22 parts of 60 % H_2TiF_6 in water solution were mixed to form a homogeneous liquid. To this was slowly added 2.44 parts of solid MnO , with stirring and cooling, and after apparent homogeneity had been achieved, stirring was continued for 30 minutes. Finally, to this mixture was added 42.2 parts of the 10.2 % solution of substituted aminomethylated phenolic polymer prepared as described in the immediately preceding paragraph.

A working composition according to the invention was made by mixing the concentrate described in the immediately preceding paragraph with deionized water to give a homogeneous liquid containing 15 % of the concentrate. For use, this working composition was put into the reservoir of an applicator as taught in U. S. Patent 5,702,759.

Conventional cold rolled steel test panels were prepared for use by rubbing with a water-wetted Scotch-Brite™ 96 General Purpose Scouring Pad (a high loft coated

abrasive pad commercially supplied by Minnesota Mining & Manufacturing Co.) once in a lengthwise direction, next in a cross-direction, and once again in a lengthwise direction, then rinsing with hot tap water and finally wiping dry with a clean, dry, lint-free towel (Kay-dry™ EX-L 34705 Delicate Task Wiper commercially supplied by Kimberly-Clark) immediately before contact with whatever treatment liquid was to be used on the particular panels. In a process example according to the invention, panels thus cleaned were wiped sparingly with the felt tip of the applicator wetted with the working composition described in the immediately preceding paragraph. Each wiping stroke of the applicator was spaced so that about half of the width of the immediately previously coated width of the substrate was overlapped, but puddling of the liquid was avoided. (The wetness of the felt tip can be controlled by activating the plunger valve of the applicator more or less often and/or by using shorter or longer periods of valve opening. If any excessive amount of liquid is deposited in a particular area, the excess amount of it can be removed by wiping with the applicator felt after its most recent supply of liquid from the reservoir of the applicator has been substantially diminished by contact with another part of the metal substrate.) The residue of liquid was then allowed to dry in the ambient air. Some streaked areas of white dust, indicative of more than optimal liquid coating thickness, were observed on the surface after drying. These areas were gently brushed away with a clean, soft, dry towel before further treatment of the panels, and underneath these formerly dusty areas the same bluish coating as on the remainder of the panel was observed.

In Comparison Example 1, abrasive cleaning and subsequent drying as described above for the example according to the invention were used without any subsequent treatment before painting. In the remaining comparison examples, the same abrasive cleaning and subsequent drying were used prior to the following prepainting treatments as specified:

Comparison Example 2: The cleaned substrates were treated with TOUCH-N-PREP® ALODINE® 1132, a commercial product available from the Henkel Surface Technologies Div. of Henkel Corporation, Madison Heights, Michigan, U.S.A. This product contains as its active ingredients hexavalent and trivalent chromium, fluorozirconic acid, and phosphoric acid and also contains a fluorinated surfactant.

Comparison Example 3: This was first treated as for Comparison Example 2, dried, and then post-rinsed with a 0.25 % solution in water of the aqueous solution in water of poly(vinyl phenol) grafted with substituted aminomethyl moieties that was used to provide component (D) for the concentrate according to the invention as

described above.

Comparison Example 4: An aqueous solution of an amine oxide type substituted phenolic polymer was prepared as follows: To a 2000 ml size three-neck round bottom flask equipped with an overhead stirrer, reflux condenser, nitrogen gas inlet, heating mantle and thermocouple, about 1300 g of deionized water, 18.8 g of sodium hydroxide, 129.1 g of N-methyl ethanolamine, and 215 g of solid poly(4-hydroxy styrene) with a weight average molecular weight of about 5000 Daltons were added and mixed till dissolved. The mixture was then heated to 65 °C with mixing. After 1.5 hours of mixing at 65 °C the polymer and all other materials added were completely dissolved. The resulting solution was then allowed to cool to 30 °C and 55.6 g of paraformaldehyde containing 92 % stoichiometric equivalent as formaldehyde (i.e., HCHO) and 114 g of additional deionized water were added with continued mixing. After 30 minutes, the paraformaldehyde was completely dissolved. The solution was then heated with mixing to 65 °C. and held at that temperature for 1.5 hours. Heating was then discontinued, and after the solution had cooled to about 30 °C, the solution was diluted with 1784 g of additional deionized water, then mixed rapidly with 104 g of 31 % H₂O₂ in water and 28 g of 14.6 % H₂O₂ in water. The solution became viscous with a heavy precipitate, but after hand mixing and continued mechanical stirring for about 40 minutes, the mixture became transparent and homogeneous again. The resulting mixture was continuously stirred mechanically for about 16 hours, and was then passed through an ion-exchange column filled with 500 ml of acid form Rohm and Haas IR-120+ cation exchange resin. A yield of 3.8 kilograms (hereinafter usually abbreviated as "kg") of solution with a pH of 8.1 resulted. To this was added 0.19 kg of 20 % H₂ZrF₆ solution in water, resulting in a total mass of 4.0 kg. Finally, to 1.0 kg of this, 46 g of additional 20 % H₂ZrF₆ solution in water was added, resulting in a final amine oxide substituted phenolic polymer solution with a pH of 3.4. The working composition for Comparison Example 4 consisted of a 0.25 % solution in water of this final amine oxide substituted phenolic polymer solution.

Comparison Example 5: For this example, an aqueous solution of a substituted amino-methylated phenolic polymer with polyhydroxyalkyl substituents on the amino nitrogen atoms was made as follows: 83 parts of propoxylated propane solvent (PROPASOL™ P from Union Carbide) and 38 parts of solid poly(4-hydroxy styrene) with a weight average molecular weight of 5000 Daltons were mixed until homogeneous. Then 62 parts of N-methyl glucamine slurried in about 100

parts of deionized water were added to this mixture, and the resulting mixture was warmed with stirring to 65 °C, after which 25 parts of a solution in water of 37 % formaldehyde (also containing 11 % of pure methanol to inhibit polymerization of the formaldehyde) were added over the course of 45 minutes.

5 The resulting reaction product containing mixture was then heated to 90 °C and held at that temperature for 6 hours. After cooling, 4.2 parts of 75 % H₃PO₄ solution in water was added, and finally the entire mixture was diluted with deionized water to constitute 1000 total parts.

To make the working treatment composition for Comparative Example 5, 602 parts of the polymer solution made as described in the immediately preceding paragraph was mixed with a precursor mixture formed by mixing 45 parts of 75 % H₃PO₄ solution in water, 82 parts of a 60 % solution of H₂TiF₆, and 24 parts of solid MnO, the latter being added slowly in solid form to the mixture of the two acids with stirring and cooling.

15 Comparison Example 6: The working treatment composition for this was GALVAPREP™ SG nickel modified zinc phosphating composition, prepared and used as directed by its manufacturer, the Henkel Surface Technologies Div. of Henkel Corporation, Madison Heights, Michigan, U.S.A.

20 Three panels were treated with the working composition according to the invention as described above and according to each of the Comparison Examples 1 - 6 as described above and were subsequently painted with PPG ED5050B cathodically electrodeposited paint. The painted panels were then submitted to accelerated corrosion testing, as further detailed in Table 1 below.

Table 1

Test Identification	Result after Test for:						
	Ex- ample	Comparative Example:					
		1	2	3	4	5	6
Ford Scab (FLTM BI 123-01)	1.6	36	1.4	1.8	33	24	24
Salt Spray (504 Hours) — Rated by:							
Maximum Creep, Millimeters	1.4 ^s	3.2 ^s	1.4 ^s	1.5 ^s	3.4 ^s	1.4 ^s	1.8 ^s
Average Creep, Millimeters	0.9	1.3	0.8	0.8	2.2	0.7	0.6
ASTM Overall	8	7	8	8	6	8	8

In addition to the results shown in Table 1, all of the samples had the same ratings in other tests or rating methods, as follows: Salt spray rated by minimum creep, 00; Conical

Mandrel, 10; Reverse Impact, 10; Knife Adhesion, 5B; and 1008 hours humidity testing according to American Society for Testing and Materials Method D2247, 10. The example according to the invention is clearly far better in performance in the aggressive cyclic "scab" test than any of the comparison examples, except for those that utilize

5 hexavalent chromium, with its accompanying pollution concerns.

The invention claimed is:

CLAIMS

1. An acidic aqueous liquid composition that is suitable for use directly, after being diluted with water, or both directly and after being diluted with water, for generating a corrosion reducing coating over a metal surface when contacted therewith, said composition comprising water and:
- (A) a component of "fluorometallate" anions, each of said anions consisting of (i) at least four fluorine atoms, (ii) at least one atom of an element selected from the group consisting of titanium, zirconium, hafnium, silicon, aluminum, and boron, and, optionally, one or both of (iii) ionizable hydrogen atoms, and (iv) one or more oxygen atoms;
- (B) a component of divalent or tetravalent cations of elements selected from the group consisting of cobalt, magnesium, manganese, zinc, nickel, tin, copper, zirconium, iron, and strontium;
- (C) a component selected from the group consisting of phosphorus-containing inorganic oxyanions and phosphonate anions; and
- (D) a component of water-soluble, water-dispersible, or both water-soluble and water-dispersible polymers of hydroxy styrene, modified by substitution on the aromatic rings of the polymers of at least mono-substituted aminomethyl moieties, in which the substituents (other than the carbon atom that is directly bonded to an aromatic ring in the polymer) on the amino nitrogen atom jointly contain at least two carbon atoms and at least one hydroxy moiety but neither of these substituents individually contains more than half as many hydroxyl moieties as it has carbon atoms, unless it contains only one carbon atom.
2. An acidic aqueous liquid composition according to claim 1, wherein:
- there is a total concentration of "fluorometallate" anions of component (A) that is at least about 0.010 M/kg;
 - there is a total concentration of metal cations of component (B) that is at least about 0.020 M/kg;
 - there is a ratio of total concentration in M/kg of metal cations of component (B) in M/kg to the total concentration in M/kg of "fluorometallate" anions of component (A) that is in a range from about 0.4:1.0 to about 1.6:1.00;
 - there is a concentration of phosphorus from component (C) that is at least about 0.017 M/kg;
 - there is a ratio of moles of phosphorus from component (C) to moles of "fluoro-

- metallate" ions from component (A) that is in a range from about 0.60:1.00 to about 2.6:1.00;
- there is a ratio of moles of phosphorus from component (C) to moles of total metal cations from component (B) that is from about 0.4:1.0 to about 2.2:1.00;
 - 5 - there is a concentration of moles of substituted phenol moieties from component (D) that is at least about 0.008 M/kg;
 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of total "fluorometallate" anions from component (A) that is from about 0.12:1.00 to about 1.5:1.00;
 - 10 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of total metal cations from component (B) that is from about 0.12:1.00 to about 1.5:1.00; and
 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of phosphorus from component (C) that is from about 0.12:1.00 to about 1.5:1.00.
- 15 3. An acidic aqueous liquid composition according to claim 2, wherein:
- component (A) is selected from the group consisting of fluorotitanate and fluoro-zirconate;
 - at least 60 % of component (B) is selected from the group consisting of divalent manganese, cobalt, nickel, and magnesium; and
 - 20 - there is not more than 0.10 % of organic materials that are liquid at 25 °C under normal atmospheric pressure and have a vapor pressure of at least 0.05 bar at 25 °C.
4. An acidic aqueous liquid composition according to claim 3, wherein component (D) has the chemical characteristics of a polymer that is a product of reaction of:
- 25 (A') at least one precursor phenolic polymer or copolymer which does not bear any substituted aminomethyl substituents on its aromatic rings;
 - (B') at least one aldehyde, ketone, or mixture thereof; and
 - (C') at least one amine.
- 30 5. An acidic aqueous liquid composition according to claim 4, wherein:
- the number of moles of carbonyl groups in component (B') has a ratio to the number of moles of primary and secondary amino nitrogen atoms in component (C') that is from about 0.5:1.00 to about 1.5:1.00;
 - the number of moles of carbonyl groups in component (B') has a ratio to the number of moles of aromatic rings in component (A') that is from about 0.20:1.00 to
- 35

- about 2.00:1.00; and
- the number of moles of primary and secondary amino nitrogen atoms in component (C') has a ratio to the number of moles of aromatic rings in component (A') that is from about 0.20:1.00 to about 2.00:1.00.
- 5 6. An acidic aqueous liquid composition according to claim 5, wherein component (D) has been made by a process comprising the following operations:
- (I') reacting the precursor phenolic polymer component (A') in water with an organic or inorganic alkalinizing agent to form an aqueous solution of the corresponding phenoxide salt;
 - 10 (II') mixing the aqueous solution from operation (I') with the amine component (C') and the component (B') of aldehyde, ketone, or mixture thereof to form a single aqueous solution in which chemical reaction among components (A'), (B'), and (C') occurs at a temperature in a range from about 20 to about 100 °C to attach substituted aminomethyl moieties to at least some of the aromatic rings in the precursor polymer and produce an aqueous solution of substituted aminomethylated phenolic polymer molecules;
 - 15 (III') adding at least one acid to the aqueous solution formed at the end of operation (II'), the quantity of acid added being sufficient to neutralize the alkalinizing agent added in operation (I') and to protonate a sufficient fraction of the amino nitrogen atoms in the substituted aminomethylated phenolic polymer to stabilize against settling the solution of the substituted aminomethylated phenolic polymer formed in operation (II'); and
 - 20 (IV') contacting the resulting aqueous solution from the end of operation (III') with a cation exchange resin in its protonated form to remove at least about 75 % of any inorganic and/or quaternary ammonium cations dissolved in said aqueous solution from the end of operation (III').
 - 25
7. An acidic aqueous liquid composition according to claim 6, wherein:
- there is a total concentration of "fluorometallate" anions of component (A) that is at least about 0.040 M/kg;
 - 30 - there is a total concentration of metal cations of component (B) that is at least about 0.054 M/kg;
 - there is a ratio of total concentration in M/kg of metal cations of component (B) in M/kg to the total concentration in M/kg of "fluorometallate" anions of component (A) that is in a range from about 1.00:1.00 to about 1.30:1.00;
 - 35 - there is a concentration of phosphorus from component (C) that is at least about

- 0.057 M/kg;
- there is a ratio of moles of phosphorus from component (C) to moles of "fluoro-metallate" ions from component (A) that is in a range from about 1.00:1.00 to about 1.40:1.00;
- 5
- there is a ratio of moles of phosphorus from component (C) to moles of total metal cations from component (B) that is from about 0.80:1.00 to about 1.25:1.00;
 - there is a concentration of moles of substituted phenol moieties from component (D) that is at least about 0.024 M/kg;
 - there is a ratio of moles of substituted phenol moieties from component (D) to
- 10
- moles of total "fluorometallate" anions from component (A) that is from about 0.40:1.00 to about 0.80:1.00;
 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of total metal cations from component (B) that is from about 0.35:1.00 to about 0.70:1.00; and
- 15
- there is a ratio of moles of substituted phenol moieties from component (D) to moles of phosphorus from component (C) that is from about 0.30:1.00 to about 0.65:1.00.
8. An acidic aqueous liquid composition according to claim 1, wherein:
- component (A) is selected from the group consisting of fluorotitanate and
- 20
- fluorozirconate;
 - at least 60 % of component (B) is selected from the group consisting of divalent manganese, cobalt, nickel, and magnesium; and
 - there is not more than 0.10 % of organic materials that are liquid at 25 °C under normal atmospheric pressure and have a vapor pressure of at least 0.05 bar at
- 25
- 25 °C.
9. An acidic aqueous liquid composition according to claim 8, wherein component (D) has the chemical characteristics of a polymer that is a product of reaction of:
- (A') at least one precursor phenolic polymer or copolymer which does not bear any substituted aminomethyl substituents on its aromatic rings;
- 30
- (B') at least one aldehyde, ketone, or mixture thereof; and
 - (C') at least one amine.
10. An acidic aqueous liquid composition according to claim 9, wherein:
- the number of moles of carbonyl groups in component (B') has a ratio to the number of moles of primary and secondary amino nitrogen atoms in component
- 35
- (C') that is from about 0.5:1.00 to about 1.5:1.00;

- the number of moles of carbonyl groups in component (B') has a ratio to the number of moles of aromatic rings in component (A') that is from about 0.20:1.00 to about 2.00:1.00; and
 - the number of moles of primary and secondary amino nitrogen atoms in component (C') has a ratio to the number of moles of aromatic rings in component (A') that is from about 0.20:1.00 to about 2.00:1.00.
11. An acidic aqueous liquid composition according to claim 10, wherein component (D) has been made by a process comprising the following operations:
- (I') reacting the precursor phenolic polymer component (A') in water with an organic or inorganic alkalinizing agent to form an aqueous solution of the corresponding phenoxide salt;
 - (II') mixing the aqueous solution from operation (I') with the amine component (C') and the component (B') of aldehyde, ketone, or mixture thereof to form a single aqueous solution in which chemical reaction among components (A'), (B'), and (C') occurs at a temperature in a range from about 20 to about 100 °C to attach substituted aminomethyl moieties to at least some of the aromatic rings in the precursor polymer and produce an aqueous solution of substituted aminomethylated phenolic polymer molecules;
 - (III') adding at least one acid to the aqueous solution formed at the end of operation (II'), the quantity of acid added being sufficient to neutralize the alkalinizing agent added in operation (I') and to protonate a sufficient fraction of the amino nitrogen atoms in the substituted aminomethylated phenolic polymer to stabilize against settling the solution of the substituted aminomethylated phenolic polymer formed in operation (II'); and
 - (IV') contacting the resulting aqueous solution from the end of operation (III') with a cation exchange resin in its protonated form to remove at least about 75 % of any inorganic and/or quaternary ammonium cations dissolved in said aqueous solution from the end of operation (III').
12. An acidic aqueous liquid composition according to claim 11, wherein:
- there is a total concentration of "fluorometallate" anions of component (A) that is at least about 0.040 M/kg;
 - there is a total concentration of metal cations of component (B) that is at least about 0.054 M/kg;
 - there is a ratio of total concentration in M/kg of metal cations of component (B) in M/kg to the total concentration in M/kg of "fluorometallate" anions of

- component (A) that is in a range from about 1.00:1.00 to about 1.30:1.00;
- there is a concentration of phosphorus from component (C) that is at least about 0.057 M/kg;
 - there is a ratio of moles of phosphorus from component (C) to moles of "fluoro-metallate" ions from component (A) that is in a range from about 1.00:1.00 to about 1.40:1.00;
 - there is a ratio of moles of phosphorus from component (C) to moles of total metal cations from component (B) that is from about 0.80:1.00 to about 1.25:1.00;
 - there is a concentration of moles of substituted phenol moieties from component (D) that is at least about 0.024 M/kg;
 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of total "fluorometallate" anions from component (A) that is from about 0.40:1.00 to about 0.80:1.00;
 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of total metal cations from component (B) that is from about 0.35:1.00 to about 0.70:1.00; and
 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of phosphorus from component (C) that is from about 0.30:1.00 to about 0.65:1.00.
13. An acidic aqueous liquid composition according to claim 1, wherein component (D) has the chemical characteristics of a polymer that is a product of reaction of:
- (A') at least one precursor phenolic polymer or copolymer which does not bear any substituted aminomethyl substituents on its aromatic rings;
 - (B') at least one aldehyde, ketone, or mixture thereof; and
 - (C') at least one amine.
14. An acidic aqueous liquid composition according to claim 13, wherein:
- the number of moles of carbonyl groups in component (B') has a ratio to the number of moles of primary and secondary amino nitrogen atoms in component (C') that is from about 0.5:1.00 to about 1.5:1.00;
 - the number of moles of carbonyl groups in component (B') has a ratio to the number of moles of aromatic rings in component (A') that is from about 0.20:1.00 to about 2.00:1.00; and
 - the number of moles of primary and secondary amino nitrogen atoms in component (C') has a ratio to the number of moles of aromatic rings in component (A') that is from about 0.20:1.00 to about 2.00:1.00.

15. An acidic aqueous liquid composition according to claim 14, wherein component (D) has been made by a process comprising the following operations:

- (I') reacting the precursor phenolic polymer component (A') in water with an organic or inorganic alkalinizing agent to form an aqueous solution of the corresponding phenoxide salt;
- (II') mixing the aqueous solution from operation (I') with the amine component (C') and the component (B') of aldehyde, ketone, or mixture thereof to form a single aqueous solution in which chemical reaction among components (A'), (B'), and (C') occurs at a temperature in a range from about 20 to about 100 °C to attach substituted aminomethyl moieties to at least some of the aromatic rings in the precursor polymer and produce an aqueous solution of substituted aminomethylated phenolic polymer molecules;
- (III') adding at least one acid to the aqueous solution formed at the end of operation (II'), the quantity of acid added being sufficient to neutralize the alkalinizing agent added in operation (I') and to protonate a sufficient fraction of the amino nitrogen atoms in the substituted aminomethylated phenolic polymer to stabilize against settling the solution of the substituted aminomethylated phenolic polymer formed in operation (II'); and
- (IV') contacting the resulting aqueous solution from the end of operation (III') with a cation exchange resin in its protonated form to remove at least about 75 % of any inorganic and/or quaternary ammonium cations dissolved in said aqueous solution from the end of operation (III').

16. An acidic aqueous liquid composition according to claim 15, wherein:

- there is a total concentration of "fluorometallate" anions of component (A) that is at least about 0.040 M/kg;
- there is a total concentration of metal cations of component (B) that is at least about 0.054 M/kg;
- there is a ratio of total concentration in M/kg of metal cations of component (B) in M/kg to the total concentration in M/kg of "fluorometallate" anions of component (A) that is in a range from about 1.00:1.00 to about 1.30:1.00;
- there is a concentration of phosphorus from component (C) that is at least about 0.057 M/kg;
- there is a ratio of moles of phosphorus from component (C) to moles of "fluorometallate" ions from component (A) that is in a range from about 1.00:1.00 to about 1.40:1.00;

- there is a ratio of moles of phosphorus from component (C) to moles of total metal cations from component (B) that is from about 0.80:1.00 to about 1.25:1.00;
 - there is a concentration of moles of substituted phenol moieties from component (D) that is at least about 0.024 M/kg;
 - 5 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of total "fluorometallate" anions from component (A) that is from about 0.40:1.00 to about 0.80:1.00;
 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of total metal cations from component (B) that is from about 0.35:1.00 to about 0.70:1.00; and
 - 10 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of phosphorus from component (C) that is from about 0.30:1.00 to about 0.65:1.00.
17. An acidic aqueous liquid composition according to claim 13, wherein component (D) has been made by a process comprising the following operations:
- 15 (I') reacting the precursor phenolic polymer component (A') in water with an organic or inorganic alkalinizing agent to form an aqueous solution of the corresponding phenoxide salt;
- 20 (II') mixing the aqueous solution from operation (I') with the amine component (C') and the component (B') of aldehyde, ketone, or mixture thereof to form a single aqueous solution in which chemical reaction among components (A'), (B'), and (C') occurs at a temperature in a range from about 20 to about 100 °C to attach substituted aminomethyl moieties to at least some of the aromatic rings in the precursor polymer and produce an aqueous solution of substituted aminomethylated phenolic polymer molecules;
- 25 (III') adding at least one acid to the aqueous solution formed at the end of operation (II'), the quantity of acid added being sufficient to neutralize the alkalinizing agent added in operation (I') and to protonate a sufficient fraction of the amino nitrogen atoms in the substituted aminomethylated phenolic polymer to stabilize against settling the solution of the substituted aminomethylated phenolic polymer formed in operation (II'); and
- 30 (IV') contacting the resulting aqueous solution from the end of operation (III') with a cation exchange resin in its protonated form to remove at least about 75 % of any inorganic and/or quaternary ammonium cations dissolved in said aqueous solution from the end of operation (III').
- 35

18. An acidic aqueous liquid composition according to claim 17, wherein:
- there is a total concentration of "fluorometallate" anions of component (A) that is at least about 0.040 M/kg;
 - there is a total concentration of metal cations of component (B) that is at least about 0.054 M/kg;
 - there is a ratio of total concentration in M/kg of metal cations of component (B) in M/kg to the total concentration in M/kg of "fluorometallate" anions of component (A) that is in a range from about 1.00:1.00 to about 1.30:1.00;
 - there is a concentration of phosphorus from component (C) that is at least about 0.057 M/kg;
 - there is a ratio of moles of phosphorus from component (C) to moles of "fluorometallate" ions from component (A) that is in a range from about 1.00:1.00 to about 1.40:1.00;
 - there is a ratio of moles of phosphorus from component (C) to moles of total metal cations from component (B) that is from about 0.80:1.00 to about 1.25:1.00;
 - there is a concentration of moles of substituted phenol moieties from component (D) that is at least about 0.024 M/kg;
 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of total "fluorometallate" anions from component (A) that is from about 0.40:1.00 to about 0.80:1.00;
 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of total metal cations from component (B) that is from about 0.35:1.00 to about 0.70:1.00; and
 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of phosphorus from component (C) that is from about 0.30:1.00 to about 0.65:1.00.
19. An acidic aqueous liquid composition according to claim 1, wherein component (D) has been made by a process comprising the following operations:
- (I') reacting a precursor phenolic polymer component (A') in water with an organic or inorganic alkalinizing agent to form an aqueous solution of the corresponding phenoxide salt;
 - (II') mixing the aqueous solution from operation (I') with an amine component (C') and a component (B') of aldehyde, ketone, or mixture thereof to form a single aqueous solution in which chemical reaction among components (A'), (B'), and (C') occurs at a temperature in a range from about 20 to about 100 °C to attach

substituted aminomethyl moieties to at least some of the aromatic rings in the precursor polymer and produce an aqueous solution of substituted aminomethylated phenolic polymer molecules;

- 5 (III') adding at least one acid to the aqueous solution formed at the end of operation (II'), the quantity of acid added being sufficient to neutralize the alkalizing agent added in operation (I') and to protonate a sufficient fraction of the amino nitrogen atoms in the substituted aminomethylated phenolic polymer to stabilize against settling the solution of the substituted aminomethylated phenolic polymer formed in operation (II'); and
- 10 (IV') contacting the resulting aqueous solution from the end of operation (III') with a cation exchange resin in its protonated form to remove at least about 75 % of any inorganic and/or quaternary ammonium cations dissolved in said aqueous solution from the end of operation (III').
20. An acidic aqueous liquid composition according to claim 19, wherein:
- 15 - there is a total concentration of "fluorometallate" anions of component (A) that is at least about 0.040 M/kg;
- there is a total concentration of metal cations of component (B) that is at least about 0.054 M/kg;
- there is a ratio of total concentration in M/kg of metal cations of component (B) in M/kg to the total concentration in M/kg of "fluorometallate" anions of component (A) that is in a range from about 1.00:1.00 to about 1.30:1.00;
- 20 - there is a concentration of phosphorus from component (C) that is at least about 0.057 M/kg;
- there is a ratio of moles of phosphorus from component (C) to moles of "fluorometallate" ions from component (A) that is in a range from about 1.00:1.00 to about 1.40:1.00;
- 25 - there is a ratio of moles of phosphorus from component (C) to moles of total metal cations from component (B) that is from about 0.80:1.00 to about 1.25:1.00;
- there is a concentration of moles of substituted phenol moieties from component (D) that is at least about 0.024 M/kg;
- 30 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of total "fluorometallate" anions from component (A) that is from about 0.40:1.00 to about 0.80:1.00;
- there is a ratio of moles of substituted phenol moieties from component (D) to moles of total metal cations from component (B) that is from about 0.35:1.00 to
- 35

about 0.70:1.00; and

- there is a ratio of moles of substituted phenol moieties from component (D) to moles of phosphorus from component (C) that is from about 0.30:1.00 to about 0.65:1.00.

5 21. An acidic aqueous liquid composition according to claim 1, wherein:

- there is a total concentration of "fluorometallate" anions of component (A) that is at least about 0.040 M/kg;
- there is a total concentration of metal cations of component (B) that is at least about 0.054 M/kg;
- 10 - there is a ratio of total concentration in M/kg of metal cations of component (B) in M/kg to the total concentration in M/kg of "fluorometallate" anions of component (A) that is in a range from about 1.00:1.00 to about 1.30:1.00;
- there is a concentration of phosphorus from component (C) that is at least about 0.057 M/kg;
- 15 - there is a ratio of moles of phosphorus from component (C) to moles of "fluorometallate" ions from component (A) that is in a range from about 1.00:1.00 to about 1.40:1.00;
- there is a ratio of moles of phosphorus from component (C) to moles of total metal cations from component (B) that is from about 0.80:1.00 to about 1.25:1.00;
- 20 - there is a concentration of moles of substituted phenol moieties from component (D) that is at least about 0.024 M/kg;
- there is a ratio of moles of substituted phenol moieties from component (D) to moles of total "fluorometallate" anions from component (A) that is from about 0.40:1.00 to about 0.80:1.00;
- 25 - there is a ratio of moles of substituted phenol moieties from component (D) to moles of total metal cations from component (B) that is from about 0.35:1.00 to about 0.70:1.00; and
- there is a ratio of moles of substituted phenol moieties from component (D) to moles of phosphorus from component (C) that is from about 0.30:1.00 to about 0.65:1.00.
- 30

22. An acidic aqueous liquid composition that is suitable for use directly, after being diluted with water, or both directly and after being diluted with water, for generating a corrosion reducing coating over a metal surface when contacted therewith, said composition having been made by mixing with water at least the following components:

- 35 (A) a source of a component of "fluorometallate" anions, each of said anions consist-

- ing of (i) at least four fluorine atoms, (ii) at least one atom of an element selected from the group consisting of titanium, zirconium, hafnium, silicon, aluminum, and boron, and, optionally, one or both of (iii) ionizable hydrogen atoms, and (iv) one or more oxygen atoms;
- 5 (B) a source of a component of divalent or tetravalent cations of elements selected from the group consisting of cobalt, magnesium, manganese, zinc, nickel, tin, copper, zirconium, iron, and strontium;
- (C) a source of a component selected from the group consisting of phosphorus-containing inorganic oxyanions and phosphonate anions; and
- 10 (D) a source of a component of water-soluble, water-dispersible, or both water-soluble and water-dispersible polymers of hydroxy styrene, modified by substitution on the aromatic rings of the polymers of at least mono-substituted aminomethyl moieties, in which the substituents (other than the carbon atom that is directly bonded to an aromatic ring in the polymer) on the amino nitrogen atom jointly
- 15 contain at least two carbon atoms and at least one hydroxy moiety but neither of these substituents individually contains more than half as many hydroxyl moieties as it has carbon atoms, unless it contains only one carbon atom.
23. An acidic aqueous liquid composition according to claim 22, wherein:
- the source of "fluorometallate" anions provides such anions in an amount corresponding to a concentration in the acidic aqueous liquid composition that is at least about 0.010 M/kg;
 - the source of metal cations of component (B) provides such cations in an amount corresponding to a concentration in the acidic aqueous liquid composition that is at least about 0.020 M/kg;
 - 25 - there is a ratio of moles of metal cations of component (B) supplied to the acidic aqueous liquid composition to moles of "fluorometallate" anions supplied to the acidic aqueous liquid composition that is in a range from about 0.4:1.0 to about 1.6:1.00;
 - the source of phosphorus for component (C) provides phosphorus in an amount corresponding to a concentration in an acidic aqueous liquid composition that is at least about 0.017 M/kg;
 - 30 - there is a ratio of moles of phosphorus from component (C) supplied to the acidic aqueous liquid composition to moles of "fluorometallate" ions from component (A) supplied to the acidic aqueous liquid composition that is in a range from about 0.60:1.00 to about 2.6:1.00;
 - 35

- there is a ratio of moles of phosphorus from component (C) supplied to the acidic aqueous liquid composition to moles of total metal cations from component (B) supplied to the acidic aqueous liquid composition that is from about 0.4:1.0 to about 2.2:1.00;
- 5 - the source of substituted phenol moieties for component (D) provides such moieties in an amount corresponding to a concentration in the acidic aqueous liquid composition that is at least about 0.008 M/kg;
- there is a ratio of moles of substituted phenol moieties for component (D) supplied to the acidic aqueous liquid composition to moles of total "fluorometallate" anions from component (A) supplied to the acidic aqueous liquid composition
- 10 that is from about 0.12:1.00 to about 1.5:1.00;
- there is a ratio of moles of substituted phenol moieties from component (D) supplied to the acidic aqueous liquid composition to moles of total metal cations from component (B) supplied to the acidic aqueous liquid composition that is
- 15 from about 0.12:1.00 to about 1.5:1.00; and
- there is a ratio of moles of substituted phenol moieties from component (D) supplied to the acidic aqueous liquid composition to moles of phosphorus from component (C) that is from about 0.12:1.00 to about 1.5:1.00.
- 24. An acidic aqueous liquid composition according to claim 23, wherein:
- 20 - component (A) is selected from the group consisting of fluorotitanate and fluorozirconate;
- at least 60 % of component (B) is selected from the group consisting of divalent manganese, cobalt, nickel, and magnesium; and
- there is not more than 0.10 % of organic materials that are liquid at 25 °C under
- 25 normal atmospheric pressure and have a vapor pressure of at least 0.05 bar at 25 °C.
- 25. An acidic aqueous liquid composition according to claim 24, wherein component (D) has the chemical characteristics of a polymer that is a product of reaction of:
- (A') at least one precursor phenolic polymer or copolymer which does not bear any
- 30 substituted aminomethyl substituents on its aromatic rings;
- (B') at least one aldehyde, ketone, or mixture thereof; and
- (C') at least one amine.
- 26. An acidic aqueous liquid composition according to claim 25, wherein:
- the number of moles of carbonyl groups in component (B') has a ratio to the
- 35 number of moles of primary and secondary amino nitrogen atoms in component

- (C') that is from about 0.5:1.00 to about 1.5:1.00;
- the number of moles of carbonyl groups in component (B') has a ratio to the number of moles of aromatic rings in component (A') that is from about 0.20:1.00 to about 2.00:1.00; and
 - 5 - the number of moles of primary and secondary amino nitrogen atoms in component (C') has a ratio to the number of moles of aromatic rings in component (A') that is from about 0.20:1.00 to about 2.00:1.00.
27. An acidic aqueous liquid composition according to claim 26, wherein component (D) has been made by a process comprising the following operations:
- 10 (I') reacting the precursor phenolic polymer component (A') in water with an organic or inorganic alkalinizing agent to form an aqueous solution of the corresponding phenoxide salt;
 - (II') mixing the aqueous solution from operation (I') with the amine component (C') and the component (B') of aldehyde, ketone, or mixture thereof to form a single
15 aqueous solution in which chemical reaction among components (A'), (B'), and (C') occurs at a temperature in a range from about 20 to about 100 °C to attach substituted aminomethyl moieties to at least some of the aromatic rings in the precursor polymer and produce an aqueous solution of substituted aminomethylated phenolic polymer molecules;
 - 20 (III') adding at least one acid to the aqueous solution formed at the end of operation (II'), the quantity of acid added being sufficient to neutralize the alkalinizing agent added in operation (I') and to protonate a sufficient fraction of the amino nitrogen atoms in the substituted aminomethylated phenolic polymer to stabilize against settling the solution of the substituted aminomethylated phenolic polymer formed
25 in operation (II'); and
 - (IV') contacting the resulting aqueous solution from the end of operation (III') with a cation exchange resin in its protonated form to remove at least about 75 % of any inorganic and/or quaternary ammonium cations dissolved in said aqueous solution from the end of operation (III').
- 30 28. An acidic aqueous liquid composition according to claim 27, wherein:
- the source of "fluorometallate" anions of component (A) provides such anions in an amount corresponding to a concentration in the acidic aqueous liquid composition that is at least about 0.040 M/kg;
 - the source of metal cations of component (B) provides such cations in an amount
35 corresponding to a concentration in the acidic aqueous liquid composition that

is at least about 0.054 M/kg;

- there is a ratio of moles of metal cations of component (B) supplied to the acidic aqueous liquid composition to the ratio of moles of "fluorometallate" anions supplied to the acidic aqueous liquid composition that is in a range from about 1.00:1.00 to about 1.30:1.00;
- the source of phosphorus for component (C) provides phosphorus in an amount corresponding to a concentration in the acidic aqueous liquid composition that is at least about 0.057 M/kg;
- there is a ratio of moles of phosphorus from component (C) supplied to the acidic aqueous liquid composition to moles of "fluorometallate" ions from component (A) supplied to the acidic aqueous liquid composition that is in a range from about 1.00:1.00 to about 1.40:1.00;
- there is a ratio of moles of phosphorus from component (C) supplied to the acidic aqueous liquid composition to moles of total metal cations from component (B) supplied to the acidic aqueous liquid composition that is from about 0.80:1.00 to about 1.25:1.00;
- the source of substituted phenol moieties from component (D) provides such moieties in an amount corresponding to a concentration in the acidic aqueous liquid composition that is at least about 0.024 M/kg;
- there is a ratio of moles of substituted phenol moieties from component (D) supplied to the acidic aqueous liquid composition to moles of total "fluorometallate" anions from component (A) supplied to the acidic aqueous liquid composition that is from about 0.40:1.00 to about 0.80:1.00;
- there is a ratio of moles of substituted phenol moieties from component (D) supplied to the acidic aqueous liquid composition to moles of total metal cations from component (B) supplied to the acidic aqueous liquid composition that is from about 0.35:1.00 to about 0.70:1.00; and
- there is a ratio of moles of substituted phenol moieties from component (D) supplied to the acidic aqueous liquid composition to moles of phosphorus from component (C) supplied to the acidic aqueous liquid composition that is from about 0.30:1.00 to about 0.65:1.00.

29. A process of forming a corrosion reducing coating over a metal surface, said process comprising contacting the metal surface with an acidic aqueous liquid composition according to any one of claims 1 through 28.

30. A process according to claim 29 that produces a dried add-on mass of coating

per unit area of surface coated that is from about 50 to about 300 mg/m².

INTERNATIONAL SEARCH REPORT

International application No.
PCT/US00/29266

A. CLASSIFICATION OF SUBJECT MATTER

IPC(7) : C23C 22/00

US CL : 148/251

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

U.S. : 148/251, 247, 259, 260

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

EAST

search terms: fluorometallate, phenol, phenolic, polymer, styrenece, aminomethylated, metal, coating

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
Y	US 5,958,511 A (DOLAN) 28 September 1999, abstract, col. 2, lines 37-68, col. 3, lines 1-9, col. 4, lines 25-41, claim 1.	1-30
Y	US 5,891,952 A (MCCORMICK et al) 06 April 1999, abstract, col. 1, line 39 to col. 5, line 33.	1-30
Y	US 5,885,373 A (SIENKOWSKI) 23 March 1999, abstract, col. 3, lines 14-54, col. 5, lines 49-65.	1-30
Y	US 5,449,415 A (DOLAN) 12 September 1995, abstract, col. 2, lines 11-51, col. 5, line 66 to col. 6, line 18, claim 1.	1-30



Further documents are listed in the continuation of Box C.



See patent family annex.

"

Special categories of cited documents:

"T"

later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

A

document defining the general state of the art which is not considered to be of particular relevance

X

document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

B

earlier document published on or after the international filing date

Y

document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art

L

document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)

O

document referring to an oral disclosure, use, exhibition or other means

G

document member of the same patent family

P

document published prior to the international filing date but later than the priority date claimed

Date of the actual completion of the international search

01 FEBRUARY 2001

Date of mailing of the international search report

15 FEB 2001

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INTERNATIONAL SEARCH REPORT

International application No.
PCT/US00/29266

C (Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of documents, with indication, where appropriate, of the relevant passages	Relevant to claim No.
Y	US 5,116,912 A (LINDERT et al) 26 May 1992, abstract, col. 3, line 13 to col. 8, line 53, example 1, claim 1.	1-30

Optical gain in silicon nanocrystals

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Adding optical functionality to a silicon microelectronic chip is one of the most challenging problems of materials research. Silicon is an indirect-bandgap semiconductor and so is an inefficient emitter of light. For this reason, integration of optically functional elements with silicon microelectronic circuitry has largely been achieved through the use of direct-bandgap compound semiconductors. For optoelectronic applications, the key device is the light source—a laser. Compound semiconductor lasers exploit low-dimensional electronic systems, such as quantum wells and quantum dots, as the active optical amplifying medium. Here we demonstrate that light amplification is possible using silicon itself, in the form of quantum dots dispersed in a silicon dioxide matrix. Net optical gain is seen in both waveguide and transmission configurations, with the material gain being of the same order as that of direct-bandgap quantum dots. We explain the observations using a model based on population inversion of radiative states associated with the Si/SiO₂ interface. These findings open a route to the fabrication of a silicon laser.

Silicon, the mainstay semiconductor in microelectronic circuitry, has been considered unsuitable for optoelectronic applications owing to its indirect electronic bandgap, which limits its efficiency as a light emitter. Recently, room-temperature light emission from silicon has been shown to be possible when the silicon is in the form of a low-dimensional system^{1–5} or when selected active impurities (such as erbium⁶) and/or new phases (such as iron disilicide⁷) are inserted into the silicon lattice. All manner of low-dimensional silicon systems—such as porous silicon^{1,2,5}, silicon nanocrystals³, silicon/insulator superlattices⁴, silicon nano-pillars⁸—are being actively investigated as a means of improving the light-emission properties of silicon. The physical mechanism underlying high external quantum efficiencies for photoluminescence in low-dimensional silicon is mainly that of the quantum confinement of excitons in a nanometre-scale crystalline structure⁹, although the silicon/dielectric interface is also thought to play an active role in both the passivation of non-radiative states and the formation of radiative states¹⁰. Such work has led to many claims of a future role for silicon in photonic applications^{3,11–15}, yet a silicon laser has remained unlikely¹⁴.

To produce a silicon-based laser, we should demonstrate its light amplification or stimulated emission¹⁵. But light amplification in silicon is difficult because (1) it has efficient free carrier absorption, which reduces the net gain available for laser action¹⁴; (2) there is

significant Auger saturation of the luminescence intensity at high power²; and (3) there is significant size-dependence of the radiative energies in Si nanostructures, which yields large inhomogeneous broadening and significant optical losses in the system¹⁶. Here we report measurements of stimulated emission and light amplification in Si nanostructures and demonstrate optical gain in a single pass configuration. Population inversion is realized between the fundamental and a radiative state associated with the nanocrystal–oxide interface¹⁰. These findings could lead the way to a silicon-based laser.

Silicon nanocrystals

Low-dimensional silicon nanocrystals have been produced by negative ion implantation (80 keV; 1×10^{17} Si ions cm⁻²) into ultra-pure quartz substrates or into thermally grown silicon dioxide layers on Si substrates, followed by high-temperature thermal annealing (1,100 °C for 1 h). Quartz wafers (hereafter referred to

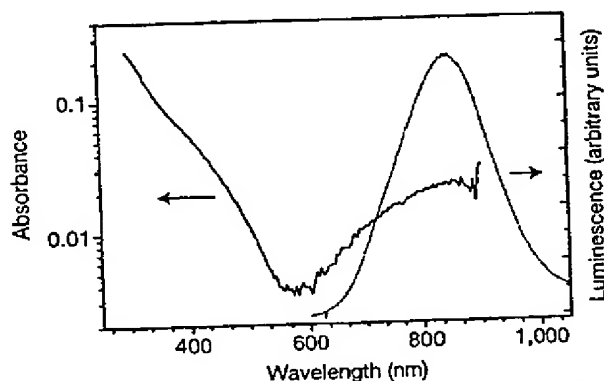


Figure 1 Room temperature absorbance and luminescence of Si nanocrystals embedded in a quartz matrix. The experimental set-up limited the absorbance measurement range. The absorbance of the quartz wafer was subtracted from the measured spectra. The 488 nm line of an Ar laser excited the luminescence.

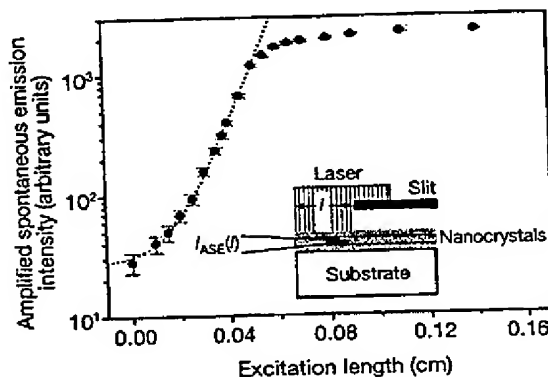


Figure 2 Amplified spontaneous emission intensity (ASE, disks) versus excitation stripe length (l) of Si nanocrystals embedded in a quartz matrix. Recording wavelength, 800 nm. A fit to the data with equation (1) is shown as the dashed line. The inset shows the experimental method. A cylindrical lens was used to focus the laser beam on the sample surface on a stripe 10- μ m wide and of variable length. Only the central part of the laterally unfocused laser spot was used to excite the sample. Measurements show that within these experimental conditions the laser power density on the sample surface is constant and independent of l . An optical 40 \times objective imaged the sample edge on a 40- μ m pinhole so that only the light coming from the near sample surface region was collected. The use of a pulsed laser avoided any thermal heating of the sample. The excitation conditions were 1 kW cm⁻² mean power at a wavelength of 390 nm.

as 'sample A') were used for optical transmission experiments, and silicon wafers (hereafter referred to as 'sample B') to demonstrate microelectronics compatibility. Transmission electron microscopy of these samples showed silicon nanocrystals embedded within the oxide matrix. They were formed in a region centred at a depth of 110 nm from the sample surface and extending for a thickness of 100 nm; they were ~3 nm in diameter, with a concentration of $2 \times 10^{19} \text{ cm}^{-3}$. If we consider the Maxwell-Garnett approximation, we can estimate an effective refractive index of 1.89 for the nanocrystal region (see Fig. 3 in the Supplementary Information). For nanocrystals produced by plasma-enhanced chemical vapour deposition (PE-CVD)¹⁷, an effective refractive index of 1.71 was measured by ellipsometry. We note that in sample A, this causes the formation of a planar waveguide with an optical filling factor of about 9.7% when a refractive index of 1.89 is considered, or of 1.17% when a refractive index of 1.71 is assumed (see Fig. 3 in the Supplementary Information).

Absorbance and luminescence spectra at room temperature for sample A are shown in Fig. 1. A single wide emission band peaked at 800 nm, characteristic of the radiative recombination of carriers in Si nanocrystals, is observed. Absorbance measurements revealed a band in the near-infrared and a rising absorption edge at shorter wavelengths. The rising edge is due to absorption in the quantum confined states of the nanocrystals¹⁸, whereas the peculiar feature of the near-infrared absorption band is caused by a Si=O interface state^{10,19,20}. As predicted by theory^{10,21} and inferred from experiment, the interface state is formed at the interface between the Si nanocrystals and the SiO₂ matrix. The microscopic nature of these interface states is still under debate^{10,21}. Very good quality SiO₂ and Si nanocrystals are needed to observe this interface state, which in other Si-based systems is hindered by interfaces with defects or the low quality of the oxide. We note the spectral coincidence of the emission band and the interface state absorption band, suggesting that radiative emission in Si nanocrystals occurs through a radiative state associated with the nanocrystal-oxide interface. Time-resolved luminescence, under picosecond excitation, on our nanocrystals showed a very fast rise time, within our experimental sensitivity (some nanoseconds)²². The decay time of the luminescence was in the microsecond range; it is dependent on the emission energy²³.

Light amplification

To measure light amplification we used the variable strip length method (see the inset of Fig. 2)²⁴. The sample is optically excited by a

doubled Ti:sapphire laser beam ($\lambda = 390 \text{ nm}$, 2-ps pulse width, 82-MHz repetition rate) in a stripe-like geometry with variable length (l). The amplified spontaneous emission intensity I_{ASE} that is emitted from the sample edge (observation angle $\phi = 0$) is measured as a function of l . From a fit of the resulting curve, the optical gain g can be deduced at every wavelength. By assuming a one-dimensional amplifier model, I_{ASE} can be related to g by^{15,24}

$$I_{\text{ASE}}(l) \propto \frac{I_{\text{SPONT}} \times l}{g - \alpha} (e^{(g - \alpha)l} - 1) \quad (1)$$

where I_{SPONT} is the spontaneous emission intensity per unit length and α an overall loss coefficient. The gain measured in this way is the modal gain¹⁵. Figure 2 shows I_{ASE} versus l in Si nanocrystals measured at a wavelength of 800 nm. For small values of l ($< 0.05 \text{ cm}$), an exponential increase of I_{ASE} is observed that indicates the occurrence of amplified spontaneous emission. A fit with equation (1) yields the net modal gain $g - \alpha = 100 \pm 10 \text{ cm}^{-1}$. For large values of l ($> 0.05 \text{ cm}$), I_{ASE} saturates as expected for any finite power supply amplification mechanism. At

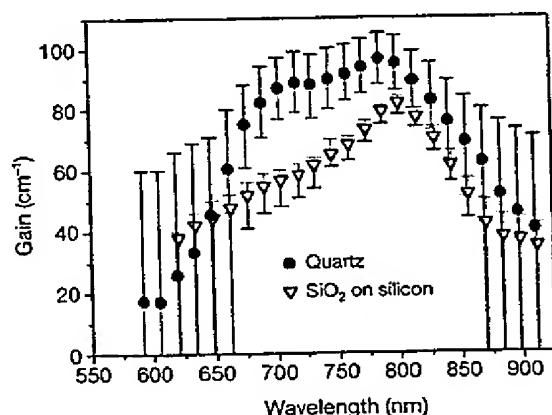


Figure 3 Spectral dependence of the net modal gain. Sample A, circles; sample B, triangles. The experimental conditions were as in Fig. 2. The large error bars result from both the low signal-to-noise ratio of the streak camera detection for low intensity signals and from the numerical procedure used for obtaining the modal gain from the ASE data^{30,31}.

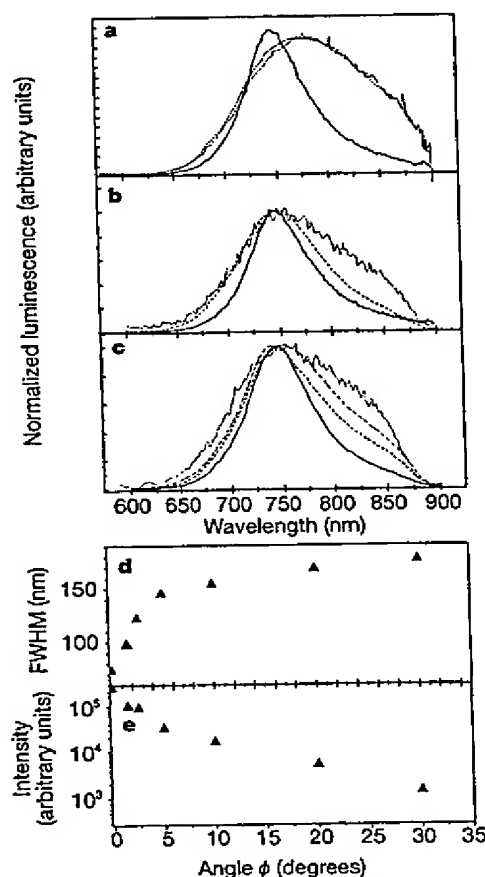


Figure 4 Amplified spontaneous emission spectra of sample A for different measurement conditions. **a**, Amplified spontaneous emission (ASE) spectra for a constant excitation length $l = 2,000 \mu\text{m}$ and various power densities P : continuous line 2.3 kW cm^{-2} , dashed line 1 kW cm^{-2} , dotted line 170 W cm^{-2} . **b**, ASE spectra for constant $P = 2.3 \text{ kW cm}^{-2}$ and various l values: $l = 2,000 \mu\text{m}$ continuous line, $l = 650 \mu\text{m}$ dashed line, $l = 200 \mu\text{m}$ dotted line. **c**, ASE spectra for constant $P = 1 \text{ kW cm}^{-2}$ and $l = 3,000 \mu\text{m}$ and various observation angles ϕ . ϕ is defined with respect to the optical axis of the one-dimensional amplifier. Continuous line $\phi = 0^\circ$, dashed line $\phi = 1.5^\circ$, dash-dotted line $\phi = 2.5^\circ$, dotted line $\phi = 20^\circ$. The low energy cut-off of the photomultiplier used to record the data deforms the spectra for wavelength longer than 880 nm. **d**, Full-width at half-maximum (FWHM) of the ASE emission as a function of the observation angle ϕ . **e**, Peak intensity of the ASE signal as a function of ϕ .

low power density, we measured absorption; when the pump power was increased, the peak net modal gain increased and then saturated at values of about 100 cm^{-1} for power densities of about 5 kW cm^{-2} .

By measuring the amplified signal for various wavelengths we obtained the gain spectrum for both samples A and B (Fig. 3; see also Supplementary Information). A wide spectral band is observed which spectrally overlaps the wavelength range of the luminescence, demonstrating that amplification is produced by the radiative state associated with the nanocrystal-oxide interface. We noticed that both samples yielded similar shapes and values for the gain curve. A confirmation of these findings was the observation of a strong emission lineshape narrowing (Fig. 4), either when the pump power density P is increased with fixed excitation length l (Fig. 4a), or when the excitation length l is increased with a fixed P (Fig. 4b). When l and P are fixed and the observation angle ϕ is changed (Fig. 4c–e), a significant intensity decrease and a broadening of the amplified emission spectrum occur as soon as there is deviation from the strict one-dimensional amplifier configuration, that is, when $\phi > 0^\circ$. These observations support also the waveguide formation in our samples.

The most direct evidence of light amplification from our systems was provided by pump and probe transmission measurements. An intense laser beam (pump) at 390 nm excites the sample in order to reach the population inversion needed for amplification, while a weak probe signal at $\sim 800 \text{ nm}$ passes through the active layer of thickness d . In the presence (absence) of the pump beam the probe beam is amplified (absorbed). In Fig. 5 we show the results. The probe signal is clearly amplified when passing through the excited nanocrystals. To our knowledge, this is the first evidence of light amplification in transmission, usually named single-pass gain, in Si-based systems. We deduced the net material gain values by using the formula given in Fig. 5; they are high enough to compare with those of self-assembled quantum dots made of III–V semiconductors^{25,26}: $10,000 \pm 3,000 \text{ cm}^{-1}$. This value has a very large error bar because of the geometry of the active nanocrystal layer and the losses in the quartz substrate. No change in probe intensity in the presence/absence of the pump beam was observed when the probe beam passed through pure quartz (without nanocrystals). Moreover, by decreasing the pump intensity (Fig. 5, right panel) we measured even absorption of the probe beam (population inversion is no

longer reached in nanocrystals). By changing the probe wavelength the net material gain decreased and eventually disappeared (amplification is lost when the probe energy is no longer in resonance with the transition for which population inversion is achieved) with an overall spectral dependence similar to that shown in Fig. 3 (see also Supplementary Information).

Gain cross-section per nanocrystal

By using the formalism of ref. 27 and the measured probe beam transmission under inversion conditions, we estimated a maximum-gain cross-section per nanocrystal $\gamma_T \approx 5 \times 10^{-16} \text{ cm}^2$. It is interesting to compare this gain cross-section per nanocrystal with the photon absorption cross-section per nanocrystal (σ). We have hence directly measured the absorption cross-sections of ion-implanted Si nanocrystals by studying the rise time of the photoluminescence intensity as a function of pump power in a fashion similar to that recently reported for porous Si²⁸.

The photoluminescence intensity is given by $I \propto N^*/\tau_R$, where N^* is the concentration of excited nanocrystals and τ_R the radiative lifetime. The rate equation for nanocrystal excitation will be:

$$\frac{dN^*}{dt} = \sigma J(N - N^*) - \frac{N^*}{\tau} \quad (2)$$

where J is the photon flux, N is the concentration of nanocrystals and τ is the decay time, taking into account both radiative and non-radiative processes. If a continuous wave (CW) pumping laser is turned on at $t = 0$, N^* will change according to equation (2) and the photoluminescence intensity will increase according to the following law:

$$I(t) = I_0 \left\{ 1 - \exp \left[- \left(\sigma J + \frac{1}{\tau} \right) t \right] \right\} = I_0 \left\{ 1 - \exp \left[- \left(\frac{t}{\tau_{\text{on}}} \right) \right] \right\} \quad (3)$$

A measure of the photoluminescence rise time as a function of J will therefore give direct information on the absorption cross-section. The inset to Fig. 6 shows $I(t)$ at 850 nm for Si nanocrystals pumped at 488 nm for different pump powers P . As predicted by equation (3), the photoluminescence rise time becomes shorter and shorter

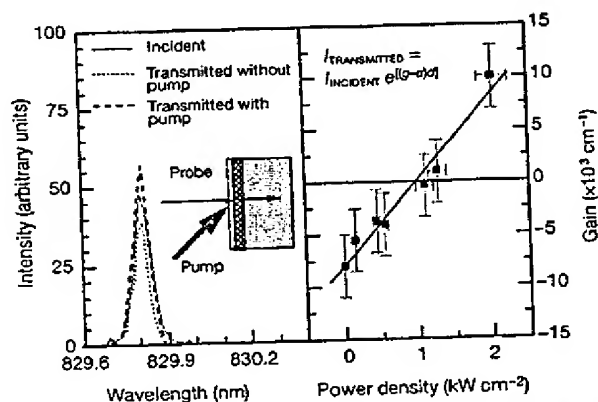


Figure 5 Gain measurements. Left panel, spectrum of the transmitted probe beam measured in presence (dashed line) or in absence (dotted line) of the pump beam. The spectrum of the transmitted probe beam in absence of the absorbing/amplifying nanocrystal medium is also reported (continuous line) and named the incident beam. The inset shows the principle of the experiment. The formula used to deduce the material gain is also reported, d being the thickness of the active region. The probe beam was provided by a Kr lamp, which was imaged to a spot size of about 0.01 mm^2 on the sample surface. The pump beam had a mean power of about 2 kW cm^{-2} and a wavelength of 390 nm . Right panel, dependence of the material gain value on the pump power density.

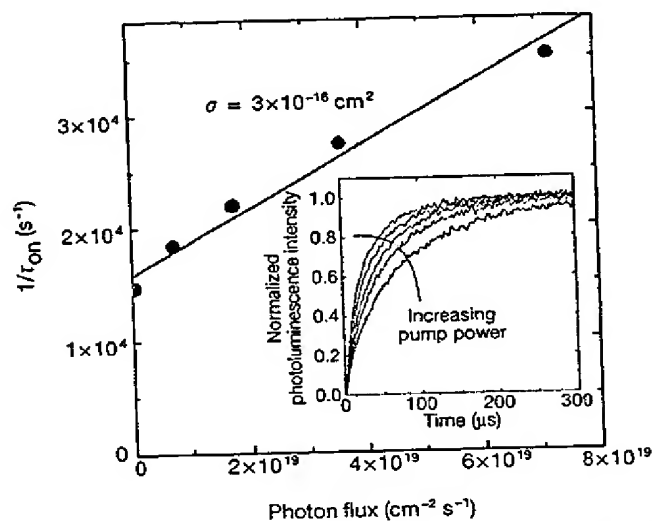


Figure 6 Reciprocal of the rise time τ_{on} as a function of the pump laser photon flux as obtained from a fit to the time resolved photoluminescence data shown in the inset. The slope gives the photon absorption cross-section σ . Inset: time resolved photoluminescence intensity at 850 nm switching on the 488 nm line of a CW Ar pumping laser at $t = 0$; data are taken at room temperature, at different pump powers (in the 0.8 – 80 mW range) and are normalized to the maximum intensity.

as P is increased. By fitting these curves with equation (3) we obtain the values of the rise time, τ_{on} , at the different P values. The reciprocal of τ_{on} is reported in Fig. 6 as a function of J . The data follow a straight line with a slope $\sigma \approx 3 \times 10^{-16} \text{ cm}^2$. The intercept of the fitted straight line with the vertical axis gives the lifetime of the Si nanocrystals in the system at the measured wavelength. The obtained value (70 μs) is in agreement with decay time measurements at 850 nm on the same sample. In this way we have been able to obtain a direct measurement of the photon absorption cross-section of the nanocrystals. Although this measurement is performed at an excitation wavelength of 488 nm, it should reflect the property of the 800 nm state because absorbance at these two wavelengths is identical (see Fig. 1). We note that, as theoretically predicted¹⁵, the measured absorption cross-section σ is of the same order of magnitude as the gain cross-section γ_T . The same agreement is found when we compare the net material gain to the absorption coefficient deduced by the absorbance data of Fig. 1.

Another important issue concerns the comparison of the gain cross-sections that are derived from the modal and the material gain. It was shown in ref. 27 that in the variable stripe-length geometry, the gain cross-section per nanocrystal (γ_{ASE}) can be derived by using

$$\gamma_{ASE} = \frac{g}{(f_c - f_v)N\Gamma} \quad (4)$$

where Γ is the optical filling factor of the amplified mode. By assuming a complete population inversion $f_c - f_v = 1$, an optical filling factor of 0.097 and the measured net modal gain $g \approx 100 \text{ cm}^{-1}$, one finds $\gamma_{ASE} \approx 5 \times 10^{-17} \text{ cm}^2$. This is a lower limit to γ_{ASE} . If we consider an incomplete inversion or a weaker confining waveguide (that is, no step index profile in the waveguide and/or a lower effective refractive index for the nanocrystal), γ_{ASE} will significantly increase. For example, by using the Γ value computed for the refractive index measured in PE-CVD nanocrystals¹⁷, we obtain $\gamma_{ASE} \approx 3 \times 10^{-16} \text{ cm}^2$. The difference between the two estimated values of γ_{ASE} may be an indication of the quantitative uncertainty on the gain cross-section values determined by the gain measurements.

A comparison of the gain cross-sections per nanocrystal derived by transmission (γ_T) and by the variable strip length method (γ_{ASE}) shows that the two values are in reasonable quantitative agreement. Indeed, the determined value of γ_{ASE} is only a lower limit and there

are large error bars on the pump and probe determined gain coefficients.

Origin of gain

The wide spectral gain of inhomogeneous nature is energetically matching both the luminescence emission and the 800-nm interface state absorption band. For these reasons, a three-level model is proposed to explain the observed gain (Fig. 7): two levels correspond to the lowest unoccupied molecular orbital (LUMO), or the bottom of the conduction band, and to the highest occupied molecular orbital (HOMO), or the top of the valence band of the nanocrystal, respectively. The third level is due to the radiative interface state observed in absorption and responsible for the luminescence emission band at 800 nm. Optical excitation populates the LUMO, emptying the HOMO. Electrons in the LUMO relax very rapidly to the interface state. Electrons in the interface state have long lifetimes. Indeed, the absorption band at 800 nm, the Stokes shift between absorption and luminescence, the fast rise and the slow decay times of the 800-nm luminescence under picosecond excitation, and the efficient luminescence emission of the 800-nm luminescence all support this energy model. Within the model, the rate of depopulation of the initial state is much faster than its filling rate via a carrier recombination mediated by the interface state. Population inversion between the HOMO and the radiative state associated with the nanocrystal-oxide interface is thus possible. This model also explains why losses due to free carrier absorption that usually exceed the gain by stimulated emission in other Si based systems²⁹, or those due to Auger recombinations²², are not effective here. In addition, model calculations¹⁰ show that the size dependence of the radiative interface state energy is smaller than that of conduction-to-valence band transitions, relaxing issues related to the broad distribution of sizes.

Using the measured absorption cross-section σ per nanocrystal, we estimate that under our peak excitation condition of about 10^{22} photons $\text{cm}^{-2} \text{ s}^{-1}$, more than 100 electron-hole pairs per nanocrystal are generated. As we have nanocrystals with about 500 Si atoms, of which about 35% are surface atoms, we have about 150 interface states available per Si nanocrystal when we assume that each surface Si atom is bound to an O atom. In Fig. 5, we show that to have optical gain the excitation level should be high enough to invert most of these states.

In Table 1 we report a compilation of data on the gain cross-section per quantum dot for some III-V semiconductors. It can be noticed that the silicon nanocrystal values are about three orders of magnitude lower than the one typically found in InAs quantum dots. We argue that this is due to the indirect bandgap of Si and to the fact that the gain is due to radiative interface states. Despite this difference, the net material gain is of the same order of magnitude between Si nanocrystals and InAs quantum dot systems, owing to the much higher areal density of nanocrystals that is achievable with the ion-implantation method used in this work. We note that the gain cross-section per nanocrystal γ is inversely proportional to the radiative lifetime τ_R (ref. 15). By looking at the γ data in Table 1, we infer large differences in lifetimes. Indeed radiative lifetimes in the microsecond range are measured for Si nanocrystals, whereas for InAs quantum dot lifetimes in the nanosecond range were reported^{25,26}.

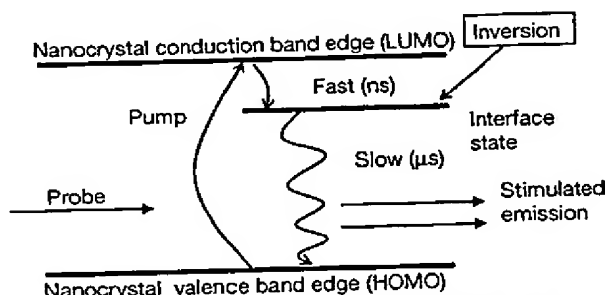


Figure 7 Schematic energy diagram for a nanocrystal showing how population inversion can be reached in this system.

Table 1 Gain cross-section per quantum dot or nanocrystal

Quantum dot material	Net modal gain (cm^{-1})	Net material gain ($\times 10^4 \text{ cm}^{-1}$)	Areal dot density (cm^{-2})	Active layer thickness (nm)	Filling factor (10^{-4})	Gain cross-section per dot (10^{-16} cm^2)	Reference
InAs single layer quantum dot	8.2	9*	1×10^{11}	1.7	1.2	1,200	25
InAs 7 stacks quantum dot	70–85	1.5*	1×10^{11}	100	48	4,000	26
GaAs single layer quantum dot	13*		1×10^{10}			450*	27
Si nanocrystals	100	1	2×10^{14}	100	970	0.5–5	This work

* Calculated approximately.

Conclusions

Modal and net material optical gains have been observed unambiguously in Si nanocrystals. Quantitative estimates of gain cross-section per nanocrystal show that the measured values are orders of magnitude lower than those found in III-V semiconductor quantum dots. However, owing to the much higher stacking density of Si nanocrystals with respect to direct-bandgap quantum dots, similar values for the material gain are observed. These findings open a route towards the realization of a silicon-based laser. □

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TOPICAL REVIEW

Will silicon be the photonic material of the third millenium?*

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Abstract

Silicon microphotronics, a technology which merges photonics and silicon microelectronic components, is rapidly evolving. Many different fields of application are emerging: transceiver modules for optical communication systems, optical bus systems for ULSI circuits, I/O stages for SOC, displays, In this review I will give a brief motivation for silicon microphotronics and try to give the state-of-the-art of this technology. The ingredient still lacking is the silicon laser: a review of the various approaches will be presented. Finally, I will try to draw some conclusions where silicon is predicted to be the material to achieve a full integration of electronic and optical devices.

(Some figures in this article are in colour only in the electronic version)

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* This review is based on the books *Light Emitting Silicon for Microphotronics* by S Ossicini, L Pavesi and F Priolo (*Springer Tracts in Modern Physics*), at press, and *Towards the First Silicon Laser* edited by L Pavesi, S Gaponenko and L Dal Negro (*NATO Science Series II*) vol 93 (Dordrecht: Kluwer), 2003.

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1. Why silicon photonics?

The big success of today's microelectronic industry is based on various factors, among others

- the presence of a single material, silicon, which is widely available, can be purified to an unprecedented level, is easy to handle and to manufacture and shows very good thermal and mechanical properties which render the processing of devices based on it easy [1],
- the availability of a natural oxide of silicon, SiO_2 , which effectively passivates the surface of silicon, is an excellent insulator, is an effective diffusion barrier and has a very high etching selectivity with respect to Si,
- the presence of a single dominating processing technology, CMOS, which accounts for more than 95% of the whole market of semiconductor chips [2],
- the possibility to integrate more and more devices, 55 000 000 transistors in PENTIUM® 4 (figure 1), on larger and larger wafers (300 mm process and 400 mm research) with a single transistor size which is decreasing (gate lengths of 180 nm are in production while 15 nm have been demonstrated) [3], yielding a significant reduction in cost per bit,
- the ability of the silicon industry to face improvements when the technology is hitting the so-called red brick wall, e.g. the use of SiGe for high frequency operation and the introduction of low k -materials and of Cu to reduce RC delays,
- an accepted common roadmap which is dictating the technology evolution for processes, architectures or equipment [3] and
- the presence of big companies which define standards and trends (almost 90% of the market is shared by ten companies).

All these factors have rendered the microelectronics industry very successful. However, in recent years some concerns about the evolution of this industry have been raised which seem related to fundamental materials and processing aspects [4]. An important example is related to the limitations of the operating speed of microelectronic devices due to the interconnect [5]. Figure 2 shows the signal delay as a function of the generation of transistors [6]. For gate length shorter than 200 nm, a situation is reached where the delay is no longer dictated by the gate switching time but by the wiring delay. In addition, as the integration is progressing the length of the interconnects on a single chip is getting longer and longer. Nowadays chips have total interconnection length per unit area of the chip of some 5 km cm^{-2} with a chip area of 450 mm^2 while in ten years from now these lengths will become 20 km cm^{-2} for a chip area of 800 mm^2 . The problem is not only related to the length of the interconnects but also to the complexity of their architecture. Nowadays, there are six layers of metal levels (figure 3), while in ten years from now there will be more than 12. All these facts introduce problems related to the delay in signal propagation causing RC coupling, signal latency, signal cross-talk and RL delays due to the reduction in dimension and increase in density of the metal line. A possible solution to these problems is looked for in optics [7]: the use of optical interconnects. Nowadays, optical interconnects through optical fibres and III-V laser sources are already used to connect different computers. It is predicted that optical interconnects will be used to connect computer boards in five years, while the use of optical interconnects within the chip is being investigated and will possibly be realized in 10–15 years from now [8]. Optical interconnects are one of the main motivations to look for silicon photonics. But this is not the only one. Photonics has seen a big development in recent years at the request of the communication market, where more and

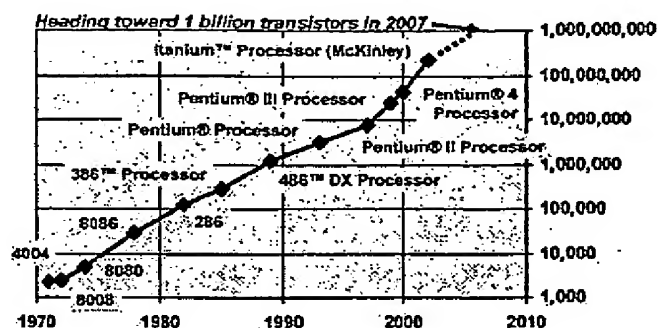


Figure 1. Evolution of the number of transistors in a single CPU (central processing unit) versus the year. This graph is based on the Intel CPU [6].

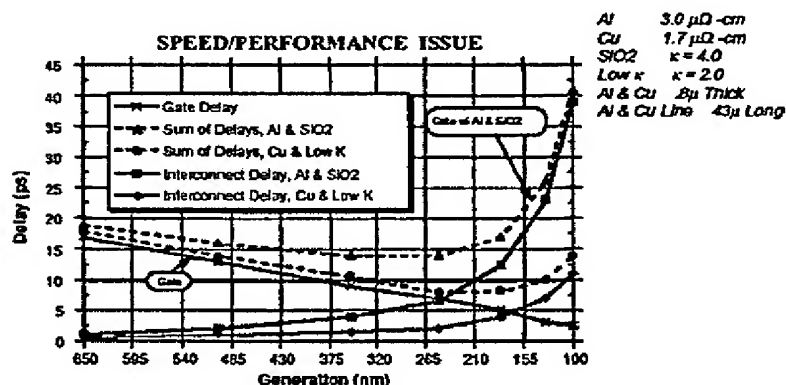


Figure 2. Calculated gate delay and wire delay as a function of the minimum feature size (device generation). From SIA Roadmap 1997 [3]. Interconnections and signal integrity, DAC tutorial. 38th Design Automation Conf. ©2001 (www.amanogawa.com/epep2000/files/jose1.pdf).

more information has to be sent at higher and higher speed. Nowadays, the capacity of optical communication on long hauls is reaching some Tb/s^{-1} over thousands of kilometres. And all these are thanks to the progress in optical fibre fabrication, the use of DWDM, of EDFA and Raman amplifiers, modulators and single frequency lasers.

If one compares the photonic industry with microelectronics today one can see many differences.

- (1) A variety of different materials is used: InP as substrate for source development, silica as material for fibres, lithium niobate for modulators, other materials for DWDM and EDFA and so on.
- (2) No single material or single technology is leading the market. Some convergence is appearing towards the use of InP as the substrate material to integrate different optical functions.
- (3) The industry is characterized by many different small companies which are specialized in specific devices: lasers, modulators etc. No big companies are dominating at present.
- (4) The production technology is still very primitive. Chip scale integration of optical components, which enables low cost and high reproducibility, is not yet achieved. Neither

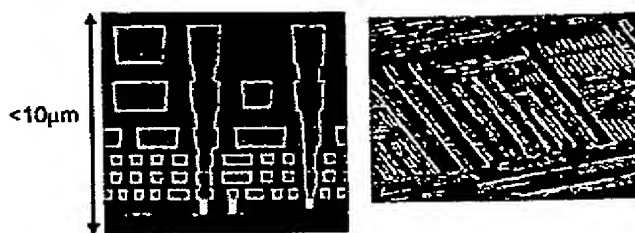


Figure 3. An example of the complexity of the metal interconnects in today's chip. Left chip cross-section: most of the chip is occupied by metal interconnect layers. Right: the complexity of the architecture of the metal line. From a talk by Joise Maiz at the Spanish Microsystems Research Centre (CMIC) on 14 June 2002 (<http://www.intel.com/research/silicon/CMIC.2002.Jose.Maiz.htm>).

standardization of processes nor packaging of optical components, which is inherent for mass production and repeatability, are present.

- (5) Roadmaps to dictate and forecast the evolution of photonics are only now being elaborated [9].

It is commonly accepted that the industrial model of microelectronics if applied to photonics will be a booster to the development and implementation of photonics. To describe this new technology the term of microphotonics has been proposed [11]. All the big players of microelectronics have aggressive programmes to develop microphotonics, mostly based on silicon [10].

The aim of this review is to try to give the state-of-the-art on the development of silicon photonics with the aim of settling the status and trying to weigh up whether silicon can be used as the photonics material. For this reason, all the different components are briefly reviewed (section 2) with a special emphasis on the subject which is at the forefront of today's discussion: the route to a silicon laser (section 3). The selection of the various experimental data is not intended to be exhaustive but simply representative of some of the more successful devices and integration schemes which have been reported. I apologize in advance to all those authors whose work I am not referring to.

2. Silicon photonics

It was predicted in the early 1990s that silicon based optoelectronics would be a reality before the end of the century [12, 13]. Indeed, all the basic components have already been demonstrated [14], except for a silicon laser.

2.1. Silicon based waveguides

The first essential component in silicon microphotonics is the medium through which light propagates: the waveguide. This has to be silicon compatible and should withstand normal microelectronics processing. Critical parameters are the refractive index of the core material, its electro-optical effects, the optical losses and the transparency region. To realize low loss optical waveguides, various approaches have been followed [15]: low dielectric mismatch structures (e.g. doped silica [16], silicon nitride [17] or silicon oxynitride on oxide [18], or differently doped silicon [19]) or high dielectric mismatch structures (e.g. silicon on oxide [11]). Low loss silica waveguides are characterized by large dimensions (see figure 4), typically 50 μm of thickness, due to the low refractive index mismatch ($\Delta n = 0.1\text{--}0.75\%$). Silica waveguides

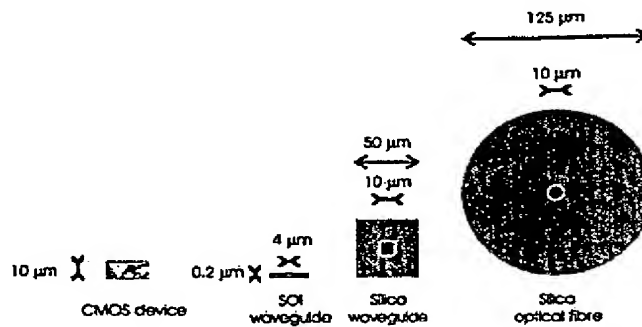


Figure 4. Comparison of the cross-sections of a CMOS chip, a typical SOI waveguide, a typical silica waveguide and a silica mono-mode optical fibre.

have a large mode spatial extent and, thus, are interesting for coupling with optical fibres but not for integration into/within electronic circuits because of a significant difference in sizes. The large waveguide size also prevents the integration of a large number of optical components in a single chip. Similar problems exist for silicon on silicon waveguides where the index difference is obtained by varying the doping density [19]. Silicon on silicon waveguides are very effective for realizing free-carrier injection active devices (e.g. modulators) as well as fast thermo-optic switches thanks to the high thermal conductivity of silicon. A major problem with these waveguides is the large free-carrier absorption which causes optical losses of some dB cm^{-1} for single-mode waveguides at $1.55 \mu\text{m}$. Silicon nitride based waveguides [17] and silicon oxynitride waveguides [18] show losses at 633 nm lower than 0.5 dB^{-1} and bending radii of less than $200 \mu\text{m}$. The nitride based waveguides are extremely flexible with respect to the wavelength of the signal light: both visible and IR.

At the other extreme, silicon on insulator (SOI) or polysilicon based waveguides allow for a large refractive index mismatch and, hence, for small size waveguides in the sub-micrometre range. This allows a large number of optical components to be integrated within a small area. Optical losses as low as 0.1 dB cm^{-1} at $1.55 \mu\text{m}$ have been reported for channel waveguides in SOI (optical mode cross-section $0.2 \times 4 \mu\text{m}^2$) [20]. Ideal for on-chip transmission, SOI waveguides have coupling problems with silica optical fibre due to both the large size difference and the different optical impedance of the two systems (figure 4). Various techniques have been proposed to solve these problems, among which are adiabatic tapers, V-grooves and grating couplers (figure 5) [21, 22]. Large single-mode stripe loaded waveguides on SOI can be achieved provided that the stripe and the slab are both made of silicon [23]. This SOI system provides low loss waveguides ($<0.2 \text{ dB cm}^{-1}$) with single-mode operation with large rib structures (optical mode cross-section $4.5 \times 4 \mu\text{m}^2$) and low birefringence ($<10^{-3}$). Appropriate geometry with the use of an asymmetric waveguide allows bend radii as short as 0.1 mm [24]. A number of photonic components in SOI have been demonstrated [23] and commercialized [24]: directional couplers, dense WDM arrayed waveguide grating, Mach-Zehnder filters, star couplers, ...

2.2. Detectors

The optical signal is converted into an electrical signal by using silicon based photodetectors. Detectors for silicon photonics are based on three different approaches [25]: silicon photoreceivers for $\lambda < 1.1 \mu\text{m}$, hybrid systems (mostly III-V on Si) and heterostructure based systems. High speed (up to 8 Gb s^{-1}) monolithically integrated silicon photoreceivers

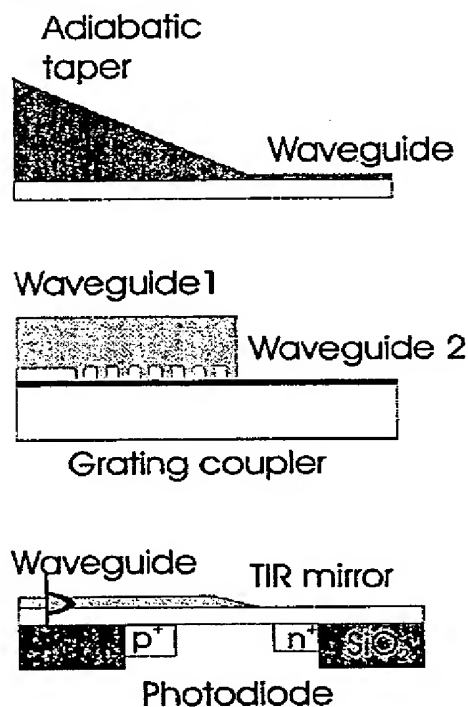


Figure 5. Various schemes to couple the light from a fibre into a waveguide by using an adiabatic taper or a grating coupler, or from a waveguide into a photodiode by using a curved TIR (total internal reflection) mirror.

at 850 nm have been fabricated by using 130 nm CMOS technology on a SOI wafer [26]. Other recent results confirm the ability of silicon integrated photoreceivers to detect signals with a high responsivity of 0.46 A W^{-1} at 3.3 V for 845 nm light and 2.5 Gb s^{-1} data rate [27]. The heterostructure approach is mainly based on the heterogrowth of Ge rich SiGe alloys: Ge-on-Si photodetectors have been reported with a responsivity of 0.89 A W^{-1} at $1.3 \mu\text{m}$ and 50 ps response time [28]. 1% quantum efficiency at $1.55 \mu\text{m}$ in an MSM (metal–semiconductor–metal) detector based on a Si/SiGe superlattice shows that promising developments are possible [29]. Similarly a waveguide photodetector with Ge/Si self-assembled islands shows responsivities of 0.25 mW at $1.55 \mu\text{m}$ with zero bias [30].

2.3. Other photonics components

Almost all the other photonics components have been demonstrated in silicon microphotronics [13, 25]. Optical modulators, optical routers and optical switching systems have been all integrated into silicon waveguides [31]. Discussion of a series of photonics components realized with SOI waveguides is given in [23] which includes plasma dispersion effect based active gratings, evanescent waveguide coupled silicon–germanium based photodetectors and Bragg cavity resonant photodetectors.

2.4. Silicon photonic integrated circuits

Based on the technologies reported in the previous sections, various demonstrations of photonic integrated circuits based on silicon have been reported. Here we discuss some examples.

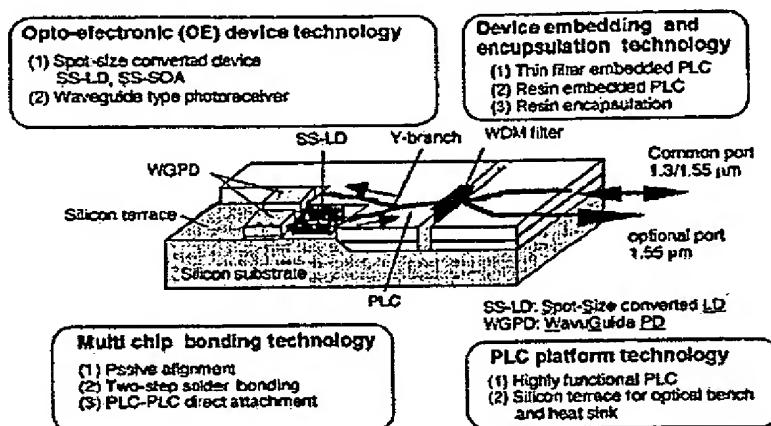


Figure 6. Example of the various devices that can be integrated on a silica based lightwave circuit. SS-LD stands for laser diode, WGPD stands for photodetectors (from [32]).

Hybrid integration of active components and silica-based planar lightwave circuits provides a full scheme for photonic component integration within a chip [32]. Passive components are realized by using silica waveguides while active components are hybridized within the silica (see figure 6). Active components (laser diodes, semiconductor optical amplifiers and photodiodes) are flip-chip bonded on silicon terraces where the optical waveguides are also formed. By using this approach, various photonic components have been integrated such as multi-wavelength light sources, optical wavelength selectors, wavelength converters, all optical time-division multiplexers etc [32]. Foreseen applications are WDM transceiver modules for fibre-to-the-home application.

A full integrated optical system based on silicon oxynitride waveguides, silicon photodetectors and CMOS transimpedance amplifiers has been realized [18]. Coupling of visible radiation to a silicon photodetector can be achieved by using mirrors at the end of the waveguide (figure 5). These are obtained by etching the end of the waveguide with an angle so that the light is reflected at almost 90° into the underlying photodetector. A schematic diagram of the cross-section of the device is shown in figure 7.

Commercial systems for the access network telecom market have been realized by using SOI waveguides and the silicon optical bench approach to interface the waveguides with both III-V laser sources and III-V photodetectors. The silicon optical bench (SOB) is a technology where the silicon wafer is used as a substrate (*optical bench*) where the various optical components are inserted by micromachining suitable lodging. In [24], lasers and photodetectors are stuck into etched holes in silicon and bump soldered in place. The system operates at $1.55 \mu\text{m}$ with a typical bit rate of 155 Mb s^{-1} [24]. A further advantage of the use of a large optical mode waveguide is the ease of interfacing to single-mode optical fibre. In the approach of [24], these are located in V-grooves etched into silicon.

A fully integrated system working at $1.55 \mu\text{m}$ has been demonstrated based on silicon waveguides with very small optical mode (cross-section $0.5 \times 0.2 \mu\text{m}^2$) which allows extremely small turn radii ($1 \mu\text{m}$) [11]. In this way a large number of optical components can be integrated on a small surface ($\approx 10\,000$ components cm^{-2}). Detectors are integrated within silicon by using Ge hetero-growth on silicon itself. Responsivity of 250 mA W^{-1} at $1.55 \mu\text{m}$ and response times shorter than 0.8 ns have been achieved [28]. A scheme for an optical clock distribution

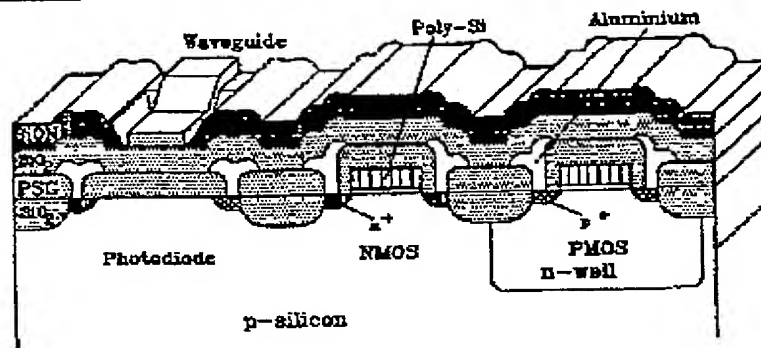


Figure 7. Cross section of an integrated device with a photodiode (PD) (left), the waveguide coupled to the PD by the TIR mirror and an amplifier stage realized with CMOS technology (from [18]).

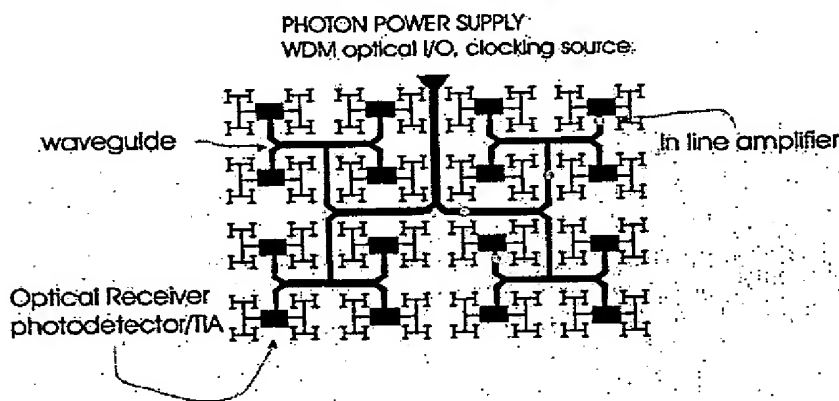


Figure 8. Scheme for an integrated optical circuit to distribute the clock signal on a chip (from [33]).

within integrated circuits based on this approach is shown in figure 8 [33]. Here the laser source is external to the chip and acts as a *photon battery* similarly to usual batteries for electrons.

A realistic bidirectional optical bus architecture for clock distribution on a Cray T-90 supercomputer board based on polyimide waveguides (loss of 0.21 dB cm^{-1} at 850 nm), a GaAs VCSEL and silicon MSM photodetectors has been investigated [34]. By using 45° TIR (total internal reflection) mirror coupling efficiencies as high as 100% among the sources or the detectors and the waveguides have been demonstrated. Examples of the connection scheme are shown in figure 9.

3. Silicon laser

To achieve monolithically integrated silicon microphotronics, the main limitation is the lack of any practical Si-based light sources: either efficient light emitting diodes (LEDs) or Si lasers. A laser is preferred as incoherent emission is probably not sufficient for dense, high speed interconnects mostly because of the basic optical inefficiencies in focusing incoherent light. A laser is ideal for optical interconnects, or more generally speaking, for silicon

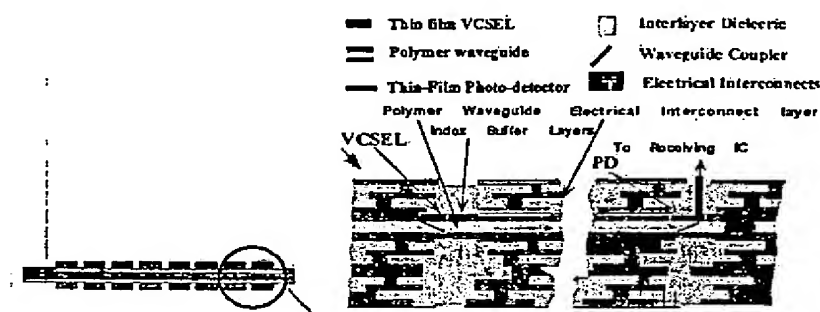


Figure 9. The optical interconnect scheme proposed in [34] for a supercomputer board: left, schematic diagram of the side view of the vertical integration layers; right, details of the schematic diagram (from [34]).

microphotonics. Unfortunately, today, the only viable solution is the hybrid approach where III-V semiconductor lasers are grown, bonded or connected to silicon photonic integrated circuits. To have a silicon laser, or in general a laser, one needs three key ingredients:

- (i) an active material which should be luminescent in the region of interest and which should be also able to amplify light,
- (ii) an optical cavity into which the active material should be placed to provide the positive optical feedback and
- (iii) a suitable and efficient pumping scheme to achieve and sustain the laser action; for integration purposes the pumping mechanism is preferable via electrical injection.

Silicon is an indirect bandgap material; light emission is a phonon-mediated process with low probability (spontaneous recombination lifetimes in the milliseconds range) [35]. In standard bulk silicon, competitive non-radiative recombination rates are much higher than the radiative ones and most of the excited e-h pairs recombine non-radiatively. This yields very low internal quantum efficiency ($\eta_i \approx 10^{-6}$) for bulk silicon luminescence. In addition, fast non-radiative processes such as Auger or free-carrier absorption severely prevent population inversion for silicon optical transitions at the high pumping rates needed to achieve optical amplification. Despite all this, during the 1990s many different strategies have been employed to overcome these materials limitations [35]. The most successful ones are based on the exploitation of low dimensional silicon where silicon is nanostructured and hence the electronic properties of free carriers are modified by quantum confinement effects [13]. A steady improvement in silicon LED performances has been achieved and silicon LEDs are now within the strict market requirements [36]. In addition, many breakthroughs have been recently demonstrated showing that this field is very active and still promising [36–40]. Figure 10 shows a schematic sketch of the various strategies that are currently followed to build a silicon laser [41]. They differ both for spectral region of emission and for the physics behind. In the following, I will review all these approaches and try to weigh them up.

3.1. Bulk silicon

Silicon is an indirect bandgap material, thus the probability for a radiative transition is very low. This is reflected in very long times for radiative recombinations. Due to these long radiative lifetimes, excited free carriers have large probabilities of finding non-radiative recombination

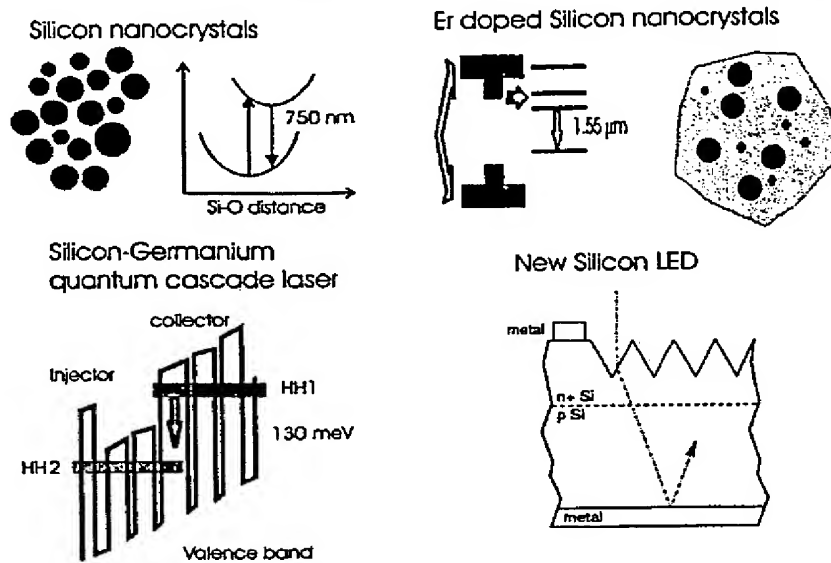


Figure 10. Various approaches proposed to realize a silicon laser.

centres and recombining non-radiatively. Room temperature emission in bulk silicon with high efficiency has only been observed in ultra-pure silicon with the surface passivated by a native oxide where excited carrier lifetimes are dominated by radiative recombination. Extremely slow recombination rates are possible with high efficiency if one is able to reduce to a minimum the competing non-radiative recombinations. This idea to increase the quantum efficiency of Si has been followed by two different approaches to develop Si based light emitting diodes [36, 42].

The first approach is based on the results achieved in high efficiency solar cells and on the consideration that, within thermodynamic arguments, absorption and emission are two reciprocal processes [36]. At first the non-radiative rates are reduced by using

- (1) high-quality intrinsic Si substrates, float zone (FZ) being preferred over Czochralski (CZ),
- (2) passivation of surfaces by high quality thermal oxide, in order to reduce surface recombination,
- (3) small metal areas and
- (4) limiting the high doping regions to contact areas, in order to reduce the Shockley–Read–Hall recombinations in the junction region.

Then, the parasitic absorption of photons once they have been generated is reduced to a minimum. For example, the reabsorption can be minimized by keeping the doping level to moderate values, such as $\sim 1.4 \times 10^{16} \text{ cm}^{-3}$. Finally, the extraction efficiency of light from bulk silicon can be enhanced by suitably texturizing the Si surface. The final device structure is shown in figure 11. Green *et al* [36] report the highest power efficiency to date for Si based LEDs, approaching 1%. Electroluminescence (EL) spectra of these devices (figure 12) are typical for band-to-band recombinations in silicon. In addition, a fully integrated opto-coupler device (LED coupled to a photodetector) was also demonstrated on the basis of this technology [43].

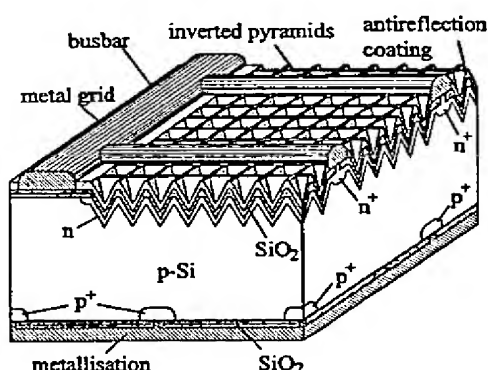


Figure 11. Design of the textured Si light emitting device after [36].

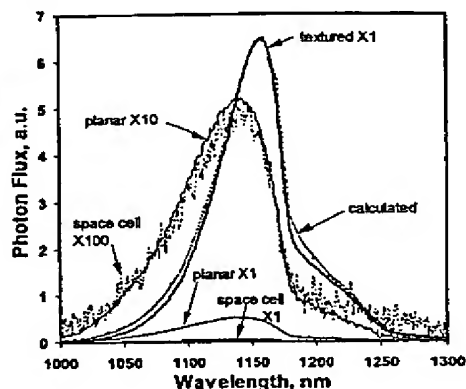


Figure 12. EL spectra for textured, planar and baseline space cell diodes under 130 mA bias current at 298 K (diode area 4 cm²). Calculated values assume a rear reflectance of 96% (after [36]).

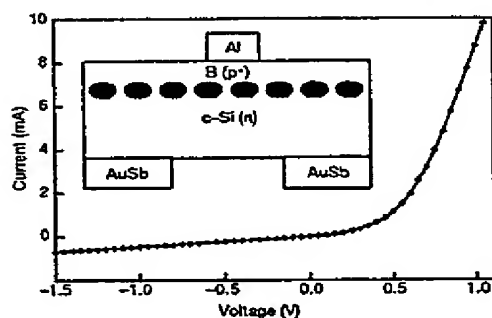


Figure 13. Current-voltage characteristics for the dislocation loop LED measured at room temperature. Inset: a schematic diagram of the LED where the grey circles evidence the region rich in dislocation loops (after [42]).

The main drawbacks of this approach for an integrated laser or light emitting diode are the following:

- (i) the need for both high purity (low doping concentration) and surface texturing renders the device processing not compatible with standard CMOS processing;
- (ii) the strong and fast free-carrier absorption typical of bulk Si, that can prohibit reaching the condition for population inversion, is not addressed [44];
- (iii) the suitable integration of the active bulk Si into an optical cavity to achieve the required optical feedback to sustain a laser action can be a problem;
- (iv) the modulation speed of the device can be limited by the long lifetime of the excited carriers (milliseconds) and by the need for a large optical cavity.

A somewhat different approach was reported in [42]; see figure 13. The idea was again a reduction of the non-radiative channels by exploiting the strain produced by localized dislocation loops to form energy barriers for carrier diffusion. Dislocations form potential pockets close to the junction which block the carriers and enhance radiative decay by localizing

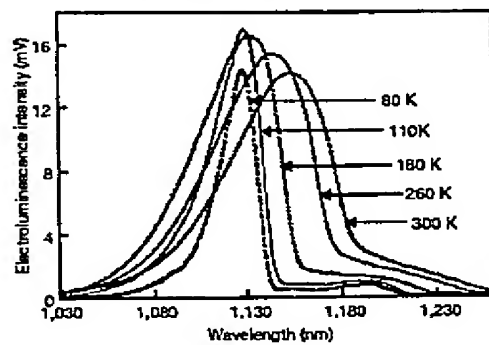


Figure 14. EL spectra against wavelength at various temperatures. The device was operated at a forward current of 50 mA for all temperatures (after [42]).

them in defect-free regions. The size of dislocation loops was in the range of 100 nm, i.e. not enough to cause a quantum confinement of the carriers, and the loop distances were of the order of 20 nm. Free carriers injected through the top electrode are not able to diffuse away and then are constrained to recombine in the near junction region. The onset of the EL at the band edge was observed as the diode turns on under forward bias. No EL was observed under reverse bias. An ultimate external quantum efficiency of about 1% is claimed for these LEDs. The EL spectrum does not present significant differences in lineshape and peak position compared to that of bulk Si (figure 14). A remarkable feature of this device is the high injection efficiency into the confined regions. This is due to the lack of quantum effects. In fact, since the density of states in the active zone is large (comparable to the bulk value), it is not a limiting factor for the free-carrier injection, in contrast to quantum confined structures. On the other hand, injection is also smooth because there is no wide bandgap material as confining barrier. Although not explained, this device has the additional and interesting feature of increasing the efficiency with temperature. The positive role of dislocation loops in enhancing luminescence from near surface silicon has been further confirmed by other authors [45, 46]. The main problem of this approach for a silicon laser is that it does not remove the two main problems of silicon which prevent population inversion, i.e. Auger recombination and free-carrier absorption [44].

Finally, a problem is also related to the wavelength of emission of these bulk silicon LEDs which is resonant with the silicon bandgap: that means that it is very difficult to control the region where the light is channelled in silicon if one wants to use these LEDs as a source for optical interconnects. Light will propagate through the wafer and will be absorbed in unwanted places.

3.2. Silicon nanocrystals

Another way to increase the emission efficiency of silicon is to turn it into a low dimensional material and, hence, to exploit quantum confinement effects to increase the radiative probability of carriers. This approach has been pioneered by the work on porous silicon (PS) [49] which shows that when silicon is partially etched in an HF solution via an electrochemical attack, the surviving structure is formed by small nanocrystals or nanowires which show bright red luminescence at room temperature. The explanation of the observed high luminescence efficiency in PS was

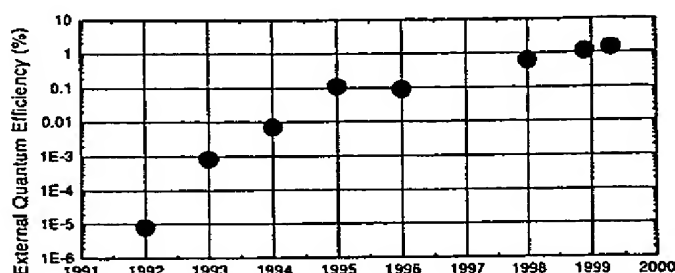


Figure 15. External quantum efficiency of PS LEDs over the year. The record in efficiency to date is that of [51].

- (i) quantum confinement which leads to an enlargement of the bandgap and to an increased recombination probability,
- (ii) the spatial confinement of the free carriers which prevents them reaching non radiative recombination centres and
- (iii) the reduction of the refractive index of the material which increases the extraction efficiency via refractive index matching.

This result has motivated many research efforts in order to exploit these properties in LEDs [50]. The evolution of PS LED performances over the year is reported in figure 15 [51].

The PS approach has however a draw-back in the high reactivity of the spongelike texture which causes the rapid ageing of the LED and an uncontrollable variation of the LED performance with time [50]. No optical gain was reported in bulk PS. From PS, silicon nanocrystals (Si-nc) can be obtained by scrapping or ultrasonically dispersing PS [52]. Then the surface chemistry can be adjusted and, in particular, oxide passivated. Evidence of amplification in these materials has been presented [53].

An alternative way is to produce silicon nanocrystals (Si-nc) in a silica matrix to exploit the quality and stability of the SiO₂/Si interface and the improved emission properties of low dimensional silicon. Many different approaches have been proposed to form the silicon nanocrystals [13, 53]. The most widely used are based on the deposition of sub-stoichiometric silica films, with a large excess of silicon, followed by a high temperature annealing [54]. The annealing causes a phase separation between the two constituent phases, i.e. silicon and SiO₂ with the formation of small silicon nanocrystals. The size and density of the Si-nc can be controlled by the deposition and the annealing parameters. Recently, the anneal of amorphous SiO/SiO₂ superlattices has been proposed to control the size distribution. Almost monodispersed size distribution has been demonstrated [55].

The luminescence properties of Si-nc are very similar to those of PS: a wide emission band is observed at room temperature whose spectral position depends on the Si-nc sizes. In these systems optical gain has been observed [37, 53, 56–62]. Optical gain in Si-nc has been revealed as a superlinear increase of the luminescence intensity as a function of the pumping rate [53, 59], as the measurements of amplified spontaneous emission (ASE) in a waveguide geometry [37, 56, 57, 60–62] (see figures 16–18), as probe amplification in transmission experiments under high pumping excitation [37] or as collimated and speckled patterned emissions which show the coherent properties of the emitted light [59]. Some concerns have been raised about the methods used to measure the gain [63]. Almost all the authors agree on the fact that the gain is due to localized state recombinations either in the form of silicon dimers or in the form of Si=O bonds formed at the interface between the Si-nc and the oxide

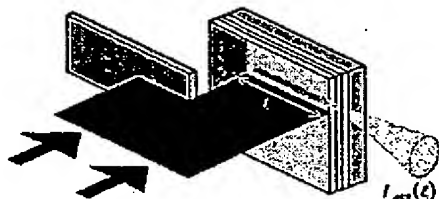


Figure 16. Sketch of the variable stripe length method to measure optical gain. The amplified spontaneous luminescence intensity I_{ase} is collected from the edge of the sample as a function of the excitation length l . The laser beam is focused on a thin stripe by a cylindrical lens.

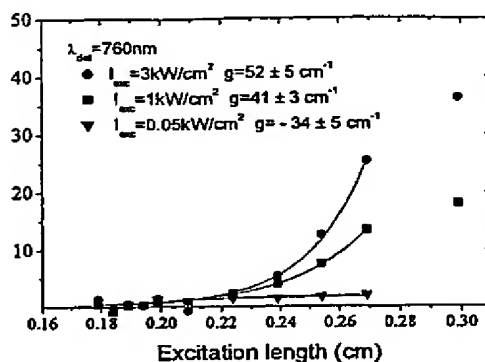


Figure 17. Room temperature VSL curves on a Si-nc sample as a function of the pumping intensities. The detection wavelength was 760 nm. By increasing the pumping intensity from 0.05 kW cm^{-2} to 1 kW cm^{-2} the optical losses turn into optical gain. The values of optical gain become saturated at an intensity of 3 kW cm^{-2} (from [62]).

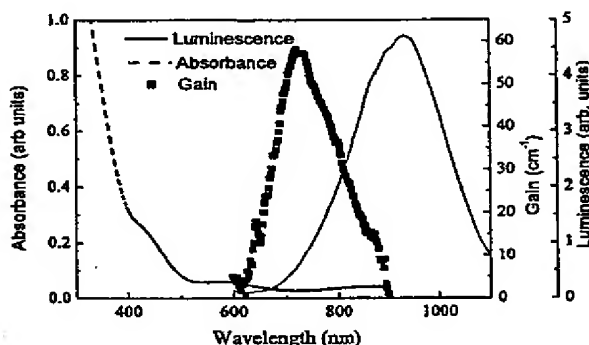


Figure 18. Absorbance (dashed curve), modal gain spectrum (solid squares) and luminescence spectrum (solid curve) for a Si-nc sample (from [62]).

or within the oxide matrix. The suggested scheme to explain population inversion, and hence gain, is a four-level model where a large lattice relaxation of the photoexcited localized centre gives rise to the four levels (figure 19) [56, 62].

Very interesting information can be achieved by time resolved experiments of the ASE from Si-nc in a waveguide geometry [56, 57, 62]. Figure 20 reports the decay lineshape of the ASE both as a function of the pumping fluences (figure 20(a)) and as a function of the excited length (figure 20(b)). In addition to the usual slow recombination of Si-nc (microseconds range), a fast contribution (nanosecond timescale) is observed which grows up either by increasing the fluence or by increasing the excitation length. This last observation rules out Auger recombination as the cause of the fast component because of its strongly non-linear dependence on the photo-excited carrier concentration, which in figure 20(b) is constant for all the various lengths. The origin of the fast component in these Si-nc is stimulated emission. This is also supported by other experimental data.

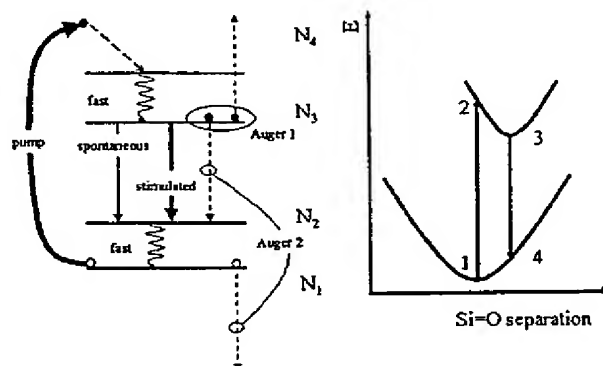


Figure 19. Left, effective four-level system based on the results of figure 18, which has been introduced to model qualitatively the recombination dynamics under gain conditions. From level 3 the excited carriers can recombine by spontaneous, stimulated or Auger recombinations. Right, schematic diagram of the energy configuration diagram of the silicon nanocrystals in an oxygen rich matrix. Localized radiative states are formed inside the nanocrystal bandgap by the interface oxygen atoms. The excited nanocrystal state can occur at a different lattice coordinate with respect to the ground state (from [62]).

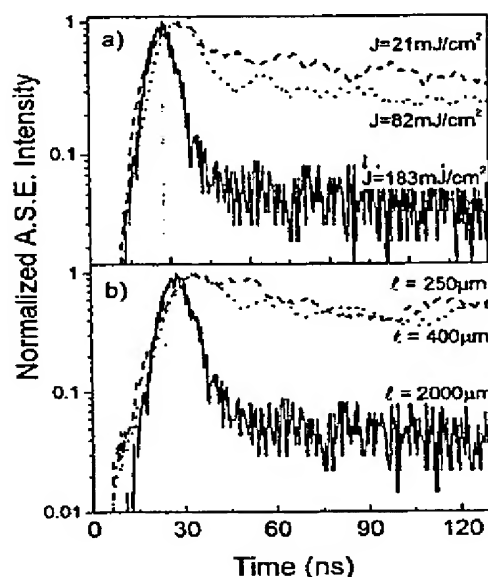


Figure 20. (a) Normalized ASE measured under VSL geometry with a pumping length $l = 2$ mm at the different pumping fluences reported in the figure. The measured sample is a Si-nc waveguide. Excitation wavelength was 355 nm. (b) Here the effect of the pumping length l on the fast ASE dynamics is shown. The pumping fluence is fixed at 183 mJ cm^{-2} and only the pumping length is varied according to the values reported in the figure (from [62]).

Figure 21(a) reports the exponential increase of the fast component intensity as a function of the photoexcited volume (which yields a net modal gain of 12 cm^{-1} under these pumping conditions). Figure 21(b) shows a clear fluence threshold over which the ASE increases

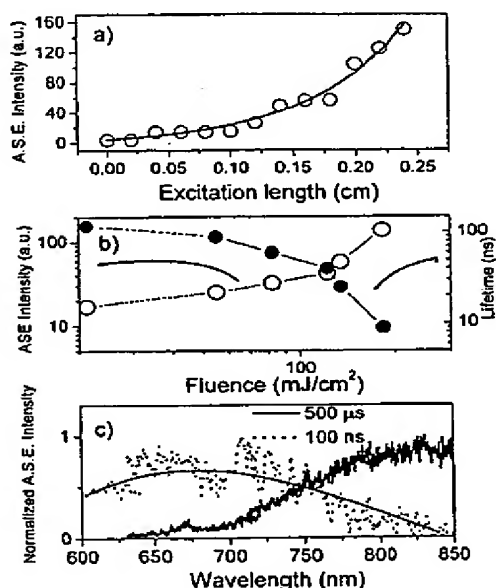


Figure 21. (a) Points: ASE peak intensity at 760 nm versus the excitation length at a pump fluence of 200 mJ cm^{-2} . Full curve: fit of the experimental data with the one-dimensional amplifier model which yields a net modal gain value of $12 \pm 3 \text{ cm}^{-1}$. (b) Open circles: ASE peak intensity of the fast component versus the pumping fluence. Black discs: $1/e$ lifetime of the ASE decay as a function of the pumping fluence. Excitation length was approximately $l = 2 \text{ mm}$. (c) ASE spectra measured for a fixed excitation length $l = 2 \text{ mm}$ and pumping fluence of 200 mJ cm^{-2} for two different integration time windows: dotted curve, 100 ns after the excitation; full curve, 500 μs after the excitation. All the data in this figure have been taken with an excitation wavelength of 355 nm (from [62]).

superlinearly with the fluences, and the decay lifetime of the emission decreases to a few nanoseconds. Figure 21(c) shows that the spectral shape of the fast component is different from the one of the slow component reflecting the typical blue shift of the gain band with respect to the luminescence (figure 18) which supports the four-level model of figure 19. The four-level model is also able to reproduce the decay of the luminescence at high fluences for Si-nc as demonstrated in figure 22. In the simulation of figure 22 both stimulated emission and Auger recombination are taken into account. At the peak fluence the lifetimes associated with these two processes are only slightly different. It is also this delicate interplay between Auger recombination and stimulated emission that governs the optical gain in Si-nc. As discussed in [62], the Si-nc density should be large enough to yield a significant optical gain. This means that optical gain cannot be achieved in all Si-nc samples. It is interesting to note that the data of figure 22 cannot be fitted with only Auger recombinations, even with peak Auger lifetimes as short as 90 ps. The contribution from stimulated emission is needed to accurately reproduce the luminescence decay.

The Si-nc system is very promising to achieve a laser. Indeed, other key ingredients for a laser have been demonstrated. Vertical optical micro-cavities based on a Fabry-Perot structure with mirrors constituted by distributed Bragg reflectors (DBRs) and where the central layer is formed by Si-nc dispersed in SiO_2 have been already fabricated [64]. The presence of the thick SiO_2 layer needed to form the DBR can be a problem for electrical injection when

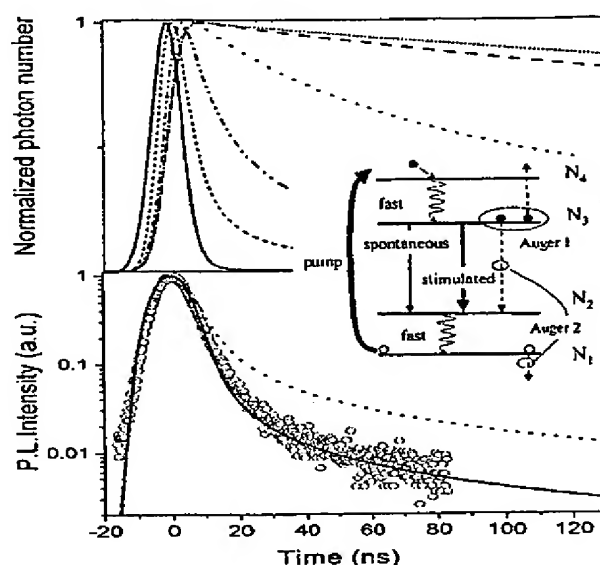


Figure 22. Top panel, simulations of the normalized PL intensity as a function of the incident photon flux ϕ_p . The peak of the incident photon flux ϕ_p was varied between 10^{16} and 10^{24} photons $s^{-1} cm^{-2}$. The main parameters used in the simulation were the pump absorption cross-section $\sigma_p = 10^{-14} cm^2$, the emission cross-section $\sigma = 10^{-17} cm^2$, the active centre concentration $N = 8 \times 10^{18} cm^{-3}$, the spontaneous emission factor $\beta = 4.5 \times 10^{-4}$ and the optical losses $\alpha = 25 cm^{-1}$. No Auger recombination has been considered here. Bottom panel, PL decay (O) of Si-nc produced by PECVD deposition of 46 at.% Si annealed at 1250 °C for 1 h. The solid line is a simulation obtained with the same parameters as in the top panel plus an effective Auger coefficient $C_A = 10 \times 10^{-10} cm^{-3} s^{-1}$ (peak Auger lifetime of 3 ns) and a pump photon flux of 5×10^{22} photons $s^{-1} cm^{-2}$. The dashed line is a simulation where no stimulated emission is present, only Auger recombination. In this case an Auger coefficient of $C_A = 2 \times 10^{-8} cm^{-3} s^{-1}$ (peak Auger lifetime of 90 ps) is needed (courtesy of L Dal Negro).

current has to flow through the DBR. Lateral injection schemes can avoid these problems. On the other hand, the electrical injection into the Si-nc is a delicate task by itself¹. Bipolar injection is extremely difficult to achieve. Despite some claims, most of the reported Si-nc LEDs are impact ionization devices: electron-hole pairs are generated by impact ionization by the energetic free carriers injected through the electrode. By exploiting impact ionization Si-nc LEDs have been demonstrated with EL spectra overlapping luminescence spectra, onset voltage as low as 5 V and efficiencies in excess of 0.1% [66]. Some unconfirmed claims of near-laser action of Si-nc LEDs have appeared in the literature [67, 68].

The problem of gain in Si-nc still has some unanswered issues:

- (i) what is the role played by the Si-nc and by the embedding medium?
- (ii) what are the key parameters which determine the presence of gain in the Si-nc?
- (iii) is the nanocrystal interaction influencing the gain?
- (iv) are low-losses active waveguides possible to achieve?
- (v) what is the precise nature of the four levels in the model, in particular the location and role of Si-O bonds?

¹ An introduction and up-to-date review can be found in [65].

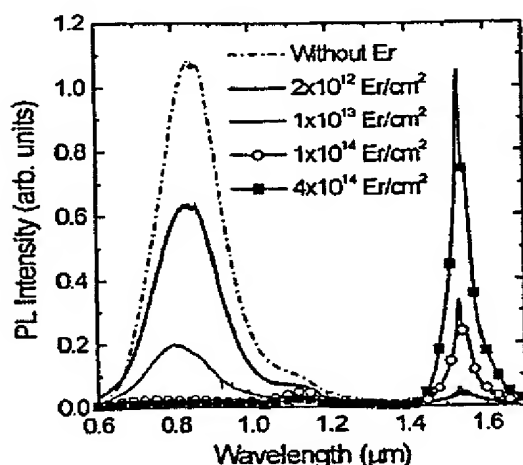


Figure 23. Room temperature PL spectra of Er implanted Si nanocrystals at different Er doses. The pump power of the laser beam was 50 mW (after [74]).

3.3. Er coupled silicon nanocrystals

The recent increase in the transmission capacity of optical fibre based communication is also related to the availability of all-optical amplifiers based on Er doped fibres [69]. In this amplifier, a silica optical fibre is doped with Er^{3+} ions, whose internal atomic-like transition at $1.54 \mu\text{m}$ is exploited to achieve light amplification. In the past, several attempts have been made to reproduce a similar materials system in silicon. Several breakthroughs have been recently achieved in the field of Er doping of crystalline Si that allowed fabrication of LEDs operating at room temperature [71–73].

What is more interesting for light amplification studies is the experimental finding of a strong enhancement of the Er luminescence when Er is implanted or deposited in a SiO_2 matrix where Si-nc have been formed, i.e. Si-nc act as sensitizers for erbium ions [74, 75]. Non-radiative de-excitation processes are reduced by widening the Si bandgap and thus avoiding one of the most detrimental sources of Er luminescence quenching. Indeed, the thermally activated back-transfer of excitation from Er^{3+} to Si-nc becomes less efficient than in bulk Si since the energy mismatch for the process becomes larger. Widening of the bandgap also produces a reduction in the free-carrier concentration, thus limiting the Auger processes. As demonstrated in figure 23, a strong luminescence comes from Er ions that are pumped through an electron-hole mediated process in which photo-excited excitons from Si nanocrystals transfer their energy to Er ions [74]. The number of Si-nc coupled to a single Er ion is still a debated issue (between one and ten) [74, 76]. As concerns where Er is placed, from high resolution luminescence it is clear that most of the Er is in the SiO_2 matrix, which is an ideal situation if one looks at reproducing the environment which is found in an Er doped fibre amplifier. Hence, Er coupled Si-nc benefits from the advantages of both silicon (efficient excitation) and SiO_2 (weak non-radiative processes, i.e. negligible temperature quenching of the luminescence), while it avoids their disadvantages (low excitation efficiency in SiO_2 and strong non-radiative processes in bulk Si). Indeed, MOS light emitting devices operating at room T have been made with this system, where a quantum efficiency larger than 1% [38] is demonstrated. Even higher efficiencies (10%) are reported for Er in silicon rich oxide films; however, in this system reliability is still an issue [39].

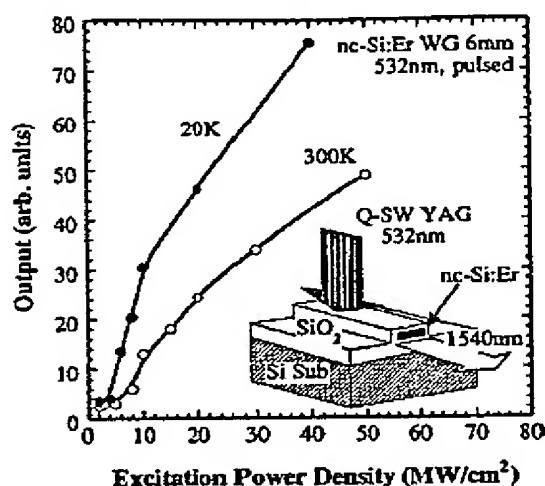


Figure 24. Pumping power density dependence of the 1540 nm emission of a 6 mm Er doped nc-Si waveguide (after [77]).

The layer co-doped with Si-nc and Er^{3+} ions has a refractive index which is larger than that of SiO_2 , i.e. waveguides can be formed with a core containing Er^{3+} coupled to Si-nc. Experiments have shown luminescence increases (figure 24) [77] or even evidence of signal enhancement (figure 25) [78] are present in these waveguides. Even though no net optical gain was measured, an enhancement in the probe transmission at $1.535 \mu\text{m}$ was observed as the pump power was increased. By rather crude approximations, it is possible to write that the probe transmission when the pump is on, $I(P)$, is related to the probe transmission when the pump is off, $I(0)$, by $\text{SE} \equiv I(P)/I(0) = \exp(2(\sigma N_2 \Gamma)L)$, where SE is the signal probe enhancement, σ is the Er^{3+} emission cross-section at $1.535 \mu\text{m}$, N_2 the density of excited Er ions, Γ the optical mode confinement factor and L the waveguide length. A fit to the experimental data yields an increased Er^{3+} emission cross-section with respect to Er ions in silica or in silicon (table 1) [40]. This is a quite unexpected result, which has however been confirmed by other research groups. The reason is still unclear; one can speculate about the role of the dielectric environment which is modified by the presence of the Si-nc [76]. What makes this finding interesting is the possibility of significantly reducing the cavity length in an amplifier or laser below the one usually employed in the silica doped fibre systems. Sizeable gain can be further obtained by low Er doping concentrations. To summarize the very interesting properties of the Er^{3+} coupled Si-nc system, table 1 compares the main cross-sections of Er^{3+} in silica and silicon and coupled with Si-nc.

The system Er^{3+} coupled to Si-nc is very promising for laser applications because the active material (Er^{3+} in SiO_2) has already shown lasing properties. In addition, the technology to produce the material is very compatible with CMOS processing. Microcavities with excellent luminescence properties have been also demonstrated [64], which allows design of both edge emitting and vertical emitting laser structures. The issue related to electrical pumping of the active material, which was believed to be a major short-cut of this approach, can be solved as extremely high efficiency LEDs have been demonstrated [38, 39]. A still open issue is to engineer the waveguide losses in order to be able to measure net optical gain and not only signal enhancement in a pump and probe experiment. This seems only a problem of time and

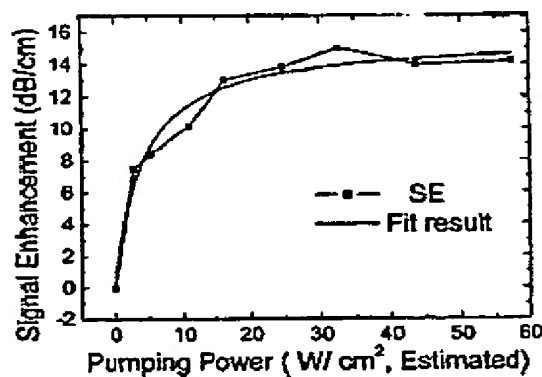


Figure 25. Pump power dependence of the signal enhancement SE and of a theoretical fit. An SE of up to 14 dB cm^{-1} , implying a possible net gain of up to 7 dB cm^{-1} , is found. From the fit, an emission cross-section of $2 \times 10^{-19} \text{ cm}^2$ and an effective excitation cross-section of $> 10^{-17} \text{ cm}^2$ at 477 nm was deduced (from [78]).

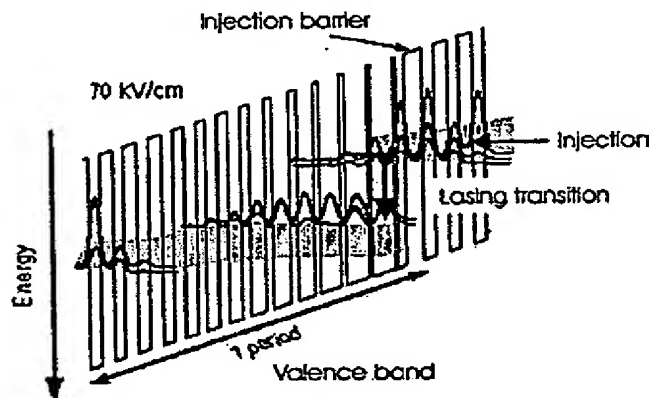


Figure 26. Schematic valence band diagram of one stage of the structure, under an applied electric field of 70 kV cm^{-1} . Only the HH band and the modulus squared of the relevant HH wavefunctions are shown for clarity. Note that the axis of the energy is turned upside down. Each period, starting from the injection barrier, consists of the following sequence of Si barrier (roman) and $\text{Si}_{0.2}\text{Ge}_{0.8}$ (bold) in Å: 25/11/4/26/5/26/6/24/7/21/8/19/9/18/10/17/11/15/12/15/13/14/15/14/16/13/17/13. The underlined numbers correspond to doped layers with a boron concentration of $5 \times 10^{17} \text{ cm}^{-3}$ (from [88]).

research efforts. Then further work should be spent to optimize the gain with respect to the waveguide parameters and develop a suitable optical cavity which can be electrically injected.

3.4. Si/Ge quantum cascade structures

One route to avoid the fundamental limitation to lasing in silicon, i.e. its indirect bandgap, is to avoid using interband transitions. Indeed, if one exploits only intraband transition, e.g. intra-valence-band transition, no fundamental problems exist to impede lasing in silicon [82]. This

Table 1. Summary of the various cross-sections related to Er^{3+} in various materials.

	Er in SiO_2 (cm^2)	Er in Si (cm^2)	Er in Si-nc (cm^2)	Reference for Er in Si-nc
Effective excitation cross-section of luminescence at a pumping energy of 488 nm	$(1-8) \times 10^{-21}$	3×10^{-15}	$(1.1-0.7) \times 10^{-16}$	[79, 80]
Effective excitation cross-section of EL		4×10^{-14}	1×10^{-14} by impact ionization	[38]
Emission cross-section at 1.535 μm	6×10^{-21}		2×10^{-19}	[40]
Absorption cross-section at 1.535 μm	4×10^{-21}	2×10^{-20}	8×10^{-20}	[81]

is indeed the approach of the quantum cascade (QC) Si/Ge system. With SiGe QC lasers, one is trying to use the concept that has already been successful in III-V semiconductors, which is advancing as a viable option for mid-IR emission, covering today a large wavelength range, 3–24 μm [83].

The idea of the device is shown in figure 26. The QC scheme can be implemented in the conduction or valence band. However, to achieve a conduction band discontinuity the growth of a Si/Ge superlattice on a relaxed SiGe buffer is necessary. For pseudomorphic growth on a Si substrate most of the band offset occurs in the valence band. Hence, the cascading scheme is usually designed in the valence band [84]. This differs from QC lasers based on III-V semiconductors that employ electron cascade structures. In figure 26, the valence band diagram of a cascading stage of a hole-injected p-i-p valence band device is shown. Injected holes make a vertical transition between subbands, and then they cascade down the electrically biased staircase. In order to assist population inversion, the lower laser level is rapidly depopulated by relaxation within the miniband. Practically, one has two identical active regions connected by an injector. EL from a SiGe QC structure grown on Si has recently been demonstrated [84, 88].

Starting from the possibility of monolithic integration with silicon microelectronics, the Si/SiGe system is more interesting than III-V heterostructures for QC laser applications. The non-polar electron-phonon interaction is the dominant loss process in III-V QC lasers. In silicon, due to the covalent bonding, the non-polar phonon scattering is absent. The optical phonon energy in Si is much higher than in GaAs (64 meV compared with 36 meV), providing a larger frequency window within which (non-polar) optical phonon scattering is suppressed. In Si the thermal conductivity is much larger than that of GaAs, giving better prospects of CW operation at non-cryogenic temperatures. On the other hand, some constraints are present [85]: the necessity to work in the valence band and thus the higher effective masses of the charge carriers, limited band offset of approximately 80 meV per 10% Ge concentration and splitting into heavy hole (HH) and light hole (LH) bands. Moreover, the high amount of strain, due to the lattice mismatch between Si and Ge, sets an upper limit to the number of wells per cascade and the number of cascades, as well as the thickness and Ge content of each individual well. Due to the mentioned constraints, the developed Si/SiGe cascade structure is a drastically simplified version of the typical III-V QC structures. As shown in figure 26, in a practical QC structure each cascade consists of only a few wells [85].

Figures 27 and 28 show the typical EL spectra recorded in QC structures grown on Si substrates [87, 88]. The levels involved are valence levels; the radiative transition is between HH states. The quantum efficiency estimate is about 10^{-5} for EL [85–88]. Temperature-

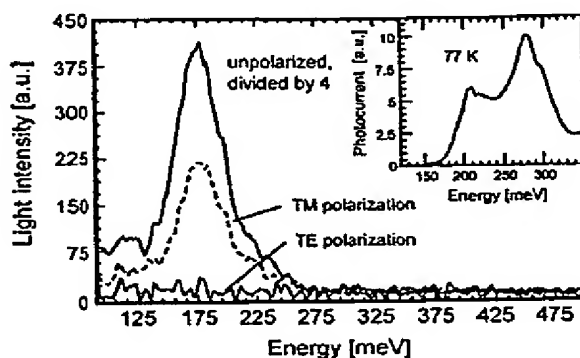


Figure 27. EL spectra of the sample with 15 repetitions, taken at 80 K with and without a polarizer placed in the light path. The parameters are 4.7 V, 550 mA, 94 kHz and a duty cycle of 10%. The polarized EL is measured at 5.2 V, 650 mA and a 20% duty cycle. The inset shows the results of a photocurrent measurement at 77 K (from [88]).

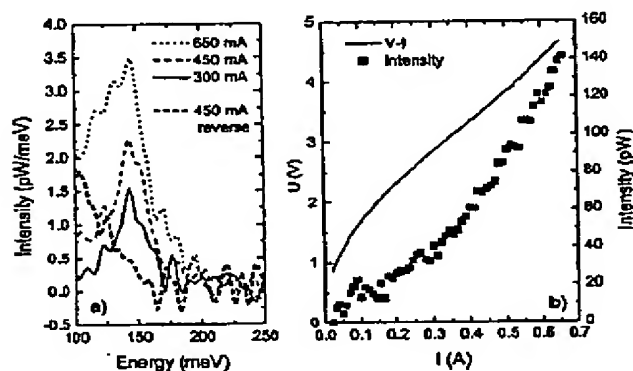


Figure 28. (a) Current-dependent EL spectra in forward bias and spectrum at reverse bias at 80 K. (b) I - V curve and integrated EL intensity (from [87]).

dependent measurements show nearly identical spectra between 20 and 90 K and a broadening and vanishing of the peak at about 160 K. It is possible to improve these results controlling the large accumulation of strain imposed by the use of a Si substrate. This has been done by using a $\text{Si}_{0.5}\text{Ge}_{0.5}$ substrate and growing on it strain compensated $\text{Si}_{0.2}\text{Ge}_{0.8}/\text{Si}$ quantum wells. Intersubband transitions have been observed by absorption measurements at 235, 262 and 325 meV changing the well width from 3.5 to 2.5 nm; peaks are observed up to room temperature [88]. For similar structures EL has been detected at 80 K [88].

The QC concept works for III-V semiconductors. The SiGe system has some advantages and a fundamental limit posed on the number of periods of successive QW cascades which is given by the critical thickness for the formation of misfit dislocation. Hence, even though these devices show interesting EL properties for the prospect of the development of a Si based laser, highly evolved cascade structures have to be realized. As the gain per single element is low due to the nature of the intraband transition, a large number of cascading structures will be needed to accumulate a macroscopic gain. In fact, no stimulated emission in SiGe

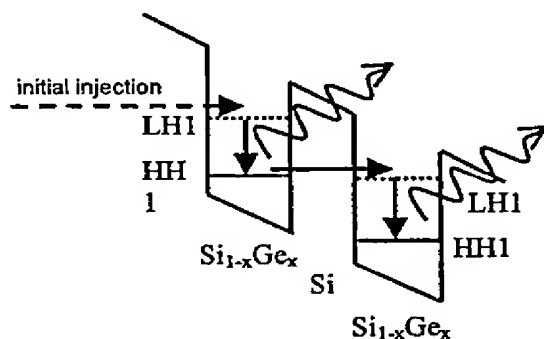


Figure 29. Schematic valence band profile of a Si/SiGe quantum staircase laser operating via radiative LH1–HH1 transitions (from [92] with kind permission of Kluwer Academic Publishers).

QC structures has been reported to date. In addition, all these have to be integrated within a waveguide cavity. In addition, the emission wavelength is different from those commonly used for optical interconnects. A waveguide for these wavelengths can be realized by using SOI substrates or thick, relaxed SiGe graded buffer. The other photonic components have still to be developed to achieve a photonic integrated system. Although some authors propose to use a QC laser for free-air optical interconnects, such a Si/Ge QC laser will be of little use for silicon photonics if all other compatible elements will not be developed.

3.5. THz emission

A gap in the frequency spectrum of electromagnetic waves opens across the THz region, where no semiconductor sources are available. At low frequencies, sources are made by electronic oscillators (high speed transistors) while at high frequencies the sources are made by injection lasers. Recently, a THz laser has been demonstrated by using III–V semiconductors which shows the way to cover this THz gap [89]. With the same aim, and using the many advantages of the SiGe system over the III–V systems for these frequencies, a research effort is made to implement the QC concept and make a laser in these frequency regions [90–92]. A typical structure is shown in figure 29 which by using p-type heterostructures is designed to emit radiation from LH–HH transitions. In this way both edge emission and surface-normal THz emission might be obtained. Growth of p-Si/SiGe QC structures comprising up to 100 periods has been demonstrated using low pressure CVD via a strain balanced approach on virtual substrates. Intersubband THz EL from a range of Si/SiGe QC structures has been observed in both edge and surface-emission geometries. An example is shown in figure 30. The LH–HH intersubband lifetime was measured to be ~ 20 ps, which is over an order of magnitude longer than high temperature values in III–V heterostructures, implying that a Si/SiGe THz QC laser may be capable of much higher operating temperatures than corresponding III–V devices. Emission power levels comparable to the one reported on III–V devices before laser processing have been measured which indicate that there are good prospects for realization of a THz Si/SiGe QCL via further optimization of the active region and appropriate cavity design [92].

Another approach to THz laser emission in silicon has been developed [93–96]. The idea is to make a THz laser using intra-shallow donor optical transitions in silicon. A band diagram showing the lasing transition is reported in figure 31. Very narrow spectral emission and the

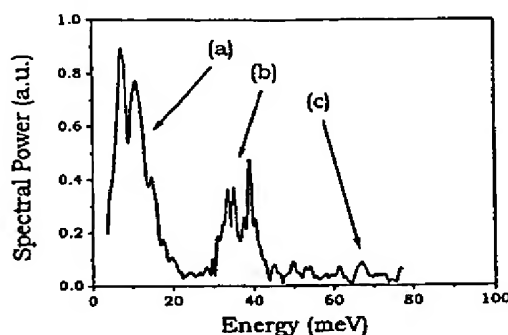


Figure 30. FTIR edge-emission spectrum for a QC structure, at a temperature of 4.2 K. The pulsed bias voltage was 7 V with a 10% duty cycle. The features marked (a), (b) and (c) correspond to the theoretically calculated emission peaks for the LH1-HH1, HH2-HH1 and LH2-HH1 intersubband transitions, respectively (from [92] with kind permission of Kluwer Academic Publishers).

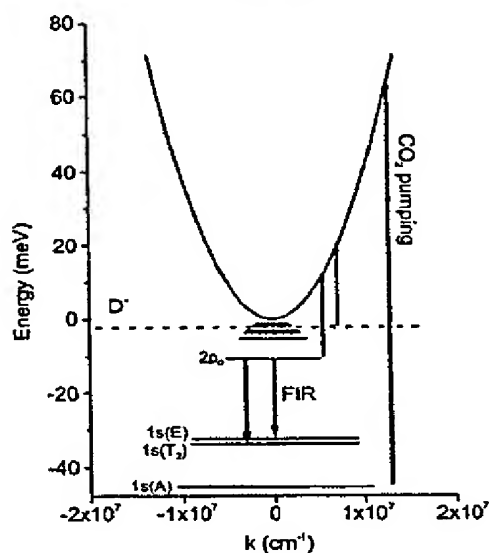


Figure 31. Optical transitions in Si:P. The dashed line represents the energy level of the D^0 centre state (from [96]).

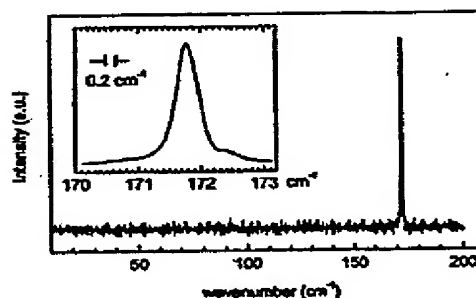


Figure 32. Stimulated emission spectrum from Si:Sb. The emission curve is identified with the $2p_0 \rightarrow 1s$ intracentre Sb transition (from [96]).

light intensity threshold versus pumping power are reported in figures 32 and 33. All these data should indicate that lasing has been achieved in this system. However, some points need to be clarified, such as the optical mode pattern in the simple cavity structure used, the evolution from spontaneous to stimulated emission and the coherent property of the light. Other concerns are related to the dilute doping of the system in order to avoid impurity-impurity interaction which will prevent population inversion and the schemes for electrical injection. It is clear that the use of THz laser sources for silicon microphotonics requires a complete reshaping of the scheme developed up to now.

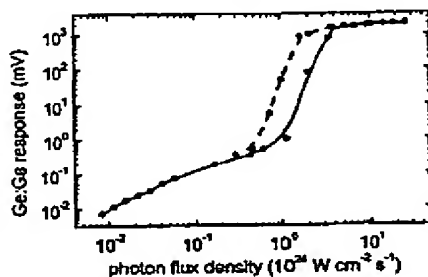


Figure 33. Dependence of the emission on the pump power for 9.6 μm excitation (dashed curve) or the 10.6 μm excitation (solid curve) (from [96]).

4. Conclusion

Throughout this review, I have tried to describe the status of silicon microphotonics and the recent advances that cause people to be optimistic to the realization of an active silicon light source. Indeed, many claims to have a silicon based laser within a short period have appeared in the literature by many of the researchers involved in this field [97]. If this objective is realized all the major building blocks for monolithic silicon microphotonics will be available.

The final vision is to have Si microphotonics participating in every global application of the photonics industry: communications, computing, information displays, optical-and-infrared imaging, medicine, optical printing, optical command-and-control, optical sensing of physical chemical and biological inputs, optical signal processing, optical storage and optical control of microwave devices or systems [98]. We indeed propose silicon as the unifying material where the next generation of photonics devices will be realized.

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Quasicontinuous gain in sol-gel derived CdS quantum dots

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We report evidence for quasicontinuous optical gain in CdS quantum dots fabricated by the sol-gel process and embedded in glass. The gain spectra are obtained using the pump and probe technique and nanosecond (quasiresonant) excitation at 11 K. The dots are in the intermediate quantum confinement regime and the concentration of CdS is relatively high. The gain, which is spectrally broad, develops on the low energy side of the absorption band edge. The reason why the gain region is broad is not only the size distribution of the dots, but also the nature of the gain, which originates from the recombination of several excited levels between two and one electron-hole pairs states (i.e., biexciton to exciton). The maximum measured gain reaches 200 cm^{-1} at 11 K and 17 cm^{-1} at 170 K. © 1995 American Institute of Physics.

Three-dimensionally quantum confined semiconductors are attractive because of the predicted enhanced optical properties with increasing confinement. Semiconductor quantum dots (QDs) have been investigated extensively,¹ but they have not shown superior emission properties. One known problem is the reduced radiative efficiency with decreasing dot as a result of surface recombination. For QDs in glass especially, limitations are the large size distribution of the dots, the presence of trap states (such as vacancies, substitutional defects, and dangling bonds), the observed photodarkening and low density of microcrystallites (active material) in the glass matrix.

Semiconductor crystallites in glass have a number of advantages including the relatively easy growth of QDs of different II-VI semiconductors by heat treatment and the possibility of fabricating low loss waveguides with ion exchange in the high quality glass matrix. The quality of our QDs has been improved with the use of the sol-gel synthesis:² high dot density, dots sizes much smaller than the bulk exciton Bohr radius a_{Bohr} , with more uniform size distribution, and an almost complete elimination of the photodarkening effect.³

Femtosecond dynamics of optical gain in strongly confined CdSe QDs have been recently studied using femtosecond (fs) optical excitation.⁴ Gain in the weak confinement regime for CuCl crystallites in glass has been also studied at 77 K.⁵

In this letter, we report our measurement of the gain spectra of sol-gel derived CdS QDs glasses in the intermediate confinement regime and compare the results with the calculated gain for a quasi-zero-dimensional electron-hole system, accounting for one and two electron-hole pair states. Using the variable-stripe length method, we analyze the temperature dependence of the gain between 11 and 170 K.

The CdS nanocrystallites we investigated were fabricated by the sol-gel process.² 5.6 wt % CdS nanocrystallites

were embedded in a fully dense $5\text{Na}_2\text{O}-15\text{B}_2\text{O}_3-80\text{SiO}_2$ (mol %) glass after heating at 590°C . The transmission electron micrograph (TEM) in Fig. 1(a) shows that the average dot radius R was 2.8 nm with a distribution of σ (standard deviation) = 1.0. The average radius R of the dots was on the order of the bulk exciton Bohr radius a_{Bohr} . Figure 1(b) shows a typical image of these nanocrystallites with high resolution transmission electron microscope (HRTEM: Topcon 002B, 200 kV). These crystallites of $R \approx 2.5\text{ nm}$ (graphitized carbon internal standard) are both oriented perpendicularly to the (101) reflection plane. The electron and x-ray diffraction pattern of our sample show that the crystallites are of hexagonal symmetry.

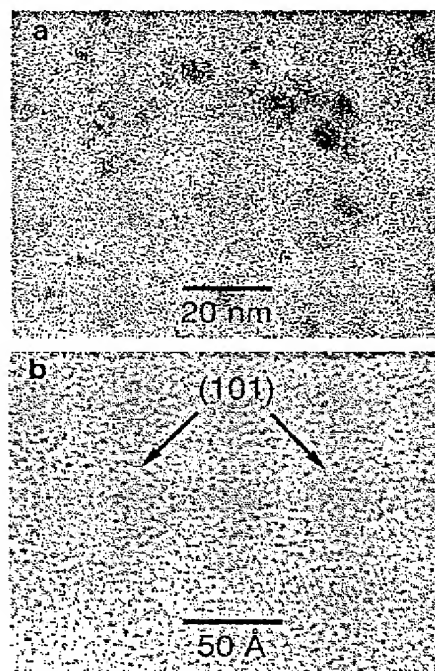


FIG. 1. (a) TEM micrograph and (b) HRTEM micrograph of the 5.6 wt % CdS-doped sodium borosilicate glass.

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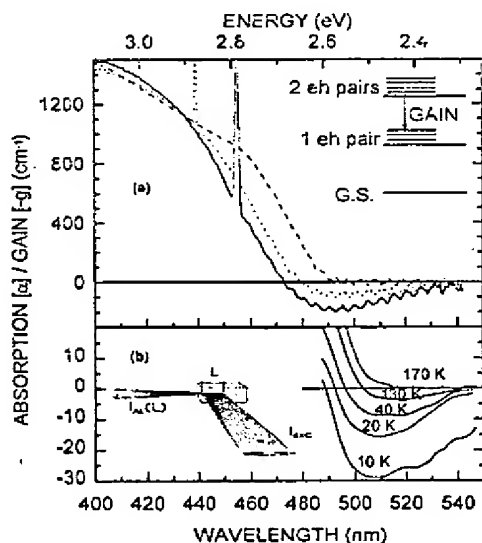


FIG. 2. (a) Absorption (gain) spectra of CdS QDs measured at 11 K for different pump excitation intensities I_{exc} . The pump wavelength is 455 nm for the solid curve ($I_{exc} = I_0$) and 440 nm for the dotted curve ($I_{exc} = I_0/7$), with $I_0 = 570$ kW/cm². The dashed curve is the linear absorption spectrum. The inset shows a scheme of the energy levels considered for the model calculations: the ground state and the one and two electron-hole pairs levels with their excited states, and one of the recombinations which provides the optical gain. (b) Temperature dependence of the gain between 10 and 170 K. The spectra were measured with the variable-stripe-length method (see the inset), with an excitation intensity $I_{exc} = 3I_0$ at 355 nm.

The glass sample was polished to a thickness of 17 ± 2 μ m and held in a cryostat at $T = 11$ K for optical characterization. Differential transmission (DTS) and absorption spectra have been measured in the common pump and probe technique, with 3 ns duration pump pulses at 10 Hz repetition rate. The probe beam, originating from the luminescence of a dye cell, had an 8 ns pulse duration and was cross polarized with respect to the pump to reduce the scattered light. The time delay between the pump and the probe pulses was less than 1 ns.

Figure 2(a) shows the absorption spectra for different pump intensities. A bleaching of the absorption spectrum is observed as the excitation intensity is increased. The position of the maximum (around 465 nm) does not change when we vary the pump wavelength, even for low excitation intensity. This can be interpreted as a low inhomogeneous broadening,⁶ giving another indication of the relatively narrow size distribution of the QDs. A broad gain region (negative absorption), ranging from ≈ 470 to ≈ 540 nm, is observed for the highest pump intensity. These results are in agreement with the results of Dneprovskii *et al.*,⁷ who have observed gain in CdSe QDs in the intermediate confinement regime. The maximum gain reaches about 200 cm⁻¹ at 493 nm for an excitation intensity of $I_0 = 570$ kW/cm². The rapid oscillations in the gain region are Fabry-Pérot interferences, whose period corresponds to the sample thickness. The pump wavelength is 455 and 440 nm, respectively, for the two input intensities I_{exc} , as indicated by the sharp scattering signals. The comparison with the linear absorption spectrum (dashed curve) shows that the gain develops on the low energy side of the absorption band edge and extends well be-

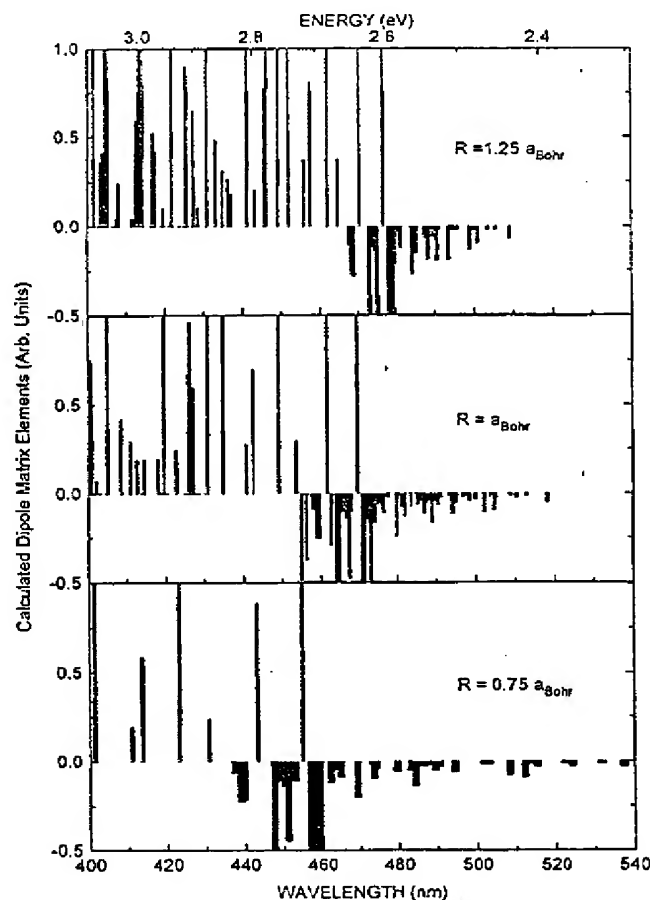


FIG. 3. Calculated absorption (gain) spectra. The bars correspond to the relative strength of the optical transitions between (i) the one electron-hole pairs and the ground state (positive=absorption) and (ii) the two electron-hole pairs and the one pairs states (negative=gain). Calculations are made for $R = 0.75 a_{Bohr}$, $R = a_{Bohr}$, and $R = 1.25 a_{Bohr}$ relative to the bulk band-gap value of CdS (2.56 eV) and without any broadening mechanism. The model takes into account the Coulomb interaction, the valence band mixing, and the surface polarization effects.

low the band edge. This behavior is expected of the gain involving two electron-hole (eh) pairs (biexciton) recombination.⁸

The temperature dependence of the gain between 11 and 170 K is presented in Fig. 2(b). The variable-stripe length method was used to measure the gain spectra.⁹ The third harmonic of the Nd:YAG laser was focused onto the sample by using a cylindrical lens, to form a narrow rectangular stripe, 0.05–2 mm long and about 20 μ m wide. The emitted light (i.e., amplified luminescence I_{AL}) was collected from the edge of the sample in the direction of the stripe, as shown in the inset of Fig. 2(b). The method allowed us to reproduce the gain spectra of Fig. 2(a) (solid line), when we pumped with $I_{exc} \approx 8I_0$ at 355 nm. As expected, the gain decreases with increasing temperature, rapidly at the beginning (it is already two times smaller at 20 K), and it disappears around 170 K. With increasing the excitation intensity to 10 MW/cm², the maximum gain at 170 K was 17 cm⁻¹.

Figure 3 shows the calculated gain and absorption spectra in the absence of any broadening mechanism. The bars

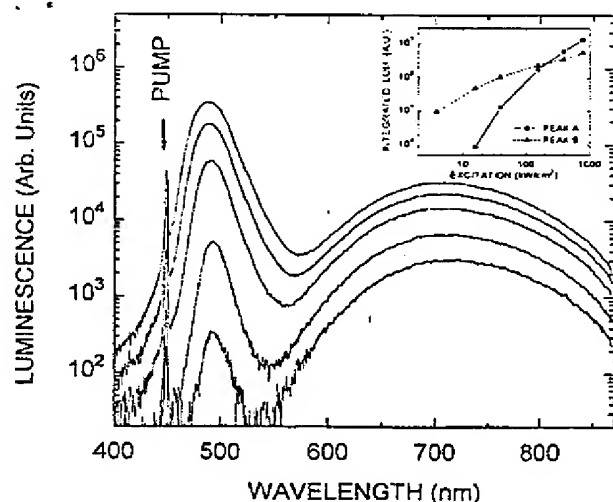


FIG. 4. Photoluminescence (PL) spectra at 11 K for different pump excitation intensities (from 4 to 770 kW/cm²). The intensities are in logarithmic scale. The narrow peak (at 490 nm) is attributed to excitonic recombination and the broad peak (around 700 nm) is probably due to trap states recombination. The pump wavelength is indicated on the figure. The inset shows the log-log plot of the integrated luminescence (areas) of the two PL peaks as a function of the pump excitation intensity. The circles are for the narrow peak surface (A) and the triangles are for the broad peak surface (B).

correspond to the calculated dipole matrix elements in relative units. The transitions between the one *eh* pair states and the ground state are shown as positive bars (above the baseline), and the transitions between two *eh* pairs and the one pair states are shown as negative bars (below the baseline). Calculations are made for three different confinement values: $R = 0.75a_{\text{Bohr}}$, $R = a_{\text{Bohr}}$, and $R = 1.25a_{\text{Bohr}}$, with an infinite quantum confinement potential. A detailed description of the calculation, which includes the numerical diagonalization of the Hamiltonian for the one and two *eh* pairs wave functions and the calculation of the eigenenergies and of the transitions dipole moments, will be published elsewhere.⁸ The valence band mixing is accounted for by using the Luttinger Hamiltonian within the axial approximation [with $m_e = 0.1m_0$; $\gamma_1 = 5$; $\gamma_2 = \gamma_3 = 1.8$, $\epsilon = \epsilon_1/\epsilon_2 = 10$, where $\gamma_1, \gamma_2, \gamma_3$ are the Luttinger parameters and $\epsilon_1(\epsilon_2)$ is the dielectric constant inside (outside) the dot]. The model takes into account the several excited levels of the two and one *eh* pairs states (i.e., biexciton and exciton), which can have a recombination energy smaller than the ground states transition energy [see the inset of Fig. 2(a)]. The population among the different excited energy levels is assumed to be the same for each quantum dot level, as there is not a thermal distribution between the levels which can be populated by only one *eh* pair of the same quantum state. We see that the gain is due to these two *eh* pairs recombination which extend on the low energy side of the first one pair transition. As the quantum confinement is increased (see the lower part of Fig. 3), the gain reaches lower energies and the absorption is shifted to higher energies. Due to the size distribution of the QDs in our sample,

the light emitted by the strongest gain region around the band edge is reabsorbed by the bigger dots (upper part of Fig. 3).

We have taken photoluminescence (PL) spectra of the same sample (see Fig. 4). They show two emission bands as it is usually the case for the QDs in glass:¹⁰ one broadband in the near-infrared, which is related to defects, and a narrower band at shorter wavelength, in the same range where the gain is observed. The excitation intensity dependence of the integrated luminescence of these two bands indicates that the narrow band is growing faster than the broad one with increasing excitation, supporting the fact that the broad emission is related to impurities or defects whereas the narrow emission band has an intrinsic character. The increase in the integrated luminescence of the narrow band is superlinear with the intensity. This can be understood because (i) the luminescence comes from biexcitonic recombination, (ii) there is stimulated emission in that region, and (iii) there are two different recombination channels, one intrinsic channel and one channel related to defects. At higher excitations, the defect-related channel saturates (see the inset in the Fig. 4).

In summary, CdS QDs in the intermediate quantum confinement regime show significant optical gain as high as 200 cm⁻¹ at low temperature. The gain persists up to 170 K, but it decreases with increasing temperature. The sample is fabricated by sol-gel process and has a relatively high concentration of CdS embedded in a borosilicate glass. Our model calculations indicate that the broad gain region we observe is not only due to the size distribution of the dots, but also to the fact that the gain originates from several excited states transitions between two and one electron-hole pairs.

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Room-temperature gain at 1.3 μm in PbS-doped glasses

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We report on room-temperature optical gain at the ground exciton transition of PbS quantum-dot-doped glasses while optical pumping into the next-higher exciton resonance. The material gain in the quantum dots is as large as 80 cm^{-1} . The dot-size selective excitation provides tunability of the optical gain. This is demonstrated by tuning the gain from 1317 to 1352 nm by changing the pump wavelength from 900 to 980 nm. © 1999 American Institute of Physics.
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A major advantage of semiconductor-doped glasses over epitaxially grown structures is that glass is an inexpensive and robust material. Recent improvements in the manufacture of quantum dots (QDs) embedded in glassy matrices have resulted in structures with more uniform-size distribution; fewer vacancies, substitutional defects, and dangling bonds; higher dot concentration; and reduced photodarkening. As in epitaxially grown QD structures, the three-dimensional quantum-confinement effects in the incorporated semiconductor QDs allow for tailoring the linear and nonlinear optical properties of these materials. Thus, semiconductor quantum-dot-doped glasses are very promising candidate materials for photonics applications and may have niche applications relative to the complicated and expensive epitaxially grown structures.

In this letter, we report on room-temperature optical gain in PbS quantum-dot-doped glasses in the communication-wavelength region. When pumping into the first-excited exciton transition, optical gain is observed in the vicinity of the ground exciton resonance. We demonstrate that the spectral position of the peak gain can be changed from 1317 to 1352 nm by tuning the pump wavelength between 900 and 980 nm, a wavelength range which is accessible with InGaAs laser diodes. This tunability relies on the strong carrier confinement and the inhomogeneous broadening in the sample due to dot-size fluctuations, whereas the actual spectral width and position of the gain is given by the pump pulse.

In our experiments, we used PbS quantum-dot-doped glasses which were fabricated by a thermal treatment of an oxide molten glass.¹ In this method, the subsequent thermal treatment of the melted glasses precipitates the microcrystalline phase. PbS quantum-dot-doped glasses exhibit strong three-dimensional quantum-confinement effects at moderate nanocrystal size because of the large bulk exciton Bohr radius of $a_B \approx 18\text{ nm}$. This, combined with the small band-gap energy of 0.41 eV (room temperature) of PbS, allows for tuning the ground exciton absorption from the visible to 3 μm . Figure 1(a) shows the room-temperature absorption spectra of PbS-doped glasses fabricated with different ther-

mal treatment schedules, which result in different average dot sizes. The strong quantum confinement in these structures is clearly observed in the large blueshift of the 1s-absorption resonance with decreasing dot size. The appearance of defined subbanded peaks in all absorption spectra demonstrate the high quality of our samples and the relatively small size distribution of the PbS QDs. The average radii R of the QDs are deduced from fitting the spectral positions of the lowest-energy absorption peaks, which have been determined from the first derivative of the absorption spectra shown in Fig. 1(a), to the calculated 1s-transition energies.² Here, we used the analytical hyperbolic band (HB) model,³ which phenomenologically includes the nonparabolicity of the band structure and provides very good estimates for the 1s-transition energies. Figure 1(b) compares the calculated dot-size-dependent energies of the 1s and 1p transition with the measured absorption maxima. As we can see, when the HB model is used to fit the 1s-transition energy, the 1p-transition energy is slightly underestimated.

To investigate the dynamics of the nonlinear absorption, we performed two-color pump-probe experiments. In these experiments, we used orthogonally polarized 130 fs pulses which are independently tunable in frequency. Pump and probe pulses are obtained from two optical parametric amplifiers, which are synchronously pumped by one regenerative Ti:sapphire amplifier at a repetition rate of 1 kHz. The zero time delay and the time resolution are given by the cross

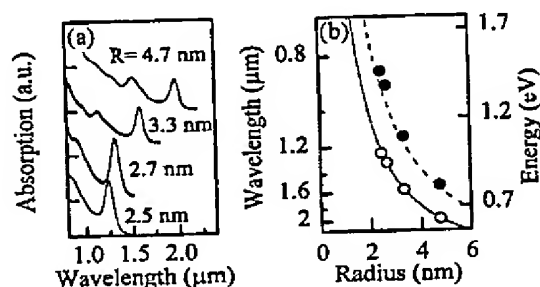


FIG. 1. (a) Room-temperature absorption spectra of PbS quantum-dot-doped glasses with different dot radii R . (b) Calculated 1s- (solid line) and 1p-transition (dashed line) energies; open (filled) circles; position of first (second) absorption peaks in (a).

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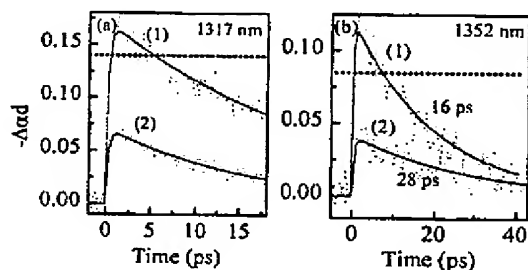


FIG. 2. Dynamics of the nonlinear absorption for the sample with $R = 2.7$ nm, (a) probe at 1317 nm, which is at the $1s$ -absorption peak, with a pump fluence of 4 mJ/cm^2 and pumping at 900 nm (1), which is at the $1p$ -absorption peak, or pumping at 980 nm (2), which is on the low-energy side of the $1p$ -absorption peak [see Fig. 3(a)]; (b) probe at 1352 nm and pump at 980 nm with pump fluences of 4 mJ/cm^2 (1) or 2 mJ/cm^2 (2). No gain is observed for probing at 1352 nm and pumping at 900 nm; the linear absorption is marked by the dotted horizontal lines.

correlation of both pulses, which has been measured to 200 fs. To minimize the fluctuation in the detected signal due to the laser-intensity noise, we utilized a dual-beam setup for the probe beam: one invariant reference path and one signal path that contains the sample. The signals of both (probe) beams are then detected with an autobalanced photoreceiver (Nirvana, New Focus), which cancels out signals that are common to both channels. With this technique, we are able to detect small transmission changes; however, the spectral resolution is limited by the pulse spectral width due to the spectral-integrated detection.

In the following, we concentrate on results obtained at the sample with an average dot radius of 2.7 nm. Figure 2 shows the dynamics of the nonlinear absorption for two different probe wavelengths: (a) resonant with and (b) 35 nm below the maximum of the $1s$ -absorption peak. In Fig. 2(a) the pump wavelength is tuned and set either to the maximum of the $1p$ -absorption resonance [curve (1)] or 80 nm below this maximum [curve (2)]. Fig. 2(b) shows the bleaching dynamics for two different pump intensities, whereas the pump wavelength is fixed at 80 nm below the maximum of the $1p$ -absorption peak.

The dynamics of the bleaching signal can be fitted using a simple asymmetric response function with exponential rise and decay times (solid lines). All transients shown in Fig. 2 exhibit the same ultrafast rise time of about 300 fs, which is only slightly above the time resolution of our experiments. The recovery dynamics of the bleaching signal changes with the pump fluence, i.e., the decay time decreases with increasing pump fluence from 28 to 16 ps. The dotted horizontal lines in Fig. 2 mark the value of the linear absorption $\alpha_0 d$ at the respective probe wavelength.

The sub-ps buildup of the bleaching signal is consistent with previous observations of the ultrafast $1s$ dynamics in glass samples⁴⁻⁶ and indicates that the phonon bottleneck⁷ is not effective in this system. Note the large excess energy of the excited carriers, which varies with the pump-probe detuning between approximately 12 and 16 LO-phonon energies.⁸ Different mechanisms have been proposed to overcome the phonon bottleneck.⁹⁻¹² Klimov and McBranch⁴ showed that in a glass sample, the observed short rise time can be explained in terms of an Auger-like mechanism,¹² which involves confinement-enhanced energy

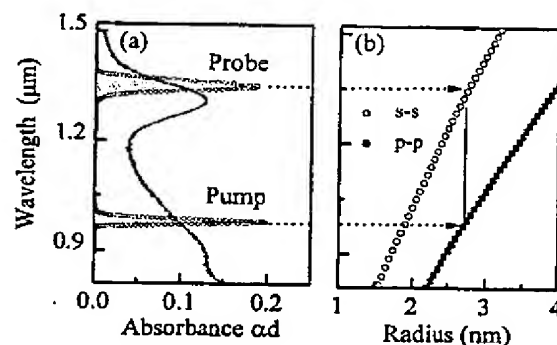


FIG. 3. (a) Absorption spectrum of the sample with $R = 2.7$ nm and pump and probe pulses of Fig. 2(b). (b) Calculated $1s$ - (open circles) and $1p$ -transition (full squares) energies; the arrows mark the positions of pump and probe pulses as shown in (a).

transfer of the electron excess energy to a hole, with subsequent fast relaxation through its dense spectrum of states. The observed ps decay of the bleaching signal agrees well with the excitonic lifetime observed in various quantum-dot glass samples.^{6,13-16} The origin of this fast and pump-fluence-dependent decay of the excitonic population has been discussed either in terms of carrier trapping effects,¹⁴⁻¹⁶ i.e., at surface located defects, or Auger recombination.¹⁷⁻¹⁹

Optical gain is observed when the negative nonlinear absorption exceeds the linear absorption (dotted horizontal lines in Fig. 2). The buildup time of the gain, as seen in each curve (1) in Figs. 2(a) and 2(b), is about 500 fs and the gain lasts for 5–6 ps. The maximum gain value is $gd = -(\alpha_0 + \Delta\alpha)d \approx 0.022$, which corresponds to 15% (a) and 28% (b) of the linear absorption. From the measured optical gain and the filling factor (0.15%) of the QDs in the glassy material, we estimate a material gain in the QDs of 80 cm^{-1} . This room-temperature material gain is large compared to earlier results of 33 cm^{-1} obtained by Butty *et al.*²⁰ in sol-gel-derived CdS QDs. The significant enhancement of the optical gain can be attributed to the stronger three-dimensional quantum confinement in our samples, which is expected to enhance the optical nonlinearities.²¹ The difference in the quantum confinement can be seen if we compare the normalized dot radius $R/a_B = 1.3$ for the sample used in Ref. 20 to $R/a_B = 0.15$ in the PbS quantum-dot-doped glass used in the experiments presented here.

In Fig. 3, we illustrate the spectral dependence of the optical gain. Here, the spectral positions of the pump and probe pulses, as used in the experiments of Fig. 2(b), are compared to the linear absorption of the sample [Fig. 3(a)] and the calculated $1s$ – $1p$ splitting [Fig. 3(b)]. When pumping into the $1p$ -absorption resonance, optical gain is found only when the pump-probe detuning is close or equal to the $1s$ – $1p$ splitting for a given dot size. More explicitly, gain is found only for pumping at 900 nm and probing at 1317 nm [see curve (1) of Fig. 2(a)] or for pumping at 980 nm and probing at 1352 nm [see curve (1) of Fig. 2(b)]. However, no gain is observed for pumping at 980 nm and probing at 1317 nm [see curve (2) of Fig. 2(a)] or for pumping at 900 nm and probing at 1352 nm. Since in our experiments, the inhomogeneous broadening of the $1p$ absorption is large compared to the spectral width of the pump pulse, the pump pulse only excites a small portion of all QDs for which the pump pulse

is resonant with the $1p$ transition. Optical gain is only possible in this subset of pump-pulse-selected QDs and will appear around the transition frequencies between the lowest confined electron and hole levels after relaxation of the excited electron-hole pairs to the exciton ground state. In this case, taking into account our dot-size-selective excitation, as the mismatch between the $1s-1p$ splitting and the pump-probe detuning increases, the ratio between gain from the excited QDs and absorption from the nonexcited QDs decreases. Consequently, the effective optical gain disappears if this mismatch is too large, as demonstrated in curve (2) in Fig. 2(a). Here, the detuning of the probe pulse from the maximum gain position is about 35 nm, which corresponds to the initial spectral distribution of the excited carriers given by the pump pulse.

So far, we neglected the Coulomb and spin-orbit interactions, which change the selection rules and, therefore, increase the number of dipole-allowed transitions.²² These additional transitions give rise to optical gain at new frequencies. For example, Hu *et al.*²³ used a microscopic model to describe optical gain observed in CdS and CdSe QDs.^{20,24} Here, the gain was spectrally broad and located on the low-energy side of the ground exciton absorption. This was explained by taking into account one and two electron-hole pair recombination, i.e., excitons and biexcitons. On the basis of these results, we would expect maximum gain at a pump-probe detuning slightly larger than the $1s-1p$ splitting when pumping into the $1p$ transition. Our limited spectral resolution and accuracy of the HB model to calculate the $1p$ -transition energies does not allow a final conclusion whether or not a shift or increased gain bandwidth due to the biexcitonic contributions is present. Furthermore, our pronounced dot-size-dependent selective excitation prohibits a direct comparison of our experimental results and the earlier experiments on CdS and CdSe QDs.

In summary, PbS quantum-dot-doped glasses show room-temperature gain of 80 cm^{-1} in the communication wavelength region around $1.3 \mu\text{m}$. We have shown that we can utilize the strong quantum-confinement effect and the inhomogeneous broadening due to dot-size fluctuations to tune the gain spectrum of an individual PbS-doped glass. Furthermore, because the exciton ground-state transitions, where the gain is observed, can be widely tuned by changing the size of the QDs, we believe that room-temperature optical gain in PbS quantum-dot-doped glasses is possible over a wide spectral range. This gain tunability together with the demonstrated possibility of optical pumping at the wave-

length of commercially available laser diodes shows that PbS quantum-dot-doped glasses may be suitable low-cost alternatives to current amplifiers and lasers for optical communication applications.

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Fabrication of germanium-coated nickel hollow waveguides for infrared transmission

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Circular hollow nickel waveguides with an inner germanium layer are fabricated by using a method based on rf sputtering, plating, and etching techniques. Transmission losses less than 0.5 dB are achieved including launching losses for straight waveguides with $1.5 \text{ mm}\phi \times 1 \text{ m}$ at $10.6\text{-}\mu\text{m}$ wavelength. Bending losses of the waveguides are also examined.

PACS numbers: 42.80.Lt, 42.60.Kg, 84.40.Ts, 84.40.Vt

Realization of flexible waveguides for CO_2 laser light becomes an important subject in industrial and medical applications. From the viewpoint of power handling capability, hollow-core waveguides seem to be suitable for practical use¹ and several kinds of hollow waveguides are proposed and fabricated such as so-called flexible-infrared transmissive waveguides,^{1,2} helical-circular waveguides,³ circular metallic waveguides,⁴ dielectric-coated metallic waveguides,^{5,6} and hollow glass waveguides.⁷ Among the various hollow-core waveguides, the dielectric-coated metallic waveguide which we proposed can transmit the HE_{11} mode with low loss and high launching efficiency from commercially available CO_2 lasers and its transmission characteristics are fully analyzed in slab^{8,9} and circular waveguides.^{5,10,11} In this letter we report the first fabrication of the dielectric-coated metallic waveguides with relatively small transmission losses.

To fabricate the dielectric-coated metallic waveguides, two points were the main concern. One is the adhesion between the metal and dielectrics, and another is to form waveguides efficiently. In order to solve these problems, we developed a method based on rf sputtering, plating, and etching techniques as schematically shown in Fig. 1. As first trial, we select germanium as a transparent dielectric at $10.6\text{-}\mu\text{m}$ wavelength. Nickel is chosen as a metal because its plating techniques are well established¹² and also the thickness of several hundred microns of metal which makes waveguides mechanically strong can be obtained by plating. These two materials are also preferable because they are free from toxicity in our fabrication process of waveguides.

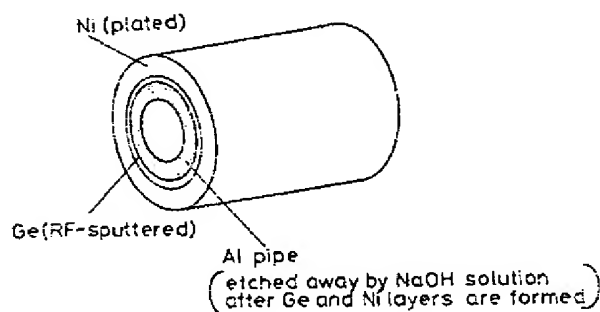


FIG. 1. Schematic method to fabricate germanium-coated nickel waveguides.

First, germanium is sputtered on a polished aluminum pipe (outer diameter of 1.5 mm, thickness of 0.2 mm, and length of 1.2 m) which is rotated with 10 rpm and moved axially with 10 cm/min in a vacuum chamber (2.0×10^{-2} Torr of argon gas) to obtain a uniform thin layer. Schematic view of the sputtering apparatus is shown in Fig. 2. Sputtered germanium on the aluminum pipe is observed to be amorphous according to the x-ray analysis, although an original material is single crystal. The sputtering rate is around $0.15 \mu\text{m/h}$ for 500-W rf power. A typical thickness of the germanium layer d is designed to be $0.45 \mu\text{m}$ so as to satisfy the minimum loss condition^{10,11}:

$$(a^2 - 1)^{1/2} k_0 d = \pm \tan^{-1} \left(\frac{a}{(a^2 - 1)^{1/4}} \right) + s\pi,$$

where k_0 is the wave number in free space, a is the refractive index of the dielectric, i.e., 4.0 for germanium, and s is the integer.

In order to form a metallic pipe, nickel is directly plated onto the germanium layer. An anode of a nickel plate is formed to a circular cylinder whose diameter is 7.5 cm and length is 1.2 m and a cathode for plating is located at its center to form nickel uniformly. The plating bath is composed of $\text{NiSO}_4 \cdot 6\text{H}_2\text{O}$ (240 g), $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ (45 g), and H_3BO_3 (30 g) in water of 1000 cm^3 in ratio. Although the pH changes from 1.3 to 3.7 during plating of several pieces of waveguides, a nickel layer of $70\text{--}200 \mu\text{m}$ is formed with sufficient mechanical strength. The current is 50 mA at an early stage of plating to form a pipe with a smooth surface and is gradually increased up to 2 A to form a pipe efficiently. Finally, the aluminum pipe is etched away by 10–20% NaOH solution and a circular hollow nickel waveguide with an inner germa-

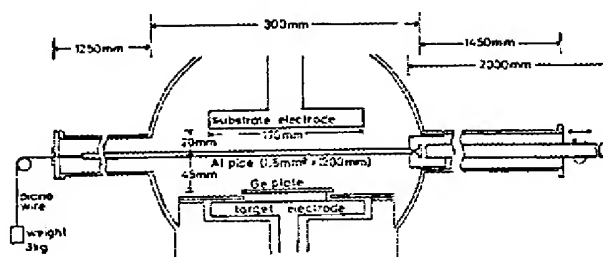


FIG. 2. Schematic view of the sputtering apparatus to form a germanium layer.

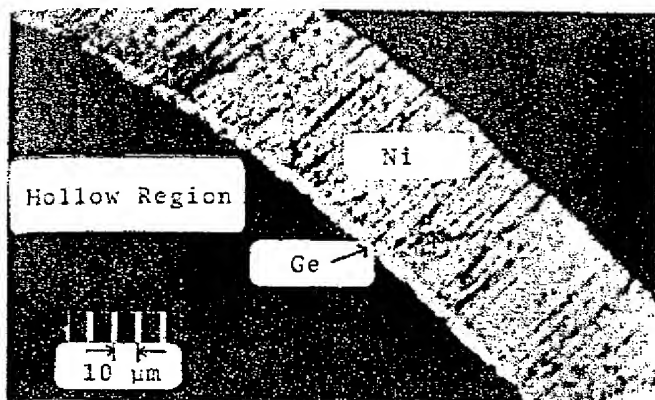


FIG. 3. Cross-sectional view of a germanium-coated nickel waveguide.

mium layer is fabricated. The cross-sectional view of the waveguide is shown in Fig. 3, where the germanium layer is made rather thick to show the adhesion between the nickel and germanium layers.

Several germanium-coated nickel waveguides are fabricated. Total transmission losses and polarization properties of the straight waveguides are measured and summarized in Table I as well as those of various hollow waveguides at 10.6 μm for comparison. Excitation of waveguides is conducted through a 20-cm-long waveguide with the same diameter of the test one by focusing a laser beam with a ZnSe lens ($f = 127$ mm). The diverging angle of the output beam from waveguides is around 16–21 mrad depending on waveguides, which shows that several higher order modes propagate, for the theoretical one of the HE_{11} mode is 10.8 mrad. It is seen that the losses of metallic waveguides can be reduced significantly by coating a dielectric material. Even when the laser beam is directly focused to the germanium-coated nickel waveguides, total transmission losses including launching losses increase by only 0.05 dB and never exceed 0.5 dB for 1-m-long waveguides. By connecting two waveguides carefully, total loss of 1 dB is achieved for 2-m-long waveguide. This

TABLE I. Transmission properties of various hollow straight waveguides. Waveguides denoted by * are commercially available. Others are fabricated by the techniques presented in this paper.

Hollow waveguides	Dimensions (mm ϕ \times m)	Total loss (dB)	Degree of polarization (%)
Pyrex glass	1.5 \times 1.01*	1.22	99.1
Teflon	1.6 \times 1.01*	4.42	99.1
Aluminum	1.6 \times 1.01*	2.26	88.9
Nickel	1.6 \times 1.01*	2.34	83.9
	1.5 \times 1.02	2.85	81.7
Germanium-coated nickel	1.5 \times 1.03	0.35	95.6
	1.5 \times 1.01	0.40	98.2
	1.5 \times 1.01	0.41	97.1
	1.5 \times 1.01	0.43	93.2

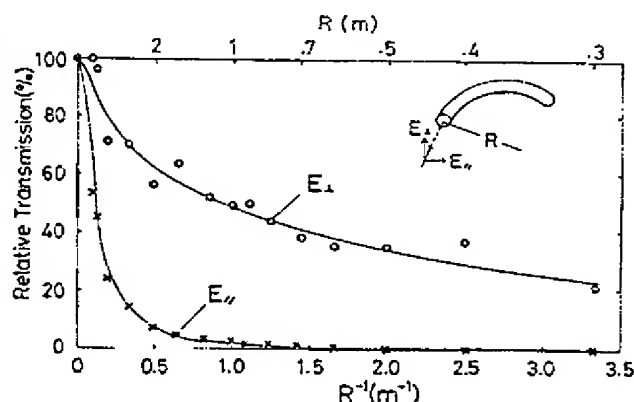


FIG. 4. Bending losses of a nickel hollow waveguide with 1.5-mm diameter and 1.02-m-long whose straight-waveguide loss is 2.85 dB.

fact suggests that a waveguide with length of several meters can be realized by using the present fabrication techniques and equipment. Although a theory^{10,11} predicts that the power loss of the HE_{11} mode is 3.2×10^{-2} dB/m, the difference between theoretical and present experimental results may be caused by roughness or gradual deformation of the waveguides and deviation of the thickness of the germanium layer from the optimum one which is not completely controlled at present.

Bending losses of nickel waveguides without and with an inner germanium layer are measured and shown in Figs. 4 and 5, where the first 20 cm is made straight and the rest of 80 cm of the waveguides is bent with a uniform curvature. When the polarization is parallel to the plane of curvature (denoted by E_{\parallel}), transmission is extremely small for the nickel waveguide even for large bending radii, whereas large improvement is done for the germanium-coated waveguide as well as when the polarization is perpendicular. If the dielectric-coated waveguides with smaller losses are realized in straight structures, the waveguides can be much sharply bent without significant loss increase.

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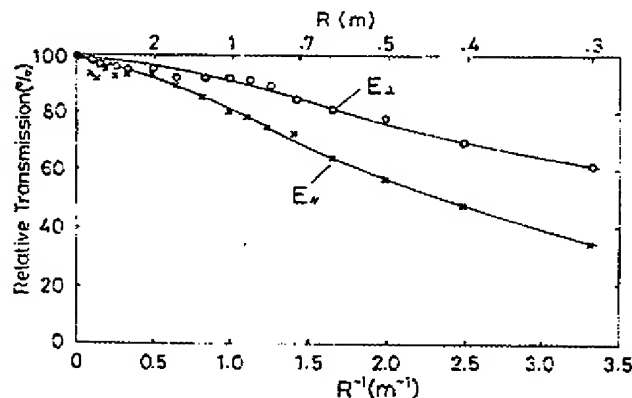


FIG. 5. Bending losses of a germanium-coated nickel waveguide with 1.5-mm diameter and 1.01-m-long whose straight-waveguide loss is 0.43 dB.

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Transverse second-order mode oscillations in a twin-stripe laser with asymmetric injection currents

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Transverse second-order mode oscillations in asymmetrically pumped twin-stripe lasers are described. Both far-field patterns and near-field patterns have two peaks, and a stronger peak appears on the weakly pumped side. The dependence of light intensity in one peak of the far-field pattern on the current into one stripe shows threshold characteristics. On application of rectangular pulse to one stripe, switching from one peak to the other occurs within 5 ns.

PACS numbers: 42.55.Px, 42.80.Lt, 42.60.Da, 42.60.He

Asymmetric waveguide structures in semiconductor lasers cause asymmetry in far-field radiation patterns. Asymmetric far-field patterns were observed for lasers with a built-in asymmetric structure,¹ and were also observed for lasers with symmetric built-in structures using asymmetric distribution of injection currents.^{2,3} The asymmetric far-field patterns were explained on the basis of Maxwell's equation for an asymmetric distribution of complex refractive indices.^{1,2,4} However, investigations have so far been limited to transverse fundamental mode oscillations in an asymmetric waveguide, and very little is known about higher-order mode oscillations.³ In the present letter, we report experimental results and some theoretical insight about transverse second-order mode oscillations in semiconductor lasers with asymmetric injection currents.

The device structure is similar to the one reported by Seifres *et al.*² The wafers were grown by liquid phase epitaxy and consisted of a Si-doped ($2 \times 10^{18} \text{ cm}^{-3}$) GaAs substrate, 2- μm -thick Te-doped ($1 \times 10^{17} \text{ cm}^{-3}$) $\text{Al}_{0.35}\text{Ga}_{0.65}\text{As}$ layer, 0.2- μm -thick undoped active $\text{Al}_{0.05}\text{Ga}_{0.95}\text{As}$ layer, 2- μm -thick Ge-doped ($5 \times 10^{16} \text{ cm}^{-3}$) $\text{Al}_{0.35}\text{Ga}_{0.65}\text{As}$ layer, and 0.3- μm -thick Ge-doped ($1 \times 10^{19} \text{ cm}^{-3}$) GaAs. A 12- μm -wide mesa was formed by etching down to 0.5 μm above the active layer. After depositing a 0.1- μm -thick Si_3N_4 film, 2- μm -wide twin stripes on 8- μm centers was etched in the film

on the mesa. The *p*-side contacts were Cr and Au. The contact between the stripes were chemically etched. 200- μm -long chips were cleaved. They were pumped with 1- μs width rectangular pulses. The total threshold current was about 60

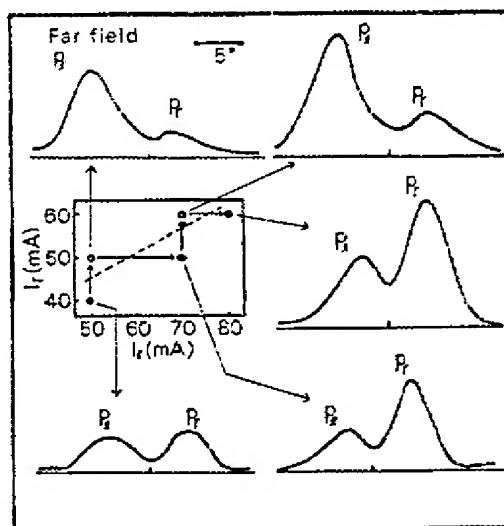


FIG. 1. Far-field patterns of a twin-stripe laser. I_1 and I_2 are currents injected into the left and right stripes. The inset shows currents at which the patterns were observed. Open circles mean that the left peak of the far-field pattern is higher than the right peak at these currents. Solid circles mean the opposite. The broken line approximately determines two current regions corresponding to the left peak (or right) peak being higher. The arrows connecting the circles show how the currents were varied.

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A Review of IR Transmitting, Hollow Waveguides

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Infrared (IR) transmitting hollow waveguides are an attractive alternative to solid-core IR fibers. Hollow guides are made from plastic, metal, or glass tubes that have highly reflective coatings deposited on the inside surface. These guides have losses as low as 0.1 dB/m at 10.6 μm and may be bent to radii less than 5 cm. For use in high-power laser delivery applications, the guides have been shown to be capable of transmitting up to 3 kW of CO_2 laser power. They are also finding uses in both temperature and chemical fiber sensor applications. This paper reviews the progress in hollow waveguide technology with emphasis on the best guides available today.

Keywords IR fibers, hollow waveguides, laser power delivery, fiber sensors

Infrared (IR) optical fibers may be defined as fiber optics transmitting wavelengths greater than approximately 2 μm . The first IR fibers were fabricated in the mid 1960's from chalcogenide glasses such as arsenic trisulfide with losses in excess of 10 dB/m [1]. During the mid 1970's, the interest in developing an efficient and reliable IR fiber for short-haul applications increased partly in response to the need for a fiber to link broadband, long wavelength radiation to remote photodetectors in military sensor applications. In addition, there was an ever-increasing need for a flexible fiber delivery system for transmitting CO_2 laser radiation in surgical applications.

Around 1975, a variety of IR materials and fibers were developed to meet these needs. These included the heavy metal fluoride glass (HMFG) and polycrystalline fibers as well as hollow rectangular waveguides. While none of these fibers had physical properties even approaching that of conventional silica fibers, they were, nevertheless, useful in lengths less than 2-3 m for a variety of IR sensor and power delivery applications [2].

Infrared fiber optics may logically be divided into three broad categories: glass, crystalline, and hollow waveguides. These categories may be further subdivided, based on either the fiber material or structure or both, as shown in Table 1. Over the past 25 years, many novel IR fibers have been made in an effort to fabricate a fiber optic with properties as close to silica as possible, but only a relatively small number have survived. A good source of general information on these various IR fiber types may be found in the literature [3-6]. In this paper, the hollow waveguide technology will be reviewed only with emphasis on the best and most

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Table 1

Three main categories of IR fibers with examples of the most common type of fiber within the category of glass, crystalline, or hollow fibers

Main	Subcategory	Examples
Glass	Heavy metal fluoride (HMFG)	ZBLAN-(ZrF ₄ -BaF ₂ -LaF ₃ -AlF ₃ -NaF)
	Germanate	GeO ₂ -PbO
	Chalcogenide	As ₂ S ₃ and AsGeTeSe
Crystal	Polycrystalline (PC)	AgBrCl
	Single crystal (SC)	Sapphire
Hollow waveguide	Metal/dielectric film	Hollow glass waveguide
	Refractive index < 1	Hollow sapphire at 10.6 μ m

practical hollow waveguide candidates available today. In general, both the optical and mechanical properties of IR transmitting hollow waveguides remain inferior to silica fibers and, therefore, the use of hollow guides is still limited to nontelecommunication, short-haul applications requiring only a few meters of waveguide, rather than kilometer lengths common in telecommunication applications. The short-haul nature of these special IR fibers results from the fact that the guides have losses in the range of a few dB/m rather than a few dB/km. Also, hollow guides have an additional loss on bending, and they are also somewhat weaker than silica fiber. These deleterious features have slowed the acceptance of hollow guides and restricted their use today to applications in chemical sensing, thermometry, and laser power delivery.

Background

Hollow waveguides present an attractive alternative to other solid-core IR fibers [2]. Key features of hollow guides are their ability to transmit wavelengths well beyond 20 μ m, their inherent advantage of having an air core for high-power laser delivery, and their relatively simple structure and potential low cost. Initially, these waveguides were developed for medical and industrial applications involving the delivery of CO₂ laser radiation, but more recently, they have been used to transmit incoherent light for broadband spectroscopic and radiometric applications [7, 8]. In general, hollow waveguides enjoy the advantages of high laser power thresholds, low insertion loss, no end reflection, ruggedness, and small beam divergence. Potential disadvantages, however, include an additional loss on bending and a small NA. Nevertheless, today they are one of the best alternatives for both chemical and temperature sensing as well as for power delivery in IR laser surgery or in industrial laser delivery systems with losses as low as 0.1 dB/m and transmitted cw laser powers as high as 2.7 kW [9].

Hollow-core waveguides may be grouped into two categories: 1) those whose inner core materials have refractive indices greater than one (leaky guides) and 2) those whose inner wall material has a refractive index less than one (attenuated total reflectance, i.e., guides). Leaky or $n > 1$ guides have metallic and dielectric films deposited on the inside of metallic, plastic, or glass tubing [10]. Attenuated

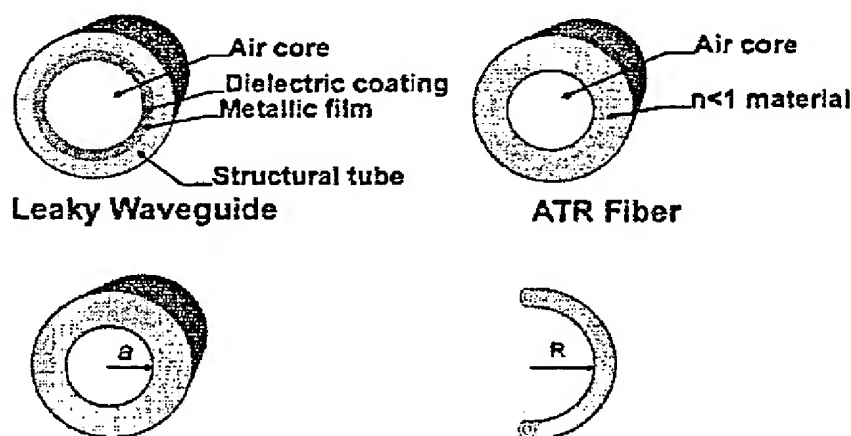


Figure 1. Structure of two types of hollow waveguides also showing the key parameters of bore and bending radius affecting the losses in hollow guides.

total reflectance guides are composed of dielectric materials with refractive indices less than one in the wavelength region of interest [11]. Therefore, $n < 1$ guides are fiberlike in that the core index ($n \approx 1$) is greater than the clad index. Hollow sapphire fibers operating at $10.6 \mu\text{m}$ ($n = 0.67$) are an example of this class of hollow guide [12]. The structure and key parameters for hollow guides are shown in Figure 1. In general, hollow structures with $n > 1$ have been made from metal, plastic, and glass tubes while the $n < 1$ or ATR guides are made of sapphire or some special $n < 1$ oxide glass.

The theory of hollow waveguide transmission has been described from the viewpoint of both wave and ray optics. Marcatili and Schmeltzer (MS) [13] have used a wave optic approach that predicts for either metallic or dielectric waveguides that $\alpha \sim 1/a^3$, where α is the attenuation coefficient and a is the bore radius. Bending the hollow waveguides increases the total loss. Recently, Miyagi et al. [14] have shown that the additional bending loss varies as $1/R$, where R is the bending radius. Therefore, we have in contrast to the solid-core fibers, a loss that depends strongly on the diameter and bending radius of the fiber. For the thin film waveguide structures, Miyagi and Kawakami [15] have shown that for dielectric coatings deposited over a metallic layer, the attenuation coefficient, is given by

$$\alpha_z = \left(\frac{U_o}{2\pi} \right)^2 \cdot \frac{\lambda^2}{a^3} \cdot \left(\frac{n}{n^2 + k^2} \right)_{\text{metal}} \cdot F_{\text{film}}, \quad (1)$$

where α_z is the loss for a straight guide; U_o is a mode-dependent parameter that for the lowest order HE_{11} mode equals 2.405; n and k in $(\dots)_{\text{metal}}$ refer to the optical constants of metal film; and F_{film} is a term that accounts for the loss due to the dielectric film(s).

Metal-Tube Waveguides

Hollow core waveguides have been fabricated using a variety of techniques. Some of the methods include physical vapor deposition of silver and dielectric layers on metallic substrates [16], sputtering of metallic, dielectric, and semiconductor films

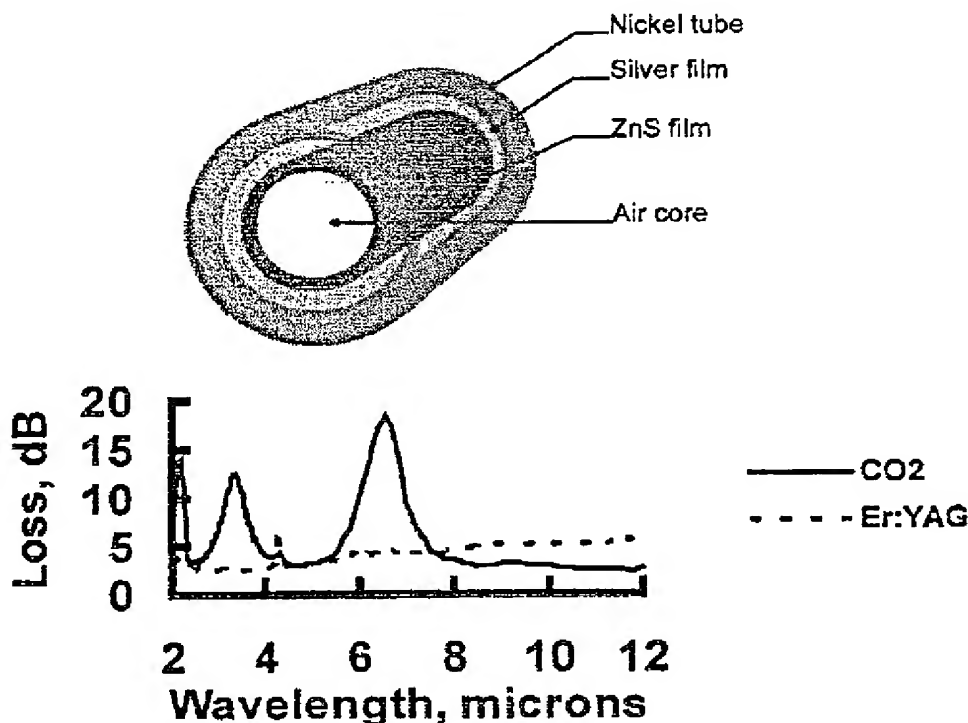


Figure 2. Structure and spectral losses for metallic waveguide fabricated by Miyagi's group at Tohoku University [22].

on a leachable mandrel followed by electroplating [17], and liquid phase formation of coatings inside plastic tubing [18], and glass tubing [19]. Most often the cross section of the guides is circular, but early work by Garmire et al. [20] and more recently by Kubo [21] on rectangular guides continues to be of interest. The advantage of the circular cross section is the ease of bending and the small overall size compared to rectangular or square cross section guides.

Professor Miyagi and his coworkers at Tohoku University [17] have pioneered the development of metallic waveguides based on a hollow nickel substrate. Their fabrication process involves three steps. In the first step, a pipe made typically of aluminum is placed in a sputtering chamber and a dielectric layer followed by a metallic film is deposited on the pipe. Next, the coated pipe is put into an electroplating tank, where a thick nickel layer is deposited on top of the sputtered layers. Finally, the pipe is etched away leaving the final structure shown in Figure 2.

In Figure 2, we also show a typical loss curve for one of their best dielectric coatings (ZnS) over silver. The data shown are for two 1000 μm -bore guides, one optimized for the 3 μm wavelength of the Er:YAG laser and the other optimized for the 10.6 μm wavelength of the CO₂ laser. The optimization for each wavelength results from adjusting the thickness of the thin-film dielectric coating. In Figure 3, bending losses for these hollow waveguides are given for both the CO₂ and Er:YAG laser wavelengths [22]. The losses are seen to be as low as 0.25 dB/m at 10.6 μm for light polarized perpendicular to the plane of bending, but slightly higher for parallel polarized light. This is as expected from waveguide theory or

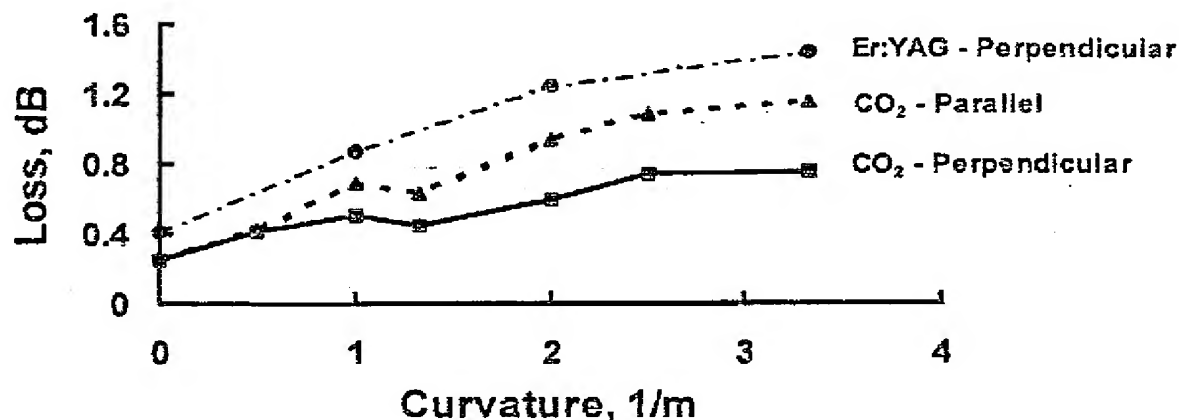


Figure 3. Losses for bent hollow metallic guides taken at both CO₂ and Er:YAG laser wavelengths [22].

from simple considerations based on Fresnel reflections from metal surfaces. The highest CO₂ laser power delivered using a 2000 μm bore metallic guide is over 3 kW [9]. Miyagi and his coworkers have also developed a hollow structure based on a square cross section [23]. To fabricate a square cross-section tube they have developed a process in which they first deposit, using evaporative techniques, thin-film coatings of ZnS, PbTe, and/or PbF₂ on phosphor bronze strips and then they solder four of these phosphor bronze metal strips together in a continuous process. The losses for these square guides are as low as 0.1 dB/m at 10.6 μm .

The waveguides developed by Morrow et al. [24] are constructed from a silver tube. Instead of depositing a metallic layer inside a hollow mandrel, they begin with an extruded silver tube and then deposit a silver halide film on the inside of the tube as shown in Figure 4. To ensure the lowest loss, Morrow et al. first etch the bore of the silver tubing to make it smooth. Then an AgBr film is applied on the inside using wet chemistry methods. The bending loss for a 1000 μm -bore tube, 1-m in length at 10.6 μm is also shown in Figure 4. It is noted from this data that the losses in the straight guide are quite low. Unfortunately, the guides tend to mix modes and, therefore, the beam output is rather multimode compared to the hollow glass or sapphire waveguides. This is due in large part to the rougher inner surface of the extruded tubing compared, for example, to the smooth inner surface of glass.

Luxar's approach is based on technology initially developed by Laakmann and her colleagues [16]. Their fabrication technique involves first depositing a silver film on a metal strip and then overcoating the silver with a thin film of PbF₂. The metal strip is then rolled and inserted into stainless steel hypodermic tubing as shown in Figure 5. The bending loss for one of their guides is also shown in Figure 5. This waveguide had a 750 μm bore, 1-m length and the losses shown are for 10.6 μm . Like the other hollow waveguides, the losses are quite low at CO₂ laser wavelengths. Furthermore, the 1/R behavior of the loss on bending is noted. This is, of course, characteristic of all hollow guides, but the magnitude of this loss depends largely on the quality of the inner surface. In general, these bending losses are tolerable for the radii normally encountered in practical applications.

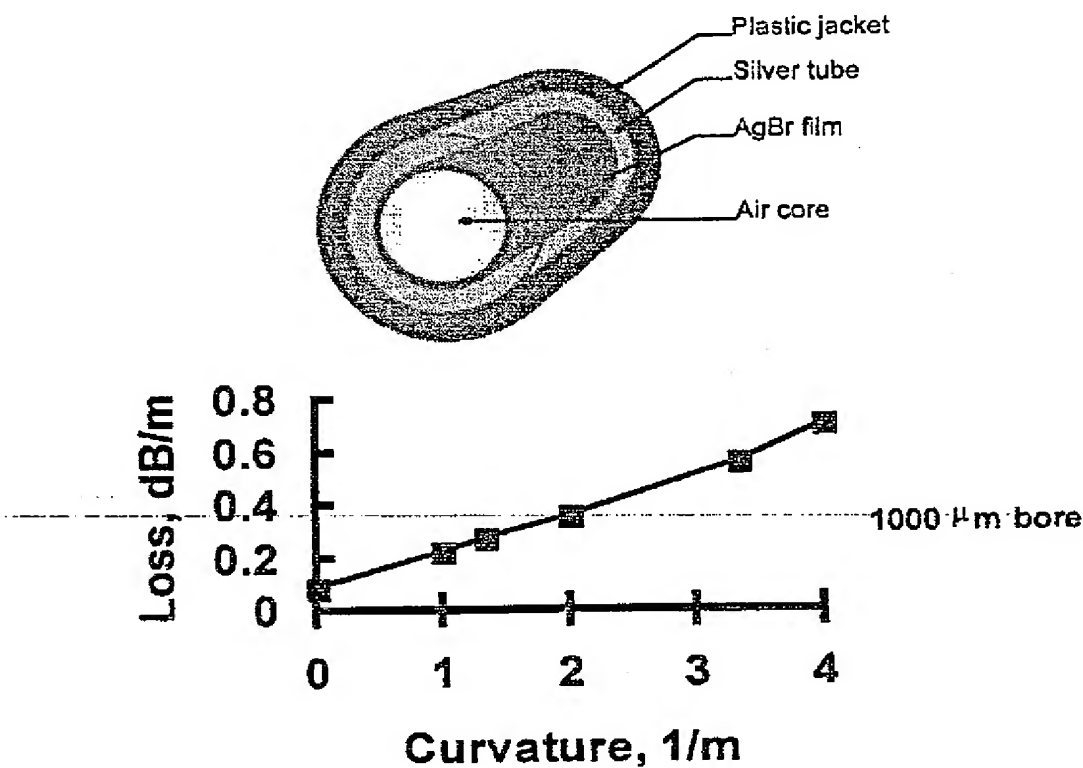


Figure 4. Structure and bending loss for hollow guide made from silver tubing with a AgBr film deposited on the inside surface [24].

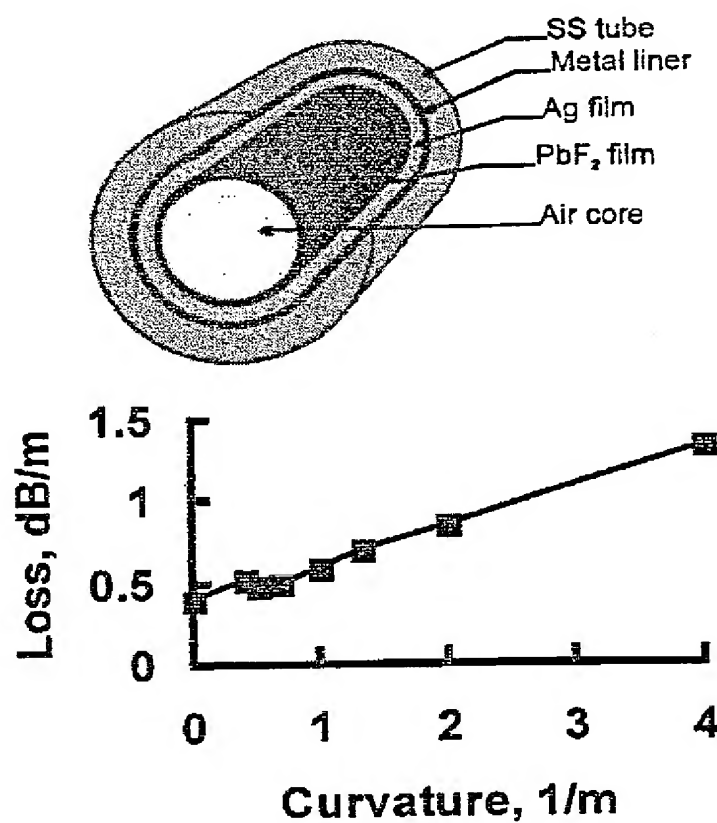


Figure 5. Structure and bending loss for hollow guides made by Luxar [16].

Plastic-Tube Waveguides

Hollow waveguides may also be formed on the inside of plastic tubing. This leads to a very flexible structure that is inexpensive to fabricate and durable enough that a reasonable laser power (safe limit < 25 W) can be transmitted through the guides. Croitoru's group at Tel Aviv University has pioneered the approach of first depositing a silver film on the inside of Teflon and polyethylene plastic tubing (see Figure 6) and then overcoating the silver film with a dielectric layer of AgI [18] using wet-chemistry techniques [25]. Initially, Croitoru used rather large bore tubing, but more recently guides with bore sizes of about $1000\text{ }\mu\text{m}$ have been fabricated. In Figure 6 the bending loss is shown for a $1000\text{ }\mu\text{m}$ bore tube measured at $10.6\text{ }\mu\text{m}$. The losses are somewhat higher than those measured for the metallic tubes. This is primarily due to increased scattering losses resulting from the rougher inner surface of the plastic tubing. In an independent measurement, the author was able to transmit over 65 W of CO_2 laser power for several minutes through an $1850\text{ }\mu\text{m}$ bore guide.

Losses in the plastic waveguides made by Croitoru may be reduced if smoother polymer tubing is chosen. Haan and Harrington [26] have used similar wet-chemistry methods to deposit Ag/AgI films inside polycarbonate tubing. Extruded polycarbonate tubing in lengths of 2 m and bore sizes from $840\text{ }\mu\text{m}$ to 2 mm was used to make the guides. Figure 7 shows the lowest straight losses for several hollow waveguide structures—two polycarbonate waveguides, a Teflon waveguide,

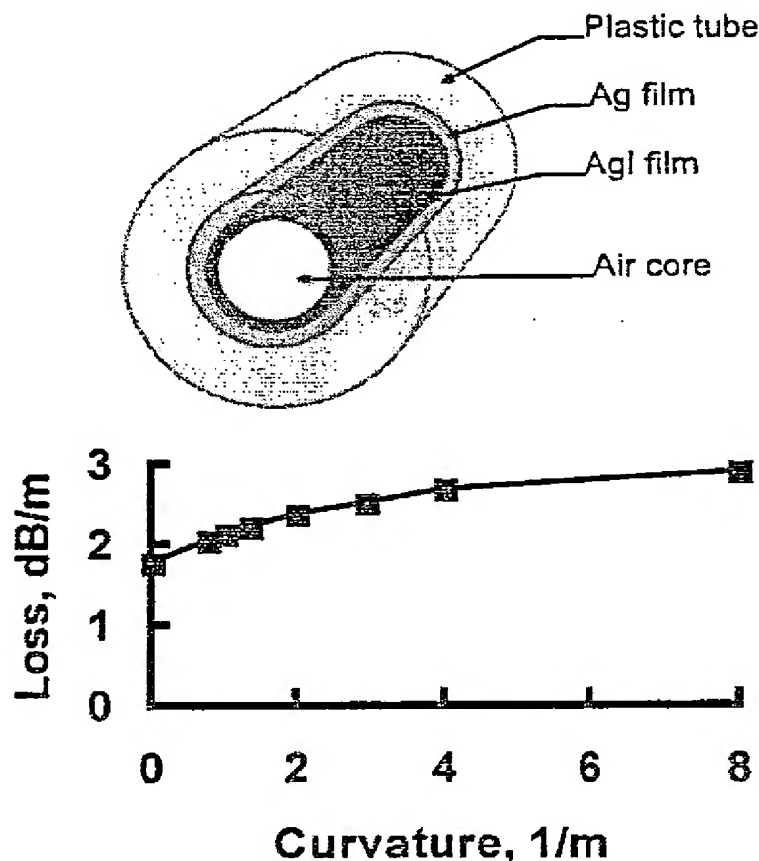


Figure 6. Structure and bending loss for hollow plastic waveguide made by Croitoru's group at Tel Aviv University [18].

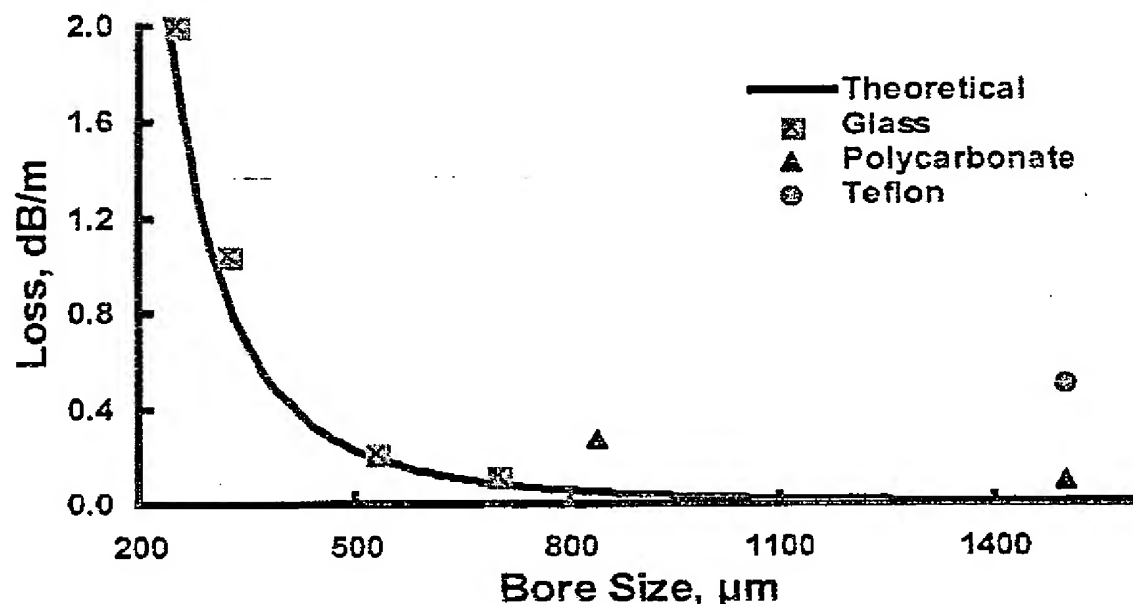


Figure 7. Composite data for hollow glass and plastic waveguides. Note the low loss at 10.6 μm for the glass guides in comparison to the plastic guides.

and hollow glass waveguides (see next section) measured using a CO_2 laser. The solid curve is calculated for the lowest order mode at 10.6 μm using Eq. 1. The attenuation values of 0.22 dB/m for an 840 μm PC guide and 0.10 dB/m for a 1500 μm PC guide show the improvement over the Teflon or polyethylene waveguides. In a simple power test using a CO_2 laser, the 1500 μm waveguide was able to withstand more than 25 W of incident laser power.

Hollow Glass Waveguides

One of the most popular hollow waveguides today is the hollow glass waveguide (HGW) developed by Harrington's group at Rutgers University. This hollow glass structure has the advantage over other hollow structures because it is simple in design, extremely flexible, and, most important, has a very smooth inner surface. Hollow glass waveguides have a metallic layer of Ag on the inside of silica glass tubing and then a dielectric layer of AgI over the metal film identical to that used to make the hollow plastic guides. Figure 8 shows a cross section of the structure of the HGWs. The fabrication of HGWs begins with silica tubing, which has a polymer (UV acrylate or polyimide) coating on the outside surface. A wet-chemistry technique (see Figure 9), similar to that used by Croitoru and his coworkers [18, 27] to deposit metal and dielectric layers on the inside of plastic tubing, is employed to first deposit a silver film using standard Ag plating technology [28]. Next, a very uniform dielectric layer of AgI is formed through an iodization process in which some of the Ag is converted to AgI [29]. Using these methods, HGWs with bore sizes ranging from 250 to 1000 μm and lengths as long as 13 m have been made.

The spectral response for HGWs depends critically on the thickness of the dielectric film. Generally, for the AgI films, the film thickness ranges from 0.2 to

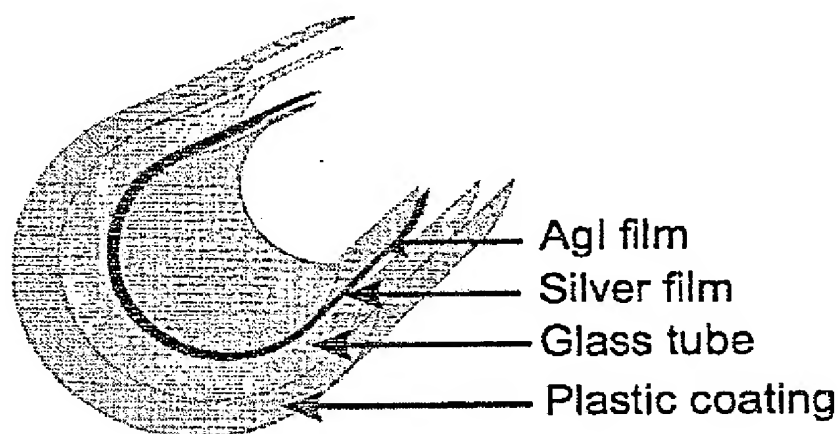


Figure 8. Structure of the HGWs showing the metallic and dielectric films deposited inside silica glass tubing.

0.8 μm . In Figure 10, we show the spectral response of two waveguides that have different thickness films deposited on the inside of a 700 μm bore silica tube, 1 m in length. The thickest film gives a minimum loss at 10.6 μm , while the thin film was selected for minimum loss near 3 μm . The latter guide has a fairly flat response beyond 3 μm and, therefore, this guide would be useful in broadband applications. The structure observed in the spectra is due to thin-film interference effects similar to that commonly observed in thin-film coatings on optical components. These effects have been observed and extensively discussed in the work of Matsuura et al. [30].

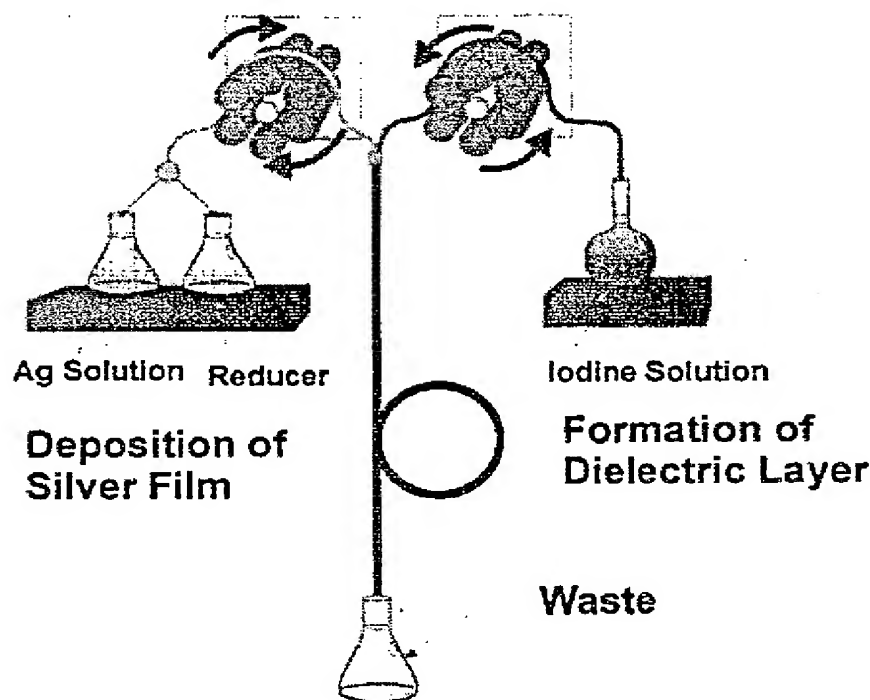


Figure 9. Schematic of the experimental set-up for depositing the Ag metallic and AgI dielectric films inside silica tubing to form the HGWs.

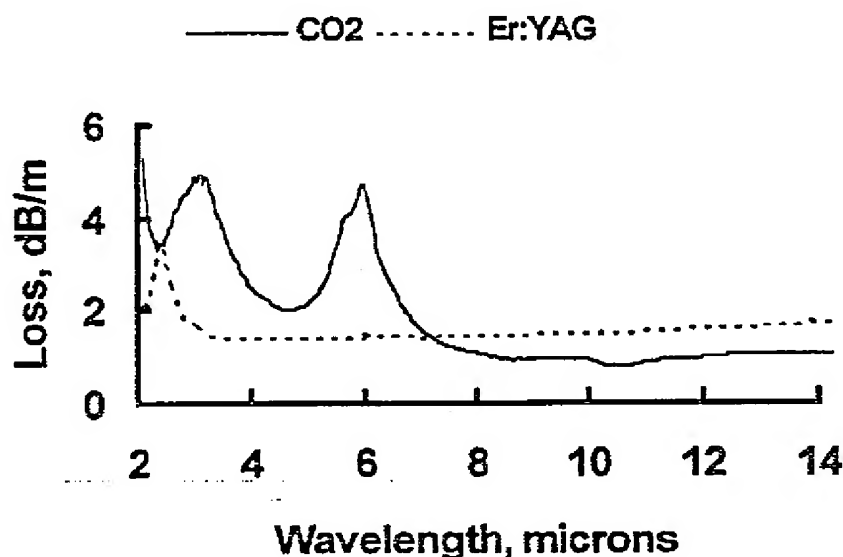


Figure 10. Spectral response of two HGWs—one designed for low loss at the CO₂ laser wavelength of 10.6 μm and the other for low loss at the Er:YAG laser wavelength of 2.94 μm .

The strong bore-size dependent loss for straight HGWs is shown for two guides in Figure 11. [28]. These data were taken using CO₂ and Er:YAG lasers and the guides were optimized for minimal loss at 10 and 3 μm , respectively. The solid curves are theoretical calculations of the losses for the lowest order HE₁₁ mode. At the CO₂ laser wavelengths, one can see not only the strong $1/a^3$ dependence predicted by MS theory, but also that there is good agreement with the experimental results. However, at 3 μm the calculated losses are much lower than the measured values. This is a result of increased scattering losses at the shorter wavelengths and the multimode character of the Er:YAG laser.

Bending increases the loss in hollow waveguides beyond that shown for the straight loss seen in Figure 12. The additional bending loss varies as $1/R$ as

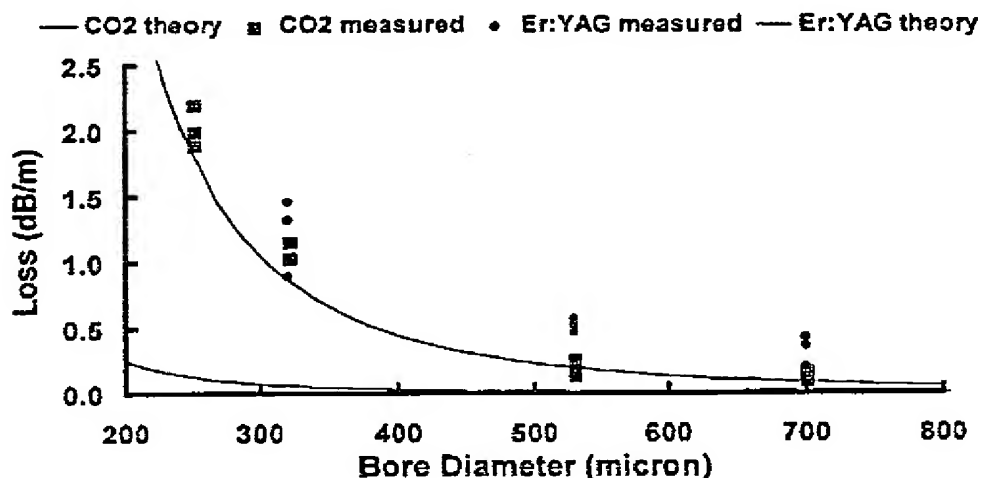


Figure 11. Measured losses for straight HGWs using CO₂ and Er:YAG lasers. Note that the predicted losses are well below the measured ones at 2.94 μm .

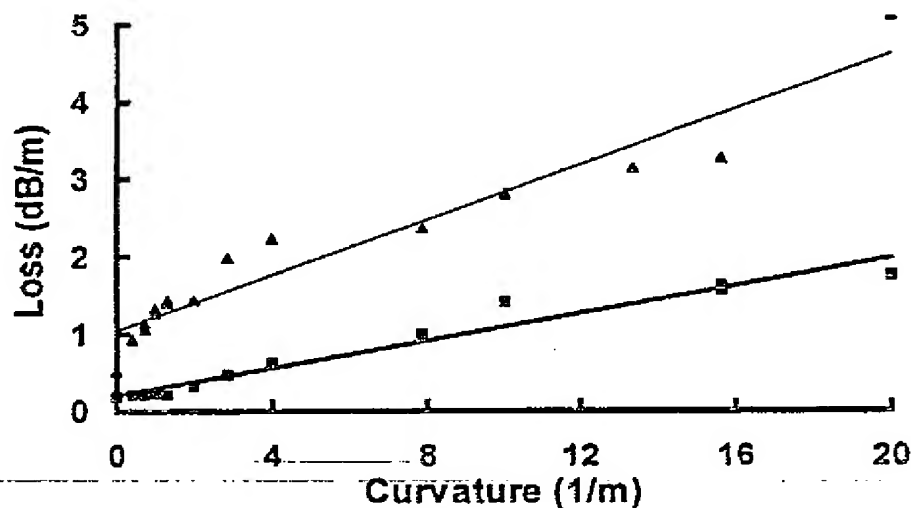


Figure 12. Bending losses for two 530- μm -bore HGWs measured at 10.6 (■) and 2.94 (▲) μm .

reflected in the data for two 530 μm bore guides in Figure 12. These data show the total loss for guides with a constant length of fiber under bending. A curvature of 20 represents a bend diameter of only 10 cm. This is sufficiently small for most applications.

Hollow $n < 1$ Waveguides

The idea of an $n < 1$ structure originated with Hidaka et al. in 1981 [31]. In this structure, the air core ($n = 1$) has a refractive index greater than the inner wall, cladding material, and, therefore, this type of waveguide is fiberlike in that $n_{\text{clad}} < n_{\text{core}}$. This is also referred to as an attenuated total reflectance (ATR) guide in contrast to the leaky structure of the $n > 1$ guides discussed above (see Figure 1). To be useful for laser transmission, the ATR guides must have the region of anomalous dispersion, where n is less than 1, fall within some useful laser wavelength range. The first $n < 1$ guides studied by Hidaka et al. [32] focused on glass tubes made from lead and germanium-doped silicates. By adding heavy ions to silica glass, he was able to shift the infrared edge to longer wavelengths so that the $n < 1$ region of anomalous dispersion occurred within the CO_2 laser wavelength band. Worrell [33] also studied $n < 1$ glasses, in particular, the germanate glasses. The losses in the hollow glass, $n < 1$ fibers, however, were very high due to a high k or extinction coefficient and this technology has largely been abandoned.

A more promising $n < 1$ structure is hollow sapphire. It was first pointed out by Harrington and Gregory [11] that sapphire or Al_2O_3 has $n < 1$ from 10 to 16.7 μm and, in addition, it has a very small k value of 0.05 at 10.6 μm . This means that the theoretical loss predicted by MS theory is very low (less than 0.1 dB/m for a 1,000- μm -bore tube) for this material. Single-crystal sapphire tubing is fabricated by Saphikon, Inc. in Milford, NH in bore sizes ranging from 250 to 1070 μm . In Figure 13 we show the measured straight losses for five different bore sizes. Also in Figure 13, the theoretical losses for both the lowest order HE_{11} and next higher order HE_{12} modes are plotted. The measured losses are somewhat higher than that predicted by MS theory as a result of the roughness of sapphire's inner wall.

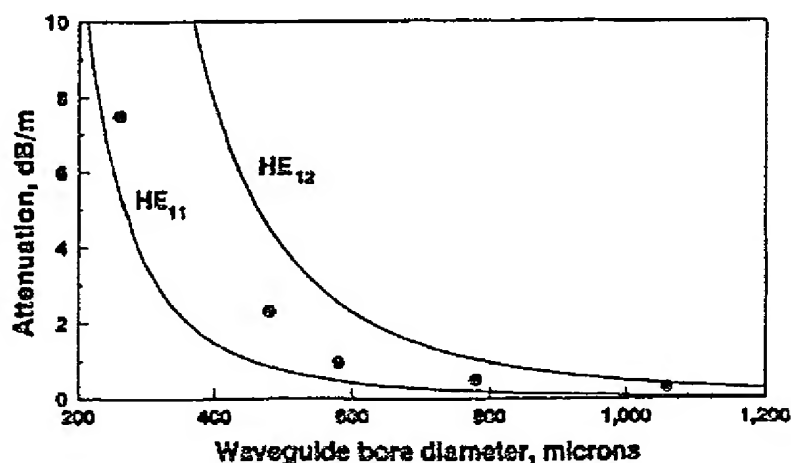


Figure 13. Measured loss at $10.6 \mu\text{m}$ for straight, hollow sapphire waveguides. Note that the losses fall between the theoretical losses for the HE_{11} and HE_{12} modes. This is a result of surface roughness of the sapphire tubing.

Gregory and Harrington [11] showed that the surface roughness of hollow sapphire accounted for the increased attenuation over that predicted by theory for the HE_{11} mode. Figure 14 shows the bending loss for the $530 \mu\text{m}$ bore tube. The curvature is not as great as it is for the HGWs, because sapphire has a high modulus and, therefore, cannot be bent to small diameters. Hollow sapphire delivery systems have been coupled to CO_2 lasers for use in gynecology and orthopedic surgery and they have been packaged in a water-cooled jacket for the delivery of over 1900 W of CO_2 laser power [34].

Laser Power Delivery in Hollow Glass Waveguides

Hollow waveguides are ideal for high power laser delivery because of the inherently high damage threshold of an air-core structure. For most medical and some low-power industrial applications, it is sufficient to be able to deliver up to 100 W of laser power. For these purposes, the HGWs can be used without cooling,

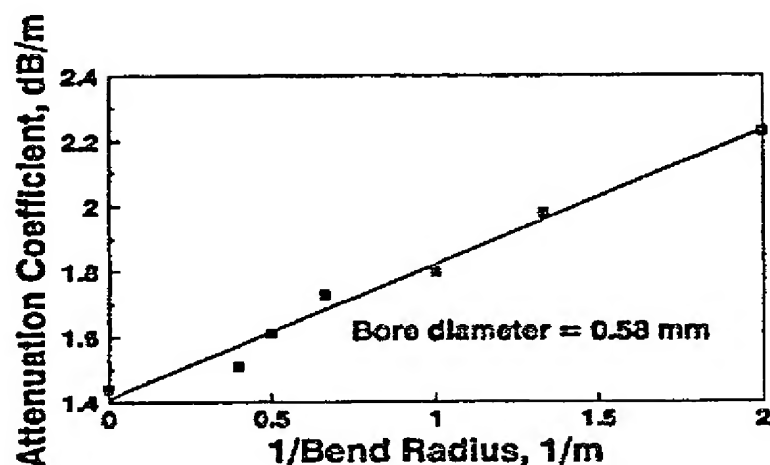


Figure 14. Bending loss for $560 \mu\text{m}$ bore, hollow sapphire tubing at $10.6 \mu\text{m}$.

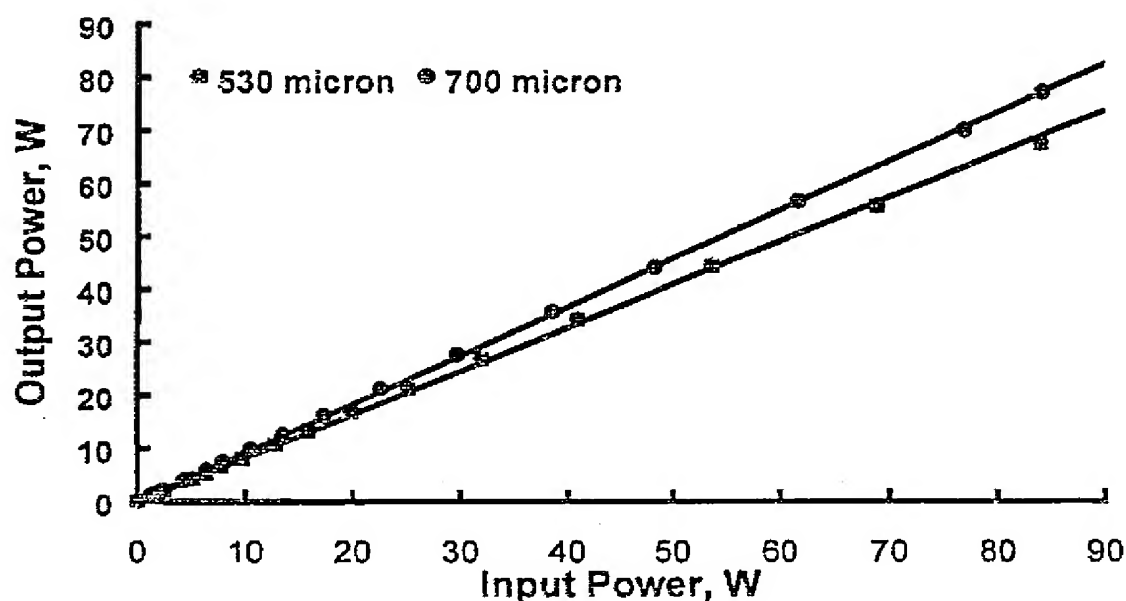


Figure 15. Low CO_2 laser power delivery for two bore size HGWs with no cooling. This is the typical laser power used in medical lasers and for low power cutting and marking.

although it is often helpful if an inert gas is used to purge the bore of the guide. In Figure 15, the results of low-power, CO_2 laser power delivery through 700 and 530 μm bore HGWs is shown. It is also possible to deliver higher CO_2 laser powers through the guides if active cooling is incorporated. In Figure 16, the results of high-power CO_2 laser transmission is shown through a 700- μm -bore HGW that

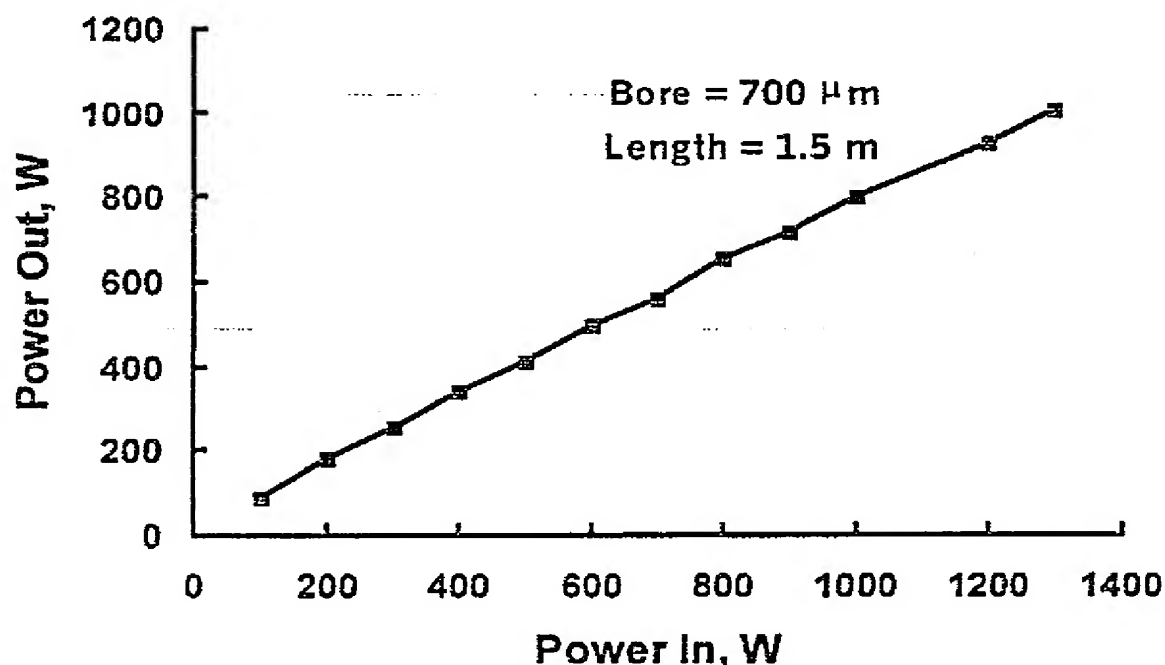


Figure 16. High CO_2 laser power delivery for a 700 μm bore HGW with a water cooling jacket. Note that the maximum power is just over 1,000 W!

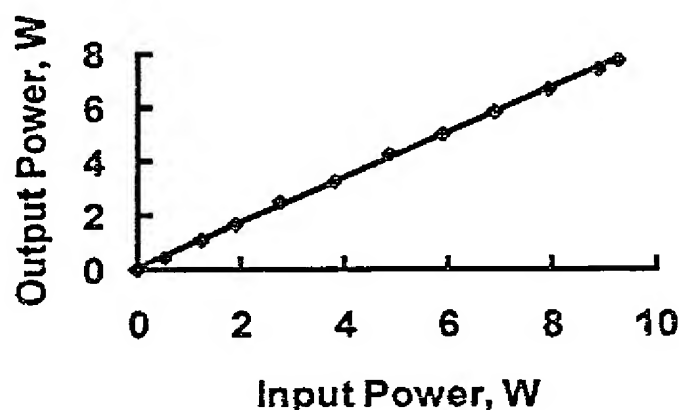


Figure 17. Power delivery through a HGW using an Er:YAG laser.

has a water jacket surrounding the guide. The maximum laser power delivered through the guide was just over 1000 W [34]!

The 2.94- μm , pulsed Er:YAG laser is becoming an important medical laser because the depth of ablation is very shallow and, therefore, this laser has great potential in surgical applications involving precise cutting and ablation. In Figure 17 the average 3 μm laser power delivered by the 1000 μm -bore HGW is shown. These data were obtained using a multimode Er:YAG laser made by Continuum. The maximum average output power of about 8 W represents a substantial average power for this wavelength. This power is sufficient for most surgical and dental applications.

The output beam profile of the HGW is important for many applications. In principle, the HGWs are nearly single-mode because the higher order modes are attenuated by the factor $(U_0)^2$ (see Eqn. 1). In practice, however, mode distortion can occur even with a $\text{TEM}_{(0)}$ input beam. The spatial profile can worsen on bending due to increased coupling into higher order modes. The amount of coupling into higher order modes is a function of the diameter of the waveguide, the roughness of the surface, and the refractive indices of the material. The spatial profile of a 530- μm -bore HGW is shown in Figures 18A and 18B. From the data one sees that it is possible to generate a single-mode HE_{11} output when the guide is straight or bent (Figure 18A), but, at other times when the guide is bent, low-order modes can be generated resulting in the modal pattern in Figure 18B. An important point is that the smaller the bore size the better the modal purity even on bending. A 250- μm bore, straight or bent guide, for example, retains a nearly perfect single mode output [35]. The near single-mode output from the glass waveguides is very important when small spot sizes are needed for precise cutting or marking.

Applications of Hollow Waveguides

Applications of hollow waveguides fall into two broad categories: laser power delivery and fiber sensors. As fiber sensors, hollow waveguides are generally used either to transmit black-body radiation for temperature measurements or as an active or passive link for chemical sensing. Saito and Kikuchi [36] give a good review of the use of hollow guides as IR fiber sensors. The use of hollow glass or

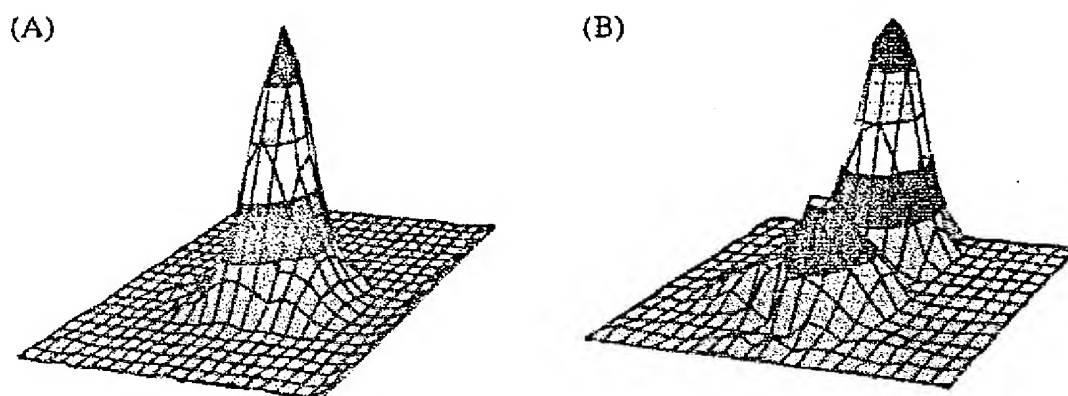


Figure 18. (A) High quality output beam profile from a bent 530 μm bore HGW and (B) A similar HGW, but with the output beam profile distorted somewhat by pushing against the side of the guide.

metallic waveguides to deliver laser power has largely been relegated to laser surgery where the required power is less than 100 W. Furthermore, most of the surgical applications to date involve the CO_2 laser, as this laser is one of the most commonly used medical lasers. As mentioned above, the HGWs are capable of delivering over a kilowatt of CO_2 laser power, yet they have not been accepted as flexible delivery systems for industrial lasers. There are two most likely explanations for this: 1. hollow guides have a somewhat higher loss when compared to current articulated arm technology; 2. industrial applications generally require a high quality (low M^2) laser output mode, whereas hollow waveguides can distort the TEM_{00} input beam of the CO_2 laser and this can lead to unacceptable kerfs and welds in cutting and welding applications.

Hollow waveguides are an ideal means of transmitting black-body radiation for thermometric measurements. In particular, the peak of black-body radiation near room temperature is around 10 μm where these guides transmit very well. They have also been used to transmit radiation above 1,000°C for the measurement of jet engine blade temperatures. As a delivery system in chemical sensing applications, hollow guides may be used merely as a passive fiber link from the chemical processing area to a remote detector, or they may play a more active role in which the guide is filled with the gas to be sensed. The latter application involves using the hollow guide itself as both the container for the gas sample and as a waveguide [26]. That is, a coiled hollow guide can replace a standard White cell to give a long pathlength and a small volume cell. Several researchers have used this method to measure small quantities of benign gases [26].

Conclusions

Hollow, IR transmitting waveguides are becoming an attractive fiber optic for the delivery of high power laser radiation as well as for important temperature and chemical fiber sensor applications. In general, these guides enjoy losses of a few tenths of a dB/m and are quite flexible. Because the energy is carried in the hollow core, there is no core material that might be easily damaged by high peak or average laser powers as there would be in a comparable solid-core IR fiber. In

addition, there is no Fresnel loss; this is especially important in comparison to the chalcogenide glass fibers, where Fresnel losses can be as high as 25%/surface. When selecting a hollow guide, it is important to remember that the optical principles are different between a hollow core guide and a solid-core fiber. The most important distinctions are that the hollow guides have a loss that varies as the reciprocal of the bore radius cubed and that there is an additional loss on bending which varies as $1/R$. These properties are not shared by conventional solid-core fibers. It is the authors' experience, however, that most applications of IR fibers do not require tight bending radii so that the additional loss on bending is not prohibitive. Two other advantages of hollow guides are important: these guides are nearly single mode and there seems to be no loss due to aging as no change in transmission has been observed after storage of the guides for over 2 years.

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Biography

James A. Harrington is Professor of Ceramic and Materials Engineering at Rutgers University in Piscataway, NJ. Dr. Harrington's basic research interests are in the area of specialty fiber optics and the optical properties of solids. Since 1977, he has worked on all aspects of infrared fibers, including fabrication, characterization, and applications. His current research interests include the development of IR fiber optics for use in the delivery of laser power in surgical and industrial applications and for use as chemical and thermal fiber sensors. Specifically, these fibers include hollow glass waveguides and single-crystal sapphire fibers.

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speed camera capable of recording up to 1000 frames per second. The resulting video is then analyzed to determine the radius r_0 of the bubble, the hole size R at which the ripples are first observed, and the number of ripples n^* . Because the hole expands very fast at first, R is much larger than R_c by the time the bubble begins to collapse. To compare the experiments with the theory, in which R enters as a parameter, we measured the latter at the onset of the instability for each given size of the bubble. The quantitative measurements are compared with the theoretical predictions for the dependence of n^* on the bubble size in Fig. 2. On a more qualitative level, the experiments show a suppression of the instability for small bubbles, in agreement with the threshold conditions above.

We conclude with a discussion of possible refinements of the theory and their relation to the geometric nature of the problem. A more complete theory would incorporate a (flattened) hemisphere as the initial condition, rather than a cone. Also, because of the progressive drainage of the liquid, the thickness l acquires a dependence on r (and time). This in turn implies nonuniform rigidity $K(r)$ and mass $\mu(r)$, leading to functions f , g , ϕ , ψ , and P of a more complicated form. On a more fundamental level, all these aspects should be addressed in terms of the coupled hydrodynamics of the slow viscous (liquid) flow and the rapid air flow (13). Yet the strong geometrical constraints involved in the problem are suggestive of the robustness of the results.

The question we have answered is akin to that of applying a curved surface onto a flat one in the most economical way, a problem that has taxed cartographers for many centuries and lies at the birth of differential geometry. It is also somewhat of an inverse counterpart to the problem of fitting a flat sheet to a three-dimensional landscape, which has been studied in various contexts (15–17) and is an issue that still vexes fashion designers. The relevance of the geometrical constraints is manifest, for example, in the strong dependence of the rippling on the size of the opening, which is closely related to a well-known theorem by Gauss (18), Jellett (19), and others, according to which (loosely put) a closed surface cannot be bent without being stretched, whereas an open surface can be bent inextensionally. Similarly, we find that a smaller hole implies a relatively stiffer bubble and hampers the rippling. Although the precise forms of the functions ϕ and ψ arise from the physical constraints and dynamics imposed by the forces and various boundary conditions, the essence is in the geometry.

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Control of Thickness and Orientation of Solution-Grown Silicon Nanowires

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Bulk quantities of defect-free silicon (Si) nanowires with nearly uniform diameters ranging from 40 to 50 angstroms were grown to a length of several micrometers with a supercritical fluid solution-phase approach. Alkanethiol-coated gold nanocrystals (25 angstroms in diameter) were used as uniform seeds to direct one-dimensional Si crystallization in a solvent heated and pressurized above its critical point. The orientation of the Si nanowires produced with this method could be controlled with reaction pressure. Visible photoluminescence due to quantum confinement effects was observed, as were discrete optical transitions in the ultraviolet-visible absorbance spectra.

One-dimensional quantum wires are expected to play a vital role as both interconnects and functional components in future mesoscopic electronic and optical devices and also to provide an opportunity to test fundamental quantum mechanical concepts (1, 2). As the wire diameter approaches the carrier de Broglie wavelength, quantum confinement effects shift band gap energies and, in Si, induce visible photoluminescence (3). The electronic and optical prop-

erties of the nanowires strongly depend on size; therefore, size control and tunability are key to the success of any method of synthesizing quantum wires. Dimensionality also affects the material properties of nanowires. The absence of translational symmetry in Si could profoundly affect the electronic properties: Bulk Si is an indirect semiconductor with a band gap of 1.1 eV, whereas linear polysilane chains exhibit a 3.89-eV direct gap (4). The lattice orientation in a wire can provide a tuning parameter, unavailable in quantum dots, to adjust material properties to suit particular applications. In carbon nanotubes, for example, the bonding geometry and orientation profoundly affect the electronic structure and can lead to either metallic or insulating behavior (5). Calculations for Si nanowires

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have shown that bond orientational order will influence their physical properties as well (6). These effects, however, have not been experimentally observed in Si nanowires, largely because of inability to obtain highly crystalline samples with narrow wire diameter distributions and tunable crystallographic orientation.

The high-aspect-ratio quantum wires (those with a high ratio of length to diameter) must have diameters less than 10 nm. Self-assembly strategies are best suited for directing materials synthesis on this mesoscopic length scale. Many successful strategies have been developed for semiconductor quantum dots (7). These methods rely on the ability to isolate nanometer-scale reaction compartments, control crystallization kinetics, provide a growth template, or some combination of these. Strategies have also been developed for semiconductor nanowire self-assembly in solution and in the gas phase; however, greater control over the wire diameter distributions and crystallinity is desired (1, 2, 8). Certainly, the production of single-walled carbon nanotubes in the gas phase demonstrates that self-assembly methods can yield very high-quality one-dimensional (1D) materials (2). Because solution-phase synthetic methods can enable great chemical flexibility and synthetic tunability, as has been demonstrated for quantum dots (7), it is desirable to extend the highly controlled solution-phase self-assembly methods to semiconductor nanowires. The work presented here demonstrates the ability to self-assemble Si nanowires in solution with a very high degree of control, producing wires with highly crystalline cores, narrow diameter size distributions, aspect ratios greater than 1000, and tunable crystallographic orientation.

We used solvent-dispersed, size-monodisperse, alkanethiol-capped gold (Au) nanocrystals to direct Si nanowire growth with narrow wire diameter distributions. Sterically stabilized Au nanocrystals were dispersed in supercritical hexane with a silicon precursor, diphenylsilane, at temperatures of 500°C and 270 bar (or 200 bar in some cases). At these temperatures, the diphenylsilane decomposes to Si atoms. The phase diagram for Si and Au indicates that at temperatures above 363°C, Si and Au form an alloy in equilibrium with pure solid Si when the Si concentration with respect to Au is greater than 18.6% (2). Under the reaction conditions used, the Si atoms most likely dissolve into the sterically stabilized Au nanocrystals until reaching supersaturation, at which point they are expelled from the particle as a thin nanometer-scale wire. Figure 1 shows a schematic of the nanocrystal-directed

nanowire self-assembly process. This fluid medium provides the high temperatures necessary to promote Si crystallization. More than 30 years ago, this phase behavior was exploited to grow micrometer-diameter Si wires on a substrate by chemical vapor deposition (9). Au droplets placed on a Si substrate directed Si "whisker" formation in the vertical direction by a process called vapor-liquid-solid (VLS) growth (10). Recently, VLS growth was applied to the gas phase to form nanometer-scale Si wires (11). Gas-phase liquid Au droplets were produced by laser ablation in the presence of silane at high temperatures. Si wires spontaneously formed. Although this method produces relatively large quantities of nanometer-scale Si wires, the liquid droplets naturally have a broad size distribution, and therefore, the nanowires also have relatively broad size distributions (10). In the supercritical (sc) fluid environment, relatively size-monodisperse Au nanocrystals can be maintained to seed nanowire growth at the high temperatures necessary to achieve wire growth.

Fig. 1. Schematic of the proposed nanowire growth process. Thermal degradation of diphenylsilane results in free Si atoms that dissolve in the Au nanocrystal until reaching a SiAu alloy supersaturation, when Si is expelled from the nanocrystal as a crystalline nanowire. This wire is depicted with a preferred <111> orientation.

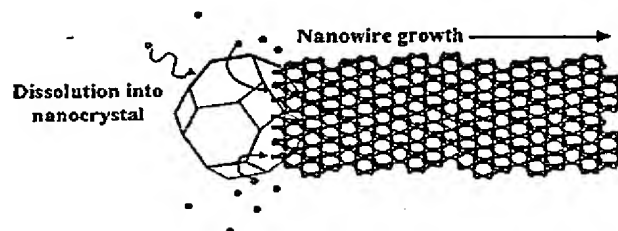


Fig. 2. TEM images of Si nanowires synthesized at 500°C in hexane at pressures of 200 bar (A and B) and 270 bar (C and D). The nanowires are highly crystalline. In both (B) and (D), the lattice fringes are the Si (111) planes, separated by 3.14 Å. The nanowire samples were prepared for imaging by dispersing them in chloroform and evaporating a drop onto a carbon-coated copper TEM grid. A JEOL 2010 transmission electron microscope with 1.7 Å point-to-point resolution operating with a 200-kV accelerating voltage with a GATAN digital photography system was used for TEM. In situ EDS (Oxford Instruments) confirmed that the samples were high in Si content with residual C and O.

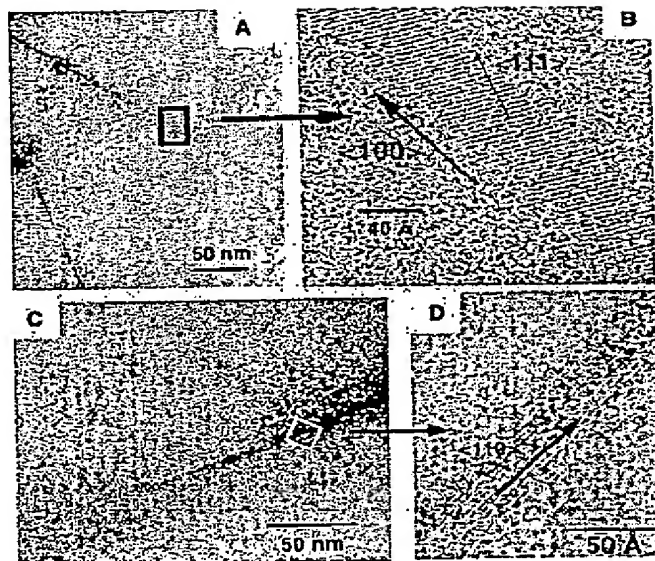


Figure 2 shows typical transmission electron microscopy (TEM) images of the wires formed with this process. The 40 to 50 Å diameter nanowires are 1D "single crystals" with extremely narrow diameter size distributions (less than $\pm 10\%$ SD about the mean wire diameter) and aspect ratios greater than 1000. These nanowires photoluminesce in the blue because of quantum confinement effects (Fig. 3), and the optical properties were found to strongly depend on the wire crystallographic orientation measured by TEM and electron diffraction: <100> oriented wires exhibit higher exciton energies than the <110> oriented wires with a diameter of 40 to 50 Å.

The Si nanowires were prepared by thermally degrading diphenylsilane in sc-hexane [supercritical temperature and pressure ($T_c = 235^\circ\text{C}$; $P_c = 30$ bar)] at 500°C and 200 or 270 bar in an inconnell high-pressure cell (12). The nanowire synthesis relies on dodecanethiol-capped Au nanocrystals, ranging between 20 and 30 Å in diameter, formed using standard arrested precip-

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itation procedures to direct wire formation (13). A light brown precipitate consisting of Si nanowires formed during the course of the reaction (TEM images in Fig. 2). No color change was observed in the absence of diphenylsilane. In the absence of Au nanocrystals, the solution quickly turned from orange to brown and then to clear as polydisperse black micrometer-sized Si particles formed and settled on the walls of the reaction vessel. Chloroform was used to extract the nanowires from the cell upon cooling and depressurization. The solvent provides a high diffusion coefficient, on the order of 10^{-3} to 10^{-4} cm² s⁻¹ (14), which is necessary for rapid reactant diffusion to help avoid Si nanocrystal growth in the bulk solvent.

The Si nanowires were characterized with a variety of techniques. Energy dispersive x-ray spectroscopy (EDS) and x-ray photoelectron spectroscopy (XPS) revealed that the nanowires consisted of a high con-

centration of Si, with residual amounts of C, O, and Au (15). As shown in Fig. 2, TEM revealed Si (111) lattice planes with 3.14 Å spacing in all of the samples, confirming that the nanowire cores did indeed consist of pure crystalline Si. The wires exhibited diameters approximately 10 Å larger than the nanocrystal diameter, and the thickness appeared to fluctuate along the length of the wire by ± 1 lattice plane; however, these variations may result from the resolution limitations of the microscope itself. The wires formed at 200 bar exhibited a preponderance of $\langle 100 \rangle$ oriented wires, whereas the samples synthesized at 270 bar contained wires oriented almost exclusively along the $\langle 110 \rangle$ direction. Pressure appears to be the determining factor in tuning the lattice orientation in the wire: At higher pressures, the brown color appeared sooner than at lower pressures, indicating that diphenylsilane degrades faster at higher pressures, and therefore it appears that the nanowire growth rate affects the nanowire orientation. A semitransparent coating of oxide or hydrocarbon on the nanowires was apparent in the TEM images of all of the wires (15). It is unclear whether this coating forms after removal from the reaction vessel or if it forms on the wire surface during nanowire growth. However, the thickness of this coating increased over time if the wires were exposed to air.

The absorption edge of the Si nanowires was strongly blue-shifted from the bulk indirect band gap of 1.1 eV and showed sharp discrete absorbance features and relatively strong "band edge" photoluminescence (PL). These optical properties likely result from quantum confinement effects, although we cannot rule out the possibility of additional surface states as well (2, 16). The $\langle 110 \rangle$ oriented nanowires exhibited distinctly molecular-type transitions, whereas the $\langle 100 \rangle$ oriented wires exhibited a strong feature reminiscent of the L \rightarrow L critical point in the Si band structure with a slowly rising phonon-assisted optical transition. The $\langle 100 \rangle$ oriented wires exhibited a significantly higher exciton energy than the $\langle 110 \rangle$ oriented wires, as predicted by Yorikawa and co-workers (6).

Thus, the tunability of the lattice orientation in the silicon nanowires leads to different optical properties that could be exploited in applications. The $\langle 100 \rangle$ nanowires exhibited a single sharp PL feature at 3.75 eV with a weak shoulder shifted to lower energy at 1.9 eV. The weak low-energy peak likely results from the surrounding oxide (16). The $\langle 110 \rangle$ oriented wires exhibited three distinct peaks at 3.35, 2.9, and 2.55 eV. The relatively robust PL exhibited initially after synthesis became completely quenched over the course of

about 1 week, but the PL energy did not appear to shift to lower energy upon exposure to air. Certainly, the lattice orientation is expected to affect electron transport in future mesoscopic interconnects employing Si and must be considered in the design and manufacture of such devices.

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- Under a nitrogen atmosphere, the nanocrystals were dispersed in diphenylsilane with a Au:Si mole ratio of 0.1%, then loaded into an Inconnel high-pressure cell (0.2 ml) and sealed under a nitrogen atmosphere. The cell was attached, via a three-way valve, to a stainless steel high-pressure tube (~ 40 cm³) equipped with a stainless steel piston. A high-pressure liquid chromatography pump (LDC Analytical) was used to pump deionized water into the back of the piston and displace oxygen-free anhydrous hexane through an inlet heat exchanger and into the reaction cell to the desired pressure of either 200 or 270 bar. The cell was covered with heating tape (0.6 m) and heated to 500°C (± 0.2 °C) using a platinum resistance thermometer and a temperature controller. The reaction proceeded at these conditions for 1 hour.
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- XPS revealed that the samples consisted of less than 0.1% Au, with ratios of Si:O of 2:1 and Si:C of 1:3. The relatively high concentrations of carbon and oxygen stem from the fact that the volume of the ~ 40 Å thick coating of carbon and oxygen surrounding the 40 Å diameter wire has a volume three times that of the silicon core.
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- K.P.J. thanks the U.S. Department of Energy and NSF for support. B.A.K. thanks DuPont for support through a Young Professor Grant.

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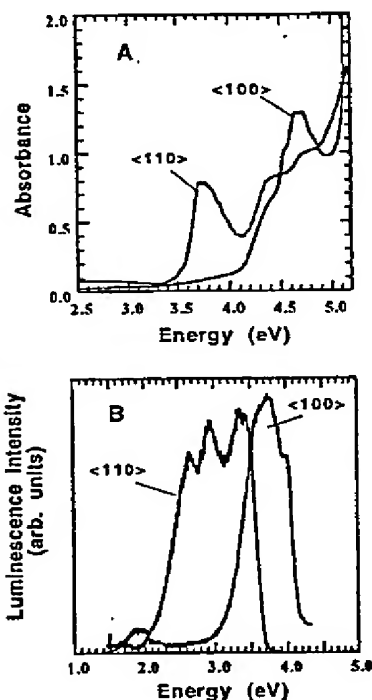


Fig. 3. Room-temperature optical properties of the Si nanowires imaged in Fig. 2. (A) Absorbance spectra (from a Varian Cary 500 UV-Vis-NIR spectrophotometer) for a dilute nanowire suspension in hexane. The spectra labeled $\langle 100 \rangle$ are for nanowires formed at 200 bar (Fig. 2, A and B) and $\langle 110 \rangle$ at 270 bar (Fig. 2, C and D). (B) Photoluminescence spectra (from a SPEX Fluorolog-3 spectrophotometer) of nanowires dispersed in hexane with excitation energies of 4.46 eV (277 nm, $\langle 100 \rangle$) and 4.12 eV (300 nm, $\langle 110 \rangle$).

The Formation of Dimensionally Ordered Silicon Nanowires within Mesoporous Silica

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Nanoscale one-dimensional (1D) structures of semiconductor nanowires are expected to play a vital role as materials for both interconnects and emerging future technologies because of their low-dimensionality and unique optical, electrical, and mechanical properties.¹ Even though the preparation of semiconductor nanowires in bulk quantities is now possible, unanswered questions relating to their processibility remain: namely, how can they be evenly distributed across a surface, or within a 3D structure, and how can they be made sufficiently robust to be used practically within planar or multidimensional devices? Consequently, a limiting step toward the development of nanotechnology is the configuration of nanowires into useful electronic device architectures.² Recently, mesoporous solids³ that contain uni-directional arrays of pores, typically 2–15 nm in diameter, running throughout the material have been exploited as templates for semiconductor materials formed from the gas phase.^{4–6} Certainly, these gas-phase methods have yielded high-quality semiconductor nanomaterials but very often high temperatures or extensive reaction times are required for successful nucleation and growth of the materials within the mesopores making these techniques often costly and time-consuming.^{4,6} In this communication we describe the use of a novel supercritical fluid solution-phase technique to rapidly fill the pores of mesoporous silica with silicon nanowires. The silica mesoporous matrix provides a means of producing a high density of stable, well-ordered arrays of semiconductor nanowires.

Hexagonal mesoporous silica was prepared by hydrolyzing tetramethoxysilane (TMOS) in the presence of a poly(ethylene oxide) (PEO)–poly(propylene oxide) (PPO) triblock copolymer surfactant (PEO₂₆PPO₃₉PEO₂₆) and HCl (0.5 M).⁷ Silicon nanowires were formed within the pores of the mesoporous silica by degrading diphenylsilane (0.022 mol in hexane) in a high-pressure

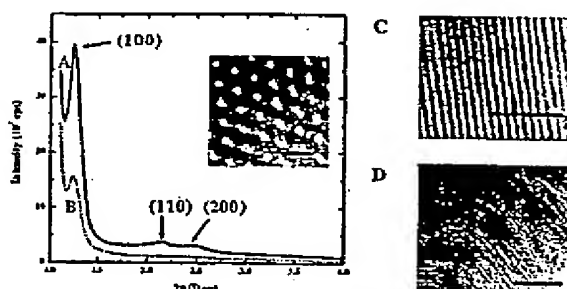


Figure 1. PXRD patterns of (a) as-synthesized calcined mesoporous silica prepared by using the PEO₂₆PPO₃₉PEO₂₆ triblock copolymer surfactant and (b) the same hexagonal mesoporous sample after silicon nanowire inclusion in the mesopores. The insert shows the ordered arrays of pores in the mesoporous matrix (scale bar = 12 nm). "Side-on" TEM image of (c) the as-synthesized calcined mesoporous silica (scale bar = 50 nm) and (d) after incorporation of silicon through diphenylsilane decomposition (scale bar = 50 nm).

reaction cell at 500 °C and a pressure of 375 bar for 15 min.⁸ The white mesoporous materials changed color from white to yellow to dark orange/red during the course of the reaction. The end product was homogeneous in appearance. No color change was observed in the absence of diphenylsilane. After the reaction had finished, the contents of the cell were washed with hexane and the relatively large (~1 mm dimensions) particles of the dark orange/red mesoporous silica incorporating silicon were manually extracted and dried for analysis.⁹

Figure 1a shows the powder X-ray diffraction (PXRD) pattern of the calcined mesoporous silica used in the present work. Three well-resolved peaks can be readily indexed to (100), (110), and (200) reflections for a hexagonal mesoporous solid.^{3,10} The position of the intense (100) peak reflects a *d* spacing of 7.1 nm corresponding to a pore-to-pore distance of 8.2 nm. Complementary to the PXRD data, transmission electron microscopy (TEM) revealed that the calcined mesoporous silica had a well-ordered mesoscopic structure with a wall thickness and pore diameter of approximately 3 and 5 nm, respectively. A pore-to-pore distance of 8.3 nm was obtained by low-angle PXRD after inclusion of silicon in the mesoporous silica (Figure 1b). The reduced intensity

(8) The cell was attached, via a three-way valve, to a stainless steel high-pressure tube (~21 mL) equipped with a stainless steel piston. An Isco high-pressure pump (Isco Instruments, PA) was used to pump CO₂ into the back of the piston and displace oxygen-free anhydrous hexane into the reaction cell to 375 bar.¹⁶ The cell was placed in a furnace and heated to 500 °C (±1 °C) using a platinum resistance thermometer and temperature controller. The reaction proceeded at these conditions for 15 min. The high pressures and temperatures used in these experiments and the volatile nature of the chemicals could potentially lead to fire or explosion. Suitable safety precautions should be taken into consideration including the use of a blast screen.

(9) A JEOL 1200 EX electron microscope operating with an 80 kV accelerating voltage was used for transmission electron microscopy (TEM). Samples were redispersed in chloroform and a drop of the mixture was placed on a carbon-coated copper TEM grid. Powder X-ray diffraction (PXRD) profiles were recorded on a Philips 3710 PWD diffractometer, equipped with a Cu K α radiation source and standard scintillation detector. Silicon magic angle spinning nuclear magnetic resonance (²⁹Si MAS NMR) spectra were obtained at room temperature using a Chemagnetics CMX Lite 300 MHz apparatus. Pulses of 30° with 4 s pulse width were used. Pulse delay times were varied in the range 60 to 10 000 s depending on the relaxation process. Samples were spun at 5 kHz. Chemical shifts are quoted relative to tetramethylsilane and referenced using tetramethoxysilane. The UV–visible spectra of the silica samples suspended in ethanol were recorded on a Hewlett-Packard HP 8543 diode array spectrophotometer. The surface areas of the samples were measured using nitrogen BET isotherms at 77 K on a Micromeritics ASAP 2010 volumetric analyser (Norcross, GA). Before the adsorption data were taken the samples were degassed for 12 h at 120 °C.

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of the (100) peak and the absence of (110) and (200) peaks noted after silylation probably arises from the inclusion of silicon in the mesopores, possibly resulting in an increased residual strain on the silica walls.¹¹

The TEM images shown in parts c and d of Figure 1 are evidence for inclusion of silicon in the mesopores. Figure 1c displays a "side-on" view of the mesoporous matrix showing good contrast between the silica walls and the empty pores. A side-on view of mesoporous silica incorporating crystalline silicon shows very little contrast between the silica walls and the silicon. This is expected since the work functions of SiO₂ and Si are similar. A comparison of these two TEM images (Figure 1c,d) is consistent with silicon filling the mesopores. We believe that such a comprehensive filling of the pores with semiconductor nanowires has not previously been observed using other inclusion methods. Moreover, the comparatively short reaction time of the present method makes it noteworthy. At diphenylsilane concentrations above 0.002 mol "whiskers" of silicon wires several microns in length (not shown) were observed to extrude from the mesoporous surface. Hence, the mesopores appear to act as directional templates for surface nanowire growth once the pores have been filled. The purity and crystallinity of the mesoporous silicon nanowires was confirmed by PXRD at high angles. The relatively sharp peaks in the PXRD pattern can be indexed to a diamond structure of silicon with a lattice constant $a = 0.357$ nm, which is in good agreement with literature values for silicon.¹²

Silicon magic angle spinning nuclear magnetic resonance (²⁹Si MAS NMR) was undertaken on a mesoporous silica sample before and after silicon inclusion. As synthesized, the mesoporous silica displays two distinct ²⁹Si chemical shifts at -103.4 and -111.4 ppm respectively assigned to species Q₃(SiO₃(OH)) and Q₄(SiO₄).^{6,13} A smaller feature apparent at -92.4 ppm was assigned to Q₂(SiO₂(OH)₂).⁶ The Q₃ and Q₂ peaks result predominantly from species at the surface of the silica walls of each pore whereas Q₄ species are within the bulk silica walls. The Q₄ to (Q₃ + Q₂) peak area ratio of 0.5 is consistent with a wall thickness that is approximately 0.33–0.5 times the size of the pore diameter¹³ and is complementary to the results obtained by PXRD and TEM. New peaks centered around -80 ppm and a complete loss in the intensity of the Q₃ peak at -103.4 ppm were observed upon inclusion of silicon in the pores. Curve fitting of the peak at -80 ppm resolved two features, a relatively sharp peak at approximately -80.8 ppm and a broader less distinct peak at -88.0 ppm. The peak at -80.8 ppm is most probably due to crystalline silicon (Si₄Si) as, in a control experiment, silicon from a powdered Si(111) wafer produced a single sharp peak at -82.5 ppm. The small downfield shift of -1.7 ppm (relative to the powdered wafer) most likely results from the lattice expansion of the SiO₂ walls after silicon inclusion resulting in lower electron densities. We have assigned the peak at -88.0 ppm to silicon atoms attached at the surface of the intrachannel silicon (Si₃(wall)-O-Si(wire)). The loss of surface Si-OH site resonances at -103.4 (Q₃(SiO₃(OH))) and -92.4 ppm (SiO₂(OH)₂) and the formation of the Si₃(wall)-O-Si(wire) peak probably infers that the silicon nanowires are anchored to the surface of the mesoporous walls and pore-filling has been achieved. ²⁹Si MAS NMR also revealed that the volume ratio of silicon to silica in the mesoporous sample was ~29%, which correlates to ~80% of the mesopores being filled with silicon.

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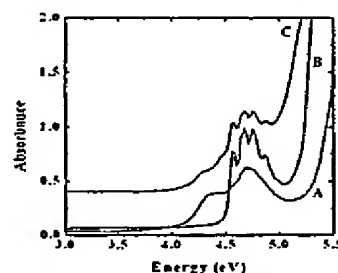


Figure 2. UV-visible absorbance spectrum of (a) as-synthesized calcined mesoporous silica mesoporous, (b) silicon nanowires grown from gold nanocrystals, and (c) silica loaded with crystalline silicon nanowires.

Nitrogen absorption-desorption isotherms obtained for a calcined mesoporous silica sample exhibited a high surface area of ~1045 m² g⁻¹. Upon inclusion of the silicon nanowires in the mesopores the surface area decreased to ~72 m² g⁻¹, which supports the hypothesis that the pores have been filled.

Figure 2 shows the UV-visible absorption of as-synthesized calcined mesoporous silica before (Figure 3a) and after incorporation of crystalline silicon (Figure 3c) and also the electronic spectrum of micron length silicon nanowires (Figure 3b), ~4 nm in diameter, grown using colloidal gold nanoparticles as nucleation seeds.¹⁴ The mesoporous silicon nanowires possess similar optical properties to the silicon nanowires seeded from the gold nanocrystals. Both samples show a strongly blue-shifted absorption band-edge from the bulk indirect gap of 1.1 eV and display sharp, discrete absorbance features. It is likely that these optical properties result from quantum confinement effects.¹⁵ Both the mesoporous silicon and Au-grown silicon nanowire samples exhibit strong absorption features centered at 4.7 eV indicative of the L → L transition previously reported for (100) orientated silicon nanowires.¹⁴ Additional peaks in the mesoporous silicon sample that are not observed in the Au-grown silicon nanowires are centered at 4.25 eV probably arise from absorption due to the silica mesoporous solid. The UV-visible spectra confirm that the mesoporous silicon nanowires possibly exhibit quantum confinement effects previously observed in colloidal grown nanowires.

In conclusion, a novel supercritical fluid solution phase technique has been utilized to fill 5 nm diameter pores of hexagonal mesoporous silica with quantum-confined silicon nanowires. This is the first example of utilizing supercritical fluids as an inclusion medium for mesoporous materials. The high diffusivity¹⁶ of the supercritical fluid phase, of the order of 10⁻³–10⁻⁴ cm² s⁻¹, enables rapid diffusion of reactant precursor into the pores of the silica matrix where swift nucleation and growth can occur resulting in a reaction rate several orders of magnitude faster than are obtainable using traditional gas-phase or solution-phase deposition methods.

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(14) Gold-grown nanowires were prepared using a modification of the method published by Holmes et al. (*Science*, 2000, 287, 1471). Under a nitrogen atmosphere, dodecanethiol-capped gold nanocrystals were dispersed in diphenylsilane with a Au:Si ratio of 0.1%, then loaded into the high-pressure cell (5 mL) and sealed under a nitrogen atmosphere. The cell was heated and pressurized as previously described.¹

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Template-directed vapor-liquid-solid growth of silicon nanowires

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The synthesis of Si nanowires in nanoporous anodic alumina membranes was demonstrated using a combination of Au electrodeposition and vapor-liquid-solid growth at 500 °C using SiH₄ as the Si source. The average diameter of the nanowires was 200 ± 54 nm which was close to the pore size distribution of the membranes. High-resolution transmission electron microscopy revealed that the nanowires consist of a crystalline Si core, oriented in the (100) or (211) growth direction, with a thin (<3 nm) native oxide coating. In this process, Au terminates both ends of the growing wires, resulting in the formation of Au-Si-Au nanowires. © 2002 American Vacuum Society. [DOI: 10.1116/1.1430240]

I. INTRODUCTION

There is currently intense interest in one-dimensional nanostructures, such as nanotubes and nanowires, due to their potential to test fundamental concepts of dimensionality and to serve as building blocks for nanoscale devices.¹ Carbon nanotubes have been the focus of much of this work, however, it has proven difficult to control conductivity in these structures^{2,3} which restricts their application. As a result, there is renewed interest in the synthesis and properties of semiconductor nanowires, such as Si, Ge, and GaAs, and the development of assembly methods to fabricate functional nanoscale electronic^{4,5} and optical devices⁶ using these structures.

The electrical and optical properties of semiconductor nanowires strongly depend on the wire diameter⁷ as well as the crystallographic orientation⁸ and defect structure of the material. Synthesis techniques have therefore been directed at producing large quantities of single crystal nanowires with uniform diameters and controlled growth orientations. In one approach, nanoporous materials, such as anodized aluminum, have been used as templates for nanotube and nanowire synthesis via electrodeposition^{9,10} or chemical vapor deposition.^{11,12} Template-directed synthesis affords good control over nanowire diameter and uniformity and provides a support structure for the fabrication of multi-layered nanowires.¹³ Direct deposition into nanopores, however, typically yields polycrystalline material. Alternatively, vapor-liquid-solid (VLS) growth has been used to produce single crystal semiconductor nanowires such as Si¹⁴ and GaAs.¹⁵ In the VLS growth of Si nanowires, a metal, such as gold (Au) is used as a catalyst agent to nucleate whisker

growth from a Si-containing vapor.¹⁶ Au and Si form a liquid alloy that has a eutectic temperature of 363 °C, which, upon supersaturation, nucleates the growth of a Si wire. The fabrication of nanometer-size wires can be achieved through careful control of the metal particle size and the growth conditions.¹⁷ VLS growth of Si nanowires has been demonstrated using techniques in which the metal particles were either fabricated by metal deposition and photolithography on a Si surface,¹⁸ produced in the gas phase via laser ablation^{19,20} or evaporation,²¹ or were in the form of nanocrystals suspended in a supercritical solution.²² These methods yield crystalline Si nanowires and afford some degree of control over the wire diameter. However, given the lack of a support structure for nanowire fabrication, it is difficult to produce multi-layered nanowires using these techniques.

This report demonstrates the synthesis of Si nanowires by VLS growth in nanoporous anodic alumina membranes. By combining template-directed synthesis with VLS growth, this technique provides a simple method to produce large quantities of crystalline Si nanowires with uniform diameters that are determined by the pore size of the membrane. Furthermore, by combining metal electrodeposition with VLS growth in nanoporous membranes, this technique enables the fabrication of metal-Si-metal nanowires.

II. EXPERIMENT

Commercially available anodic alumina membranes (Whatman Scientific) with a nominal pore diameter of 200 nm and thickness of 60 μm were used as templates in this initial study. Alumina membranes with pore diameters ranging from 4 nm to greater than 200 nm can be produced through the anodization of aluminum in various acids as described by Routkevitch *et al.*²³ The Au catalyst for VLS

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growth was electrodeposited into the membranes using a process described by Martin *et al.*²⁴ In order to control the placement of the Au catalyst relative to the top surface of the pore, a 10–30 μm long segment of silver (Ag) was initially deposited in the pores, followed by a thin (0.2–1.2 μm) segment of Au. The Ag was then removed by etching in 8.0 M HNO_3 , leaving only the thin Au segment near the center of the membrane. The membrane was cleaned with distilled water in an ultrasonic bath and dried under ambient conditions. Vapor-liquid-solid growth was carried out at 500 °C in a low pressure, isothermal reactor using a 5% mixture of SiH_4 in H_2 .

Plan view and cross sectional images of the membranes after VLS growth were analyzed using a 30 kV Philips XL20 scanning electron microscope (SEM). A transmission electron microscope (TEM) sample was prepared by scraping a portion of the nanowires off the top surface of the membrane onto a TEM grid. High-resolution transmission electron microscope (HREM) images and electron diffraction patterns of the Si nanowires were recorded at 200 kV using a Hitachi HF2000 field emission transmission electron microscope.

III. RESULTS AND DISCUSSION

In traditional VLS growth, the metal particle is either supported on a surface or produced in the gas phase and is readily accessible to the vapor phase growth species. In our application, the gold catalyst was buried deep within the pore. Consequently, a careful choice of reaction conditions was required to ensure that the vapor phase species diffused far into the pore and preferentially reacted with the metal catalyst rather than the internal walls of the pore or the top surface of the membrane. A model of gas phase diffusion and reaction in a cylindrical pore²⁵ was used to predict the SiH_4 concentration profile within a 200 nm diameter pore as a function of temperature and pressure in order to define initial conditions for VLS growth. At a pressure of 10 Torr, SiH_4 was predicted to diffuse the entire pore length (60 μm) with minimal Si deposition for temperatures less than 600 °C. In order to extend this technique to 10 nm diameter pores, the growth temperature must be reduced to less than 400 °C to prevent Si deposition on the internal walls of the pore. Initial growth experiments, carried out at 500 °C, 13 Torr total pressure and a SiH_4 partial pressure of 0.65 Torr, on membranes that did not contain Au resulted in negligible Si deposition on the membrane surfaces for growth times up to 45 min.

Growth experiments were then carried out under identical conditions (500 °C, 13 Torr, 0.65 Torr SiH_4) using the Au-impregnated nanoporous anodic alumina membranes. Gold catalyzes the decomposition of SiH_4 under these conditions resulting in the rapid growth of Si nanowires within the pores. Nanowires began to emerge from the top of the pores after approximately 15 min [Fig. 1(a)] and populated the surface after 20 min of growth [Fig. 1(b)]. Si nanowires were observed growing in nominally every pore of the membrane. The presence of Au at the tip of the nanowires, as confirmed

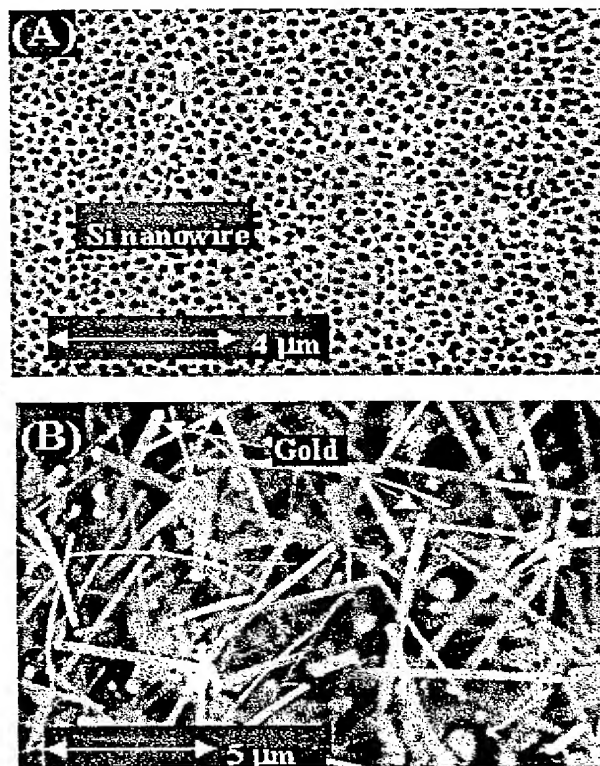


FIG. 1. Scanning electron microscopy images of the top surface of Au-impregnated alumina membranes during various stages of VLS growth from the nanopores. After 15 min (A) Si nanowires began to emerge from the pores and fully covered the surface after 20 min (B). Gold balls present at the tips of the nanowires in (B) are evidence of a VLS growth mechanism.

by energy dispersive x-ray spectroscopy, provided evidence of a VLS growth mechanism.

During growth, the Si nanowire fills the entire internal diameter of the pore and pushes a portion of the Au–Si alloy toward the pore surface as the wire grows in length. As a result, the final diameter of the Si nanowire is determined exclusively by the initial diameter of the pore. The pore diameter of the anodic alumina membranes ranged from 110 nm to 270 nm (sample of 280 pores) with an average diameter of 194 ± 33 nm. Correspondingly, the diameter of the Si nanowires ranged from 100 nm to 340 nm (sample of 50 nanowires) with an average diameter of 200 ± 54 nm.

A low magnification TEM image of a sample of nanowires removed from the membrane surface is shown in Fig. 2. The long wire near the center of the image is aligned with the beam parallel to the $[01\bar{1}]$ zone axis. The growth direction of the nanowire is $[100]$. Crystallographic analysis of the TEM images revealed Si nanowires characterized by two distinct growth axes: $\langle 100 \rangle$ and $\langle 211 \rangle$. The $\langle 211 \rangle$ orientation has previously been observed in Si nanowires fabricated by VLS growth from a surface²⁶ or via laser ablation,^{19,20} while $\langle 100 \rangle$ -oriented wires were obtained in supercritical fluid growth.²² HREM analysis of the Si nanowires, shown in Fig. 3, revealed a nearly defect-free Si core with a 2–3 nm thick amorphous native oxide coating. All of the $\langle 100 \rangle$ -oriented

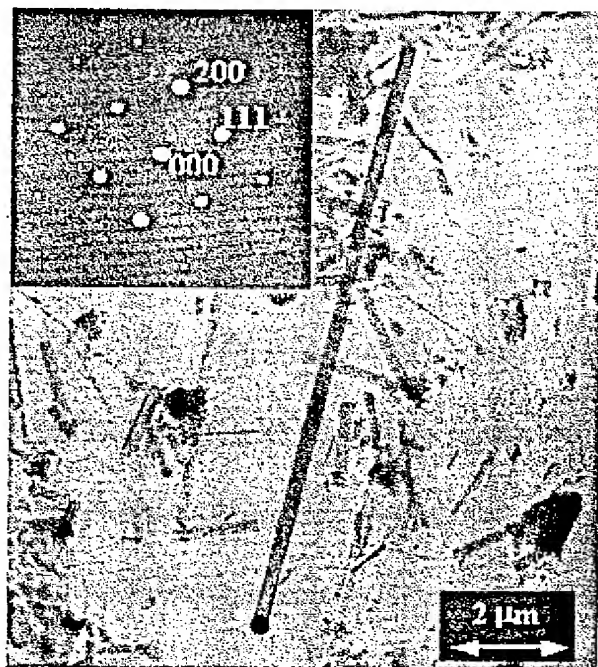


FIG. 2. Low magnification transmission electron microscopy (TEM) image of nanowires removed from the top surface of the membrane and supported on a grid. The insert is the parallel-beam diffraction pattern along the $[01\bar{1}]$ zone axis of the Si nanowire shown at the center of the image, which has a $[100]$ growth direction.

wires observed thus far by TEM appear to be single crystals, based on diffraction contrast imaging, high-resolution imaging, and electron diffraction.

The nanowires were removed from the anodic alumina membranes by wet etching in 5.0 M NaOH for approximately two hours at room temperature. Figure 4 is a SEM photo of a Au-Si-Au nanowire produced by template-directed VLS growth after removal from the membrane. In this case, the initial gold segment in the membrane was 1.2 μm long and the growth time sufficiently short (15 min) so

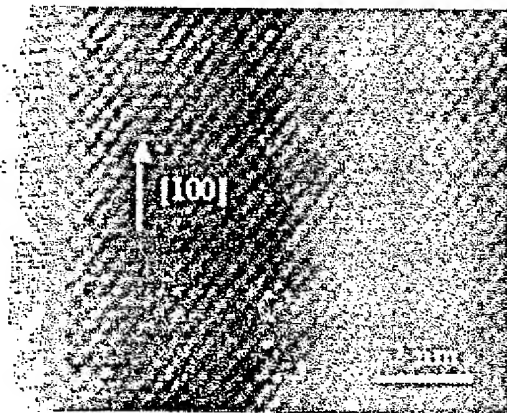


FIG. 3. High-resolution transmission electron microscopy (HREM) image near the edge of a nanowire. A thin (<3 nm) sheath of amorphous native oxide was found at the edge of the nanowire.

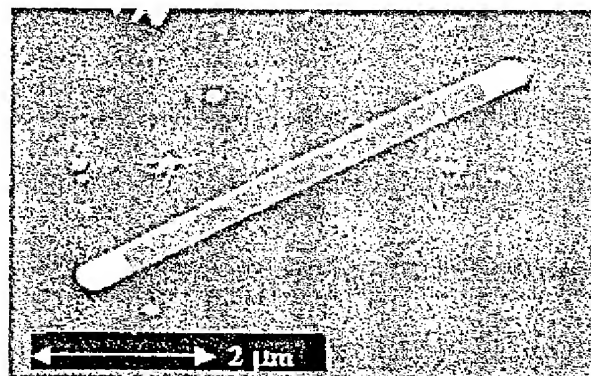


FIG. 4. Scanning electron microscopy image of a Au-Si-Au nanowire after release from the membrane by wet etching. Both ends of the nanowire are capped with approximately $0.6 \mu\text{m}$ long segments of Au.

that the length of the wire did not extend beyond the top surface of the membrane. As shown in Fig. 4, the nanowire consists of an approximately $4 \mu\text{m}$ long central region of Si with $0.6 \mu\text{m}$ long segments of Au at each end of the nanowire. Gold tips are always observed at both ends of the Si nanowires produced in this process. These results indicate that Si nucleates near the center of the Au-Si liquid alloy in this process and the crystal grows outward along the pore axis, leaving a segment of Au at each end of the nanowire. This is in contrast to Si nanowires produced by VLS growth from a surface¹⁸ or by laser ablation²⁰ in which the metal particle is present at only one end of the wire. Interactions between the liquid alloy drop and the internal pore walls are likely responsible for the modification in VLS growth observed in this study.

IV. CONCLUSIONS

The fabrication of Si nanowires has been demonstrated using a combination of template-directed synthesis and VLS growth. The use of nanopores for VLS growth provides control over nanowire diameter while also enabling the production of single crystal material. This work demonstrated the synthesis of nominally 200 nm diameter Si nanowires in nanoporous anodic alumina membranes using electrodeposited Au as the metal catalyst for VLS growth. The nanowires consist of a crystalline Si core with a $\langle 100 \rangle$ or $\langle 211 \rangle$ growth direction surrounded by a thin (<3 nm) oxide coating. The ends of the Si nanowires are capped with segments of Au, which result from the VLS mechanism of growth. It is anticipated that this technique can be extended to the synthesis of smaller diameter nanowires through the use of anodized aluminum oxide templates with pore sizes in the range of 10–100 nm.⁹ The results of this study demonstrate the potential for using template-directed VLS growth in combination with metal electrodeposition for the fabrication of single crystal Si nanowires and multi-layered metal-Si-metal nanowires.

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